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1. NANOPHYSICS

1.1. Phenomena and Effects

1.1.1.1. Heavy hole- and light hole-trion states in ellipsoidal quantum dot. /Y. Bleyan, D. Hayrapetyan/.

Nano Studies. – 2019. – #19. – pp. 7-10. – eng.; abs.: eng.

The heavy hole and light hole negative trions, which consist of two electrons bound by a hole, are considered in strongly prolate ellipsoidal quantum dot. Fig. 3, Ref. 9.

Keywords: heavy hole, light hole, negative trions, ellipsoidal dot

References

1. W. Xie. Chem. Phys., 2012, 408, 69-74.
2. M. A. Lampert. Phys. Rev. Lett., 1958, 1, 450.
3. B. Stebe, A. Ainane. Superlatt. & Microstruc., 1989, 5, 545.
4. A.J. Shields, M. Pepper, M.Y. Simmons, D.A. Ritchie. Phys. Rev. B, 1995, 52, 7841.
5. B. Stebe, A. Ainane, F. Dujardin. J. Phys. Cond. Matter, 1996, 8, 5383.
6. W. Xie. SolidState Comm., 2001, 118, 305.
7. E. Anisimovas, F. M. Peeters. Phys. Rev. B, 2003, 68, 115310.
8. C. Riva, F.M. Peeters, K. Varga. Phys. Rev. B, 2000, 61, 873.
9. B. Szafran, B. Stebe, J. Adamowski, S. Bednarek. Phys. Rev. B, 2002, 66, 165331.

1.1.1.2. Theoretical investigation of the Faraday-rotation effect in atomic alkali nano-layers. /A. Amiryan/.

Proceedings of NAS RA. Physics. – 2018. – v. 53. – #3. – pp. 281-293. – rus.; abs.: rus., arm., eng. Theoretical studies of the Faraday-rotation (FR) effect in alkali vapors contained in extremely thin cells, are presented. It is shown that the spectra of the FR signal are well frequency resolved despite the huge number of atomic transitions. This allows one to study the evolution of the Cs D_1 ($\lambda = 895$ nm) line hyperfine structure of $F_g = 4 \rightarrow F_e = 3, 4$ atomic transitions in magnetic fields. The presented theoretical model predicts the coherent Dicke narrowing effect and its revival with a periodicity $L = \lambda$. The practical applications of the FR are noted. Fig. 6, Ref. 19.

Keywords: Faraday rotation, atomic alkali vapors, nano layers, Dicke effect

References

1. D. Budker, D.F. Kimball, D.P. DeMille. Atomic Physics: an Exploration through Problems and Solutions. Oxford University Press, 2004.
2. W. Demtroder. Laser Spectroscopy: Basic Concepts and Instrumentation. Springer-Verlag, Berlin, 2004.
3. M. Fichet, G. Dutier, A. Yarovitsky, P. Todorov, I. Hamdi, I. Maurin, S. Saltiel, D. Sarkisyan, M.P. Gorza, D. Bloch, M. Ducloy. Europhys. Lett., 77, 54001 (2007).
4. А. Саргсян, Г. Ахумян, А. Амирян, К. Леруа, Д. Саркисян. Известия НАН Армении, Физика, 50, 428 (2015).
5. А. Саргсян, А. Амирян, С. Карталева, Д. Саркисян. ЖЭТФ, 125, 43 (2017).
6. А. Саргсян, А. Амирян, К. Леруа, Т. Вартанян, П. Петров, Д. Саркисян. Оптический журнал, 83, 11 (2016).
7. А. Саргсян, Г. Ахумян, А. Амирян, А. Саркисян, Д. Саркисян. Известия НАН Армении, Физика, 51, 424 (2016).
8. А. Саргсян, А. Амирян, К. Леруа, Т. Вартанян, Д. Саркисян. Оптика и спектроскопия, 123, 113 (2017).
9. А. Саргсян, Р. Мирзоян, А. Саркисян, А. Амирян, Д. Саркисян. Известия НАН Армении, Физика, 49, 31 (2014).
10. А. Саргсян, Е. Пашаян-Леруа, К. Леруа, Ю. Малакян, Д. Саркисян. Письма в ЖЭТФ, 102, 549 (2015).
11. A. Sargsyan, E. Klinger, G. Hakumyan, A. Tonoyan, A. Papoyan, C. Leroy, D. Sarkisyan. JOSAB, 34, 776 (2017).
12. А. Саргсян, А. Амирян, Т. Вартанян, Д. Саркисян. Оптика и спектроскопия, 121, 790 (2016).
13. K. McDonald. Physics Examples and Other Pedagogic Diversions. Faraday Rotation, <http://www.physics.princeton.edu/mcdonald/examples/>
14. D. Budker, W. Gawlik, D. Kimball, S.R. Rochester, V.V. Yashchuk, A. Weis. Rev. Mod. Phys., 74, 1153 (2002).
15. B. Zambon, G. Neinhuis. Opt. Commun., 143, 308 (1997).
16. G. Dutier, S. Saltiel, D. Bloch, M. Ducloy. JOSA B, 20, 793 (2003).

17. P. Tremblay, A. Michaud, M. Levesque, S. Thériault, M. Breton, J. Beaubien, N. Cyr. Phys. Rev. A, 42, 2766 (1990).
18. D. A Steck. Cesium D Line Data, <http://steck.us/alkalidata/cesiumnumbers.pdf>
19. A. Sargsyan, Y. Pashayan-Leroy, C. Leroy, D. Sarkisyan. J. Phys. B: Atom., Mol. Opt. Phys., 49, 075001 (2016).

1.1.1.3. Numerical Study of Josephson nanostructures Using Parallel Computing. (Proceedings of the Int. Conference on “The Problems of Modern Condensed Matter Physics”). /I. Rahmonov, E. Zemlyanaya, M. Bashashin, P. Atanasova, A. Rahmonova, Yu. Shukrinov/. Armenian Journal of Physics. – 2019. – v. 12. – #3. – pp. 233-239. – rus.; abs.: eng.

We investigate the phase dynamics of the stack of long JJs, the length of which exceeds the Josephson penetration depth λ_J , taking into account the inductive and capacitive couplings between junctions and diffusion current. Numerical simulation of current-voltage characteristics of the stack is based on numerical solution of a system of nonlinear partial differential equations by the fourth order Runge-Kutta method and finite-difference approximation. The calculations are performed using the MPI technique for parallel implementation. The methodical calculations on multi-processor cluster (LIT JINR) with a different number of parallel MPI-processes are carried out. We have shown that the developed parallel algorithm provides about 7 time acceleration in comparison with serial simulation. Fig. 2, Tab. 1, Ref. 22.

Keywords: Josephson junction, inductive coupling, capacitive coupling

References

1. R. Kleiner, F. Steinmeyer, G. Kunkel, P. Müller, Phys. Rev. Lett. 68 (1992) 2394.
2. Yu.M. Shukrinov and F. Mahfouzi, Phys. Rev. Lett. 98 (2007) 157001.
3. Yu.M. Shukrinov, F. Mahfouzi, and M. Suzuki, Phys. Rev. B 78 (2008) 134521.
4. I.R. Rahmonov, Y.M. Shukrinov, A. Irie, JETP Letters 99 (2014) 632.
5. Yu.M. Shukrinov, A.E. Botha, S.Yu. Medvedeva, M.R. Kolahchi and A. Irie, Chaos 24 (2014) 033115.
6. T.A. Fulton and R.C. Dynes, Solid St. Commun. 12 (1972) 57.
7. N.F. Pedersen and D. Welner, Phys. Rev. B 29 (1984) 2551.
8. D.W. McLaughlin and A.C. Scott, Phys. Rev. A 18, (1978) 1652.
9. S. Lin, X. Hu, Phys. Rev. Lett. 100 (2008) 247006.
10. R. Kleiner, T. Gaber, G. Hechtfischer, Phys. Rev. B 62 (2000) 4086.
11. Y. Matsuda, M.B. Gaifullin, K. Kumagai, K. Kadokawa, T. Mochiku, Phys. Rev. Lett. 75, (1995) 4512.
12. A.A. Yurgens, Supercond.Sci.Technol. 13, (2000) R85.
13. T.M. Benseman, A.E. Koshelev, K.E. Gray, W.K. Kwok, U. Welp, K. Kadokawa, M. Tachiki, T. Yamamoto, Phys. Rev. B 84 (2011) 064523.
14. L. Ozyuzer, A.E. Koshelev, C. Kurter et al., Science 318 (2007) 1291.
15. I.R. Rahmonov, Yu.M. Shukrinov, P.Kh. Atanasova, E.V. Zemlyanaya, M.V. Bashashin, JETP 124(131 (2017).
16. I.R. Rahmonov, Yu. Shukrinov, P. Atanasova, E.V. Zemlyanaya, O.I. Streletsova, M. Zuev, A. Plecenik, and A. Irie, EPJ Web of Conferences 173, 06011 (2018).
17. S. Sakai, P. Bodin, N.F. Pedersen, J. Appl. Phys. 73 (1993) 2411.
18. M. Machida, S. Sakai, Phys. Rev. B 70 (2004) 144520.
19. Yu.M. Shukrinov, F. Mahfouzi, Physica C 434 (2006) 6.
20. Y.M. Shukrinov, I.R. Rahmonov, JETP 115 (2012) 289.
21. Y.M. Shukrinov, F. Mahfouzi, Phys. Rev. Lett. 98 (2007) 157001.
22. Y.M. Shukrinov, M. Hamdipour, JETP Letters 95 (2012) 307.

1.2. Properties of Materials and Structures

1.1.2.1. Structural studies of polyimide films; size effect. /V. Geidarov, I. Braude, N. Gal'tsov, Yu. Pohribnaya, V. Lototskaya, N. Aksanova/. Nano Studies. – 2019. – #19. – pp. 11-14. – eng.; abs.: eng. Amorphous polymerized films of 4,4'-difinyleneoxidepyromellitimide with thickness of 125 and 75 μm , after stretching, were investigated by the structural method. The findings of the study results in that stretching of polymer films with thickness of 75 μm do not cause cardinal changes in the structure, while

the deformation of polymer films with thickness of 125 μm leads to the appearance of regions with long-range order in the sample. Fig. 3, Ref. 17.

Keywords: polyimide films, side effects, long-range order

References

1. V.V. Danchuk, N.N. Gal'tsov, M.A. Strzhemechny, A.I. Prokhvatilov. Low Temp. Phys., 2004, 30, 118.
2. L.A. Alekseeva, M.A. Strzhemechny. Low Temp. Phys., 2012, 38, 534.
3. Yu.A. Freiman, H.J. Jodl, Y. Crespo. Phys. Rep., 2018, 743, 1.
4. N.N. Galtsov, O.A. Klenova, M.A. Strzhemechny. Low Temp. Phys., 2002, 28, 365.
5. G.A. Vdovichenko, A.I. Krivchikov, O.A. Korolyuk, O.O. Romantsova. Low Temp. Phys., 2014, 40, 1112.
6. J.F. Gebbia, M.A. Ramos, D. Szewczyk, A. Jezowski, A.I. Krivchikov, Y.V. Horbatenko, T. Guidi, F.J. Bermejo, J.L. Tamarit. Phys. Rev. Lett., 2017, 119, 215506.
7. V.G. Manzhelii, A.I. Prokhvatilov, V.G. Gavrilko, A.P. Isakina. Handbook of Structure and Thermodynamic Properties of Cryocrystals, 1999, New York, Begell House Inc.
8. R. Radebaugh, N.V. Frederick, J.D. Siegwarth. Cryogenics, 1973, 13, 43.
9. I.S. Braude, N.N. Gal'tsov, V.G. Geidarov, G.I. Kirichenko, V.V. Abraimov. Low Temp. Phys., 2016, 42, 204.
10. I.S. Braude, N.N. Gal'tsov, V.G. Geidarov, G.I. Kirichenko, V.A. Lototskaya, Yu. M. Plotnikova. Low Temp. Phys., 2017, 43, 1226.
11. V.A. Karachevtsev, A.M. Plokhotnichenko, A.Yu. Glamazda, V.S. Leontiev, I.A. Levitsky. Phys. Chem. 2014, 16, 10914.
12. V.I. Gul. Materials of Polymer Films, 1972, Moscow, Khimya.
13. V.A. Lototskaya, L.F. Yakovenko, E.N. Aleksenko, V.V. Abraimov, W.Zh. Shao. East. Eur. J. Phys., 2017, 4, 44.
14. E. Lifshin. X-ray Characterization of Materials, 1999, New York, Wiley-VCH.
15. L.I. Mirkin. Handbook of X-ray Analysis of Powder Crystals, 1961, Moscow, Fizmatgiz.
16. S. Isoda, H. Shimada, M. Kochi, H. Kambe. J. Pol. Sci., 1981, 19, 1293.
17. V.G. Geidarov, I.S. Braude, N.N. Gal'tsov, Yu.M. Pohribnaya. Mol. Crys. & Liquid Crys., 2018, 661, 20.

1.1.2.2. About physical nature of some peculiarities of primary cosmic radiation nuclei and gamma quanta spectra./T. Barnaveli, N. Eristavi, I. Khaldeeva/. Nano Studies. – 2019. – #19. – pp. 203-212. – eng.; abs.: eng.

About twenty years ago authors published the data concerning some peculiarities of the behavior of cosmic radiation EAS hadron component spectra. The results pointed to the possible existence in the interstellar space of the background of weakly interacting objects of the mass (the energy of the resonance oscillations) of the order of 37 eV. On the other hand, the experimental data of the last years are pointing to the existence of cosmic gamma radiation with the specific spectrum having the steep right front again in the region of the order of 37 eV and the left front falling down to the energies of the order less than 10^{-6} eV. Obviously, no elementary object may possess such spectrum of frequencies or a decay spectrum. Such spectrum may have some certain system or construction consisting of many elements possessing their own resonance frequencies and together composing the spectra observed. The possibility is presented of exactly such explanation of the cosmic rays primary radiation spectra peculiarities experimentally observed. It is based on the hypothesis about the discreteness of the space and existence in it of the topological defects distributed with sufficient density. In the frames of the proposed model some essential experimental peculiarities of the primary cosmic radiation nuclei and gamma quanta spectra find the unified explanation. Fig. 4, Ref. 15.

Keywords: cosmic radiation, gamma quanta spectra, EAS hadron, cosmic radiation, spectrum of frequencies

References

1. T.T. Barnaveli, I.V. Khaldeeva, Z.T. Shergelashvili, N.A. Eristavi. Phys. Lett. B, 1995, 346, 178.
2. T.T. Barnaveli, T.T. Barnaveli Jr, A.P. Chubenko, N.A. Eristavi, I.V. Khaldeeva, N.M. Nesterova, Yu.G. Verbetsky. Phys. Lett. B, 1996, 369, 372.

3. T.T. Barnaveli, T.T. Barnaveli Jr, N.A. Eristavi, I.V. Khaldeeva, Yu.G. Verbetsky. Phys. Lett. B, 1996, 381, 307.
4. T.T. Barnaveli, T.T. Barnaveli Jr, N.A. Eristavi, I.V. Khaldeeva, Yu.G. Verbetsky. In: Very high Energy Phenomena in the Universe, 1997, Les Arcs, Rencontres de Moriond.
5. T.T. Barnaveli, T.T. Barnaveli Jr, A.P. Chubenko, N.A. Eristavi, I.V. Khaldeeva, N.M. Nesterova, Yu.G. Verbetsky. ArXiv, 2002, astro-ph.0208275.
6. T.T. Barnaveli, T.T. Barnaveli Jr, N.A. Eristavi, I.V. Khaldeeva. ArXiv, 2003, astroph.0310524.v1.
7. Y. Inoue. ArXiv, 2014, astro-ph. HE. 1412. 3886v1.
8. B.A. Khrenov, M.I. Panasiuk. Priroda, 2006, 2, 17.
9. S. Hossenfelder. ArXiv, 2014, hep-ph.1401.0276.v1.
10. S. Hossenfelder. ArXiv, 2014, hep-ph.1309.0311.v2.
11. S. Hossenfelder. ArXiv, 2014, hep-ph.1309.0314.v2.
12. A.B. Migdal. J. Exp. & Theo. Phys., 1971, 61, 2209.
13. A.B. Migdal. Uspekhi - Phys., 1977, 103, 369.
14. Ya.B. Zeldovich, V.S. Popov. Uspekhi - Phys., 1971, 105, 6.
15. T.T. Barnaveli, N.A. Eristavi, I.V. Khaldeeva. Arxiv, 2016, astro-ph.HE.1604.04152.

1.1.2.3. Interpolation formula of magnetic nanofluids magnetization. /A. Ugulava, S. Chkhaidze, G. Mcchedlishvili, R. Abramishvili/. Nano Studies. – 2019. – #19. – pp. 37-44. – eng.; abs.: eng.

The internal energy of the magnetic anisotropy of some nanoparticles predominates over the thermal energy even at room temperature. It is known that the magnetic anisotropy axes of the magnetic nanofluids of nanoparticles, which in the absence of a magnetic field are oriented at random, at a sufficiently strong magnetic field, are polarized along the magnetic field and the so-called mechanical anisotropy state originates. In this paper, it is shown that the magnetization curve of the subsystem of polarized particles in the case when the anisotropy energy exceeds thermal energy has a special shape similar to the shape of a hyperbolic tangent. In the present model of a magnetic nanofluid it consists of two components (subsystems) with a variable number of particles: a subsystem of particles with randomly directed axes and subsystem of particles with polarized axes. The change of the magnetic field value causes the change of the number of particles in the subsystems followed by the change the magnetization of the whole magnetic nanofluids. For a given value of the magnetic field interpolation formula of magnetic nanofluids magnetization is found from the condition that the chemical potentials of subsystems are equal. It is shown that the magnetization curve obtained on the basis of a two-component model of magnetic nanofluids, is located between the Langevin curve and the hyperbolic tangent and with increasing anisotropy takes progressively the hyperbolic tangent shape. It is also shown that in strong magnetic anisotropy saturated state of magnetization of nanoparticles occurs much earlier than it follows from the Langevin theory. This significantly increases the Curie constant. Fig. 1, Ref. 9.

Keywords: interpolation formula, magnetic nanofluids, Langevin curve, Langevin theor

References

1. X. Batlle, A. Labarta. J. Phys. D: Appl. Phys., 2002, 35, R15-R42.
2. R. Skomski. J. Phys.: Cond. Matter, 2003, 15, R841-R896.
3. M.I. Shliomis. Sov. Phys. Uspekhi, 1974, 112, 153-169.
4. S.P. Gubin, Yu.A. Koksharov, G.B. Khomutov, G.Yu. Yurkov. Russ. Chem. Rev., 2005, 74, 489-520.
5. S.V. Vonsovskij. Magnetism, 1974, New York. John Wiley.
6. S.P. Gubin, Yu.A. Koksharov. Inorg. Mater., 2002, 38, 1085-1099.
7. F. Schwabl. Statistical Mechanics. 2006, Berlin–Heidelberg, Springer.
8. M. Abramowitz, I.A. Stegun. Handbook of Mathematical Functions with Formulas, Graphs and Mathematical Tables, 1972, New York, Dover, 295-319.
9. A. Ugulava, S. Chkhaidze, Sh. Kekutia, M. Verulashvili. Physica B: Cond. Mater., 2014, 454, 249-252.

1.1.2.4. Role of interfaces in spin dynamics and flows in magnetic nanostructures with normal metal sublayers. /A. Korostil/. Nano Studies. – 2019. – #19. – pp. 45-64. – eng.; abs.: eng.

The interconnection between the spin current and spin dynamics via the spin-dependent scattering and an accompanying by spin torque effect in ferromagnetic/normal metal based magnetic multilayer nanostructures is studied including a high fast out-of-equilibrium spin dynamics. Features of the spin transport through interfaces and its impact on spin dynamics are described on the base of the scattering matrix formalism for spin flows. The dependence of the spin torque effect on conductance character of the normal metal layers is considered. The exchange processes between the itinerant s and the localized d electrons are described by kinetic rate equations for electron-magnon spin-flop scattering. It is shown that the magnon distribution function remains nonthermalized on the relevant time scales of the demagnetization process, and the relaxation of the out-of-equilibrium spin accumulation among itinerant electrons provides the principal channel for dissipation of spin angular momentum from the combined electronic system. Fig.7, Ref. 48.

Keywords: spin dynamics, flow, magnetic nanostructures, magnon distribution, out-of-equilibrium spin

References

1. J. Akerman. Science, 2005, 308, 508.
2. L. Berger. Phys. Rev. B, 1996, 54, 9353.
3. J. Slonczewski. J. Magn. Magn. Mater., 1996, 159, L1.
4. M. Tsoi, A. G. M. Jansen, J. Bass, W.C. Chiang, M. Seck, V. Tsoi, P. Wyder. Phys. Rev. Lett., 1998, 80, 4281.
5. Z. Sun. J. Magn. Magn. Mater., 1999, 202, 157.
6. E.B. Myers, D.C. Ralph, J.A. Katine, R.N. Louie, R.A. Buhrman. Science, 1999, 285, 867.
7. J.A. Katine, F.J. Albert, R.A. Buhrman, E.B. Myers, D.C. Ralph. Phys. Rev. Lett., 2000, 84, 3148.
8. T. Valet, A. Fert. Phys. Rev. B, 1993, 48, 7099.
9. H.E. Camblong, P.M. Levy, S. Zhang. Phys. Rev. B, 1995, 51, 16052.
10. M.D. Stiles. Phys. Rev. B, 1996, 54, 14679.
11. K.M. Schep, J.B.A.N. van Hoof, P.J. Kelly, G.E.W. Bauer, J.E. Inglesfield. Phys. Rev. B, 1997, 56, 10805.
12. A. Brataas, Y.V. Nazarov, G.E.W. Bauer. Phys. Rev. Lett., 2000, 84, 2481.
13. X. Waintal, E.B. Myers, P.W. Brouwer, D.C. Ralph. Phys. Rev. B, 2000, 62, 12317.
14. G.M. Choi, B.C. Min, K.J. Lee, D.G. Cahill. Nat. Commun., 2014, 5, 4334.
15. A. Kirilyuk, A.V. Kimel, T. Rasing. Rev. Mod. Phys., 2010, 82, 2731.
16. A. Kirilyuk, A.V. Kimel, T. Rasing. Rep. Prog. Phys., 2013, 76, 026501.
17. E.M. Lifshitz, L.P. Pitaevskii. Course of Theor. Physics, Vol. 9: Statistical Physics, Part 2, 1980, Oxford, Pergamon.
18. J. Xiao, G.E.W. Bauer, K.C. Uchida, E. Saitoh, S. Maekawa. Phys. Rev. B, 2010, 81, 214418.
19. G.E. Bauer, E. Saitoh, B.J. van Wees. Nat. Mater., 2012, 11, 391.
20. B. Koopmans, J.J.M. Ruigrok, F. Dalla Longa, W.J.M. de Jonge. Phys. Rev. Lett., 2005, 95, 267207.
21. J. Walowski, G. Muller, M. Djordjevic, M. Munzenberg, M. Klaui, C.A.F. Vaz, J.A.C. Bland. Phys. Rev. Lett., 2008, 101, 237401.
22. A. Brataas, A.D. Kent, H. Ohno. Nat. Mater., 2012, 11, 372.
23. E.G. Tveten, A. Brataas. Phys. Rev., 2015, 92, 180412.
24. L. Berger. J. Appl. Phys., 1984, 55, 1954.
25. C.W.J. Beenakker. Rev. Mod. Phys., 1997, 69, 731.
26. K.B. Hathaway, J.R. Cullen. J. Magn. Magn. Mater., 1992, 104-107, 1840.
27. J.C. Slonczewski. J. Magn. Magn. Mater., 1993, 126, 374.
28. P. Bruno. Phys. Rev. B, 1995, 52, 411.
29. D.M. Edwards, F. Federici, J. Mathon, A. Umerski. Phys. Rev. B, 2005, 71, 054407.
30. A. Brataas, Yu.V. Nazarov, G.E.W. Bauer. Phys. Rev. Lett., 2000, 84, 2481.
31. P.W. Brouwer, C.W.J. Beenakker. J. Math. Phys., 1996, 37, 4904.
32. J. Slonczewski. J. Magn. Magn. Mater., 1999, 195, L261.
33. E. Beaurepaire, J.C. Merle, A. Daunois, J.Y. Bigot. Phys. Rev. Lett., 1996, 76, 4250.
34. J.Y. Bigot, M. Vomir, E. Beaurepaire. Nat. Phys., 2009, 5, 515.
35. A.J. Schellekens, B. Koopmans. Phys. Rev. B, 2013, 87, 020407.
36. A. Eschenlohr, M. Battiatto, P. Maldonado, N. Pontius, T. Kachel, K. Holldack, R. Mitzner, A. Fohlisch, P.M. Oppeneer, C. Stamm. Nat. Mater., 2013, 12, 332.
37. C. Illig, M. Haag, M. Fahnle. Phys. Rev. B, 2013, 88, 214404.

38. Y. Zhang, T.H. Chuang, K. Zakeri, J. Kirschner. Phys. Rev. Lett., 2012, 109, 087203.
39. Y. Tserkovnyak, A. Brataas, G.E.W. Bauer, B.I. Halperin. Rev. Mod. Phys., 2005, 77, 1375.
40. S.A. Bender, Y. Tserkovnyak. Phys. Rev. B, 2015, 91, 140402.
41. A.P. Jauho, N.S. Wingreen, Y. Meir. Phys. Rev. B, 1994, 50, 55285544.
42. B.Y. Mueller, A. Baral, S. Vollmar, M. Cinchetti, M. Aeschlimann, H.C. Schneider, B. Rethfeld. Phys. Rev. Lett., 2013, 111, 167204.
43. J. Barker, U. Atxitia, T.A. Ostler, O. Hovorka, O. Chubykalo-Fesenko, R.W. Chantrell. Sci. Rep., 2013, 3, 3262.
44. Y. Tserkovnyak, E.M. Hankiewicz, G. Vignale. Phys. Rev. B, 2009, 79, 094415.
45. R. Meservey, P.M. Tedrow. Phys. Rev. Lett., 1978, 41, 805.
46. U. Atxitia, O. Chubykalo-Fesenko, J. Walowski, A. Mann, M. Munzenberg. Phys. Rev. B, 2010, 81, 174401.
47. A. Weber, F. Pressacco, S. Gunther, E. Mancini, P.M. Oppeneer, C.H. Back, Phys. Rev. B, 2011, 84, 132412.
48. Y. Liu, Z. Yuan, R.J.H. Wesselink, A.A. Starikov, P.J. Kelly. Phys. Rev. Lett., 2014, 113, 207202.

1.1.2.5. Activation volume of plastic deformation of ultrafine-grained copper. /T. Hryhorova/. Nano Studies. – 2019. – #19. – pp. 65–76. – eng.; abs.: eng.

Tensile and stress relaxation tests in the temperature range 77–295 K were carried out to study the mechanical properties and kinetics of plastic deformation of ultrafine-grained (UFG) copper prepared by equal-channel angular hydroextrusion. The temperature and strain-rate sensitivities of flow stress were analyzed to identify main mechanisms of plastic deformation. It was shown that as temperature increases the flow stress decreases but the activation volume changes non-monotonously reaching the peak value at about 200 K. In contrast to conventional copper polycrystals both temperature dependencies for UFG copper can be explained in terms of two mechanisms of thermally activated plastic deformation: “forest” intersection and dynamic recovery. The latter mechanism is necessary to take into account for a correct estimation of activation volume from stress relaxation tests even at low temperatures. Tab. 1, Fig. 7, Ref. 25.

Keywords: plastic deformation, ultrafine-grained copper, peak value, activation volume

References

1. U.F. Kocks, A.S. Argon, M.F. Ashby. Progr. Mater. Sci., 1975, 19, 288.
2. J. Bonneville, B. Escaig, J. L. Martin. Acta Metall., 1988, 36, 1989.
3. H. Conrad. Mater. Sci. Eng. A, 2003, 341, 216.
4. H. Conrad, K. Jung. Scripta Mater., 2005, 53, 581.
5. H. Conrad, D. Yang. J. Electronic Mater., 2002, 31, 4.
6. J. Chen, L. Lu, K. Lu. Scripta Mater., 2006, 54, 1913.
7. M. Kato, T. Fujii, S. Onaka. Mater. Trans., 2008, 49, 1278.
8. W. Blum. Scripta Mater., 2018, 146, 27.
9. B. Wielke. Acta Metall., 1978, 26, 103.
10. C.M. Kuo, C.H. Lin, Y.C. Huang. Mater. Sci. Eng. A, 2005, 396, 360.
11. A. Beloshenko, V.N. Varyukhin, V.Z. Spuskanyuk. Theory and Applications of Hydrostatic Extrusion, 2007, Kyiv, Naukova Dumka.
12. N.V. Isaev, T.V. Grigorova, O.V. Mendiuk, O.A. Davydenko, S.S. Polishchuk, V.G. Geidarov. Low Temp. Phys., 2016, 42, 825.
13. N.V. Isaev, T.V. Grigorova, S.E. Shumilin, S.S. Polishchuk, O.A. Davydenko. Low Temp. Phys., 2017, 43, 1420.
14. N.V. Isaev, T.V. Hryhorova, O.A. Davydenko, S.S. Polishchuk. Low Temp. Phys., 2018, 44, 1540.
15. G. Sargent, G. Jones, H. Conrad. Scripta. Met., 1969, 3, 481.
16. S. Oculo, H. Abe, Y. Miyajima, T. Fujii, S. Onaka, M. Kato. Mater. Trans., 2014, 55, 1525.
17. J. Gubicza, N.Q. Chinh, J.L. Labar, S. Dobatkin, Z. Hegedus, T.G. Langdon. J. Alloys & Comp., 2009, 483, 271.
18. V.I. Dotsenko, A.I. Landa. Mater. Sci. Eng., 1976, 22, 101.
19. M. Bocek, H. Schneider, P. Sindelar. Mater. Sci. Eng., 1971, 8, 161.
20. J. Friedel. Dislocations, 1964, London, Pergamon Press.
21. S. Basinski. Phil. Mag., 1959, 4, 393.
22. T. Kruml, O. Coddet, J.L. Martin. Acta Mater., 2008, 56, 333.
23. E.W. Hart, H.D. Solomon. Acta Metall., 1973, 21, 295.
24. P. Groh, R. Conte. Acta Metall., 1971, 19, 895.
25. S.V. Lubenets. Phys. Solid States, 2002, 44, 72.

1.1.2.6. GaN monolayer films on GaAs surface (001): Obtaining and optical studies. /V. Berkovits, V. Ulin, G. Iluridze, T. Minashvili, K. Davitadze, A. Gigineishvili, Z. Jabua/. Nano Studies. – 2019. – #19. – pp. 111-114. – rus.; abs.: rus.

Using anisotropic reflection spectroscopy, monolayer films of gallium nitride formed on the GaAs crystals (1) surface were studied by chemical nitridation in hydrazine-sulphide solutions. It was found that the anisotropic reflection spectra of nitrided GaAs samples contain only volume anisotropy signals in the region of the E_1 and $E_1+\Delta_1$ transitions in GaAs, and the layer of nitrogen atoms on the surface is optically isotropic. It was shown as well that as a result of nitridation, the bending of zones near the surface of n-type GaAs crystals decreases by ~25%. Fig. 2, Ref. 8.

Keywords: gallium nitride (GaN) monolayer, GaAs surface, nitrogen atoms, hydrazine-sulphide solution

References

1. H. Okumura, K. Ohta, G. Feuillet, K. Balakrishnan, S. Chichibu, H. Hamagushi, P. Hacke, S. Yoshida. J. Cryst. Growth, 1997, 178, 113.
2. V.L. Berkovits, V.P. Ulin, M. Losurdo, P. Capezzuto, G. Bruno. Appl. Phys. Lett., 2002, 80, 3739.
3. V.L. Berkovits, D. Paget, A.N. Karpenko, V.P. Ulin, O.E. Tereshchenko. Appl. Phys. Lett., 2007, 90, 022103.
4. V.L. Berkovits, V.N. Bessolov, T.N. L'vova, V.I. Safarov, R.V. Khasieva, B.V. Tsarenkov. J. Appl. Phys., 1991, 70, 3707.
5. P. Weightman, D.S. Martin, R.J. Cole, T. Farrel. Rep. Prog. Phys., 2005, 68, 1251.
6. V.L. Berkovits, V.A. Kosobukin, A.B. Gordeeva. J. Appl. Phys., 2015, 118, 245305.
7. В.Л. Берковиц, А.Б. Гордеева, Т.В. Львова, В.П. Улин. ФТП, 2012, 46, 1463.
8. В.Л. Берковиц, В.Н. Бессолов, Т.В. Львова, Е.Б. Новиков, В.Л. Сафаров, Р.В. Хасиева, Б.В. Царенков. ФТП, 1991, 25, 1406.

1.1.2.7. Location of interstitial light atoms in vanadium hydrides (deuterides). /N. Namoradze, I. Ratishvili/. Nano Studies. – 2019. – #19. – pp. 125-130. – eng.; abs.: eng.

The local mode frequencies of the light atoms inserted in the interstitial positions of a bcc metal lattice are considered. A significant difference between the oscillations of a light atom located in the octahedral and in the tetrahedral interstitial positions is indicated. An example of the $V_2H(D)$ system is analyzed and it is suggested that the neutron inelastic diffraction experiments indicate that the light atoms are located in the octahedral interstitial positions. Fig. 4, Ref. 12.

Keywords: interstitial light atoms, vanadium, neutron inelastic diffraction, octahedral positions

References

1. Eds. G. Alefeld, J. Volkl. Hydrogen in Metals, I. Topics Appl. Phys., 28, 1978, Springer Verlag.
2. Eds. G. Alefeld, J. Volkl. Hydrogen in Metals, II. Topics Appl. Phys., 29, 1978, Springer Verlag.
3. В.А. Соменков, С.Ш. Шильшткін. Фазовые превращения водорода в металле (Препринт ИАЕ им. И. Курчатова), 1978, Москва.
4. T. Schober, H. Wenzl. The systems NbH(D), VH(D), TaH(D): Structures, phase diagrams, morphologies, methods of preparation". In: Hydrogen in Metals, II (Eds. G. Alefeld, J. Volkl), Topics Appl. Phys., 29, 1978, Springer Verlag 11-71.
5. A.G. Khachaturyan. Structural Phase Transformations in Solids, 1983, New York, Wiley.
6. S. Hayashi. Diffusion of hydrogen isotopes and their mutual perturbation in $Ti_{0.33}V_{0.67}H_xDy$ ($x+y \approx 0.9$) studied by 1H and 2H NMR. J. Solid State Chem., 2003, 170, 87-91.
7. S. Hayashi. Deuterium diffusion in vanadium deuterides (VD_x ; $0.4 \leq x \leq 0.6$) studied by 2H NMR. J. Solid State Chem., 2004, 177, 824-833.
8. В.П. Тараков, Л.Н. Падурец, А.Л. Шилов, Г.А. Киракосян. Динамика дейтонов в у-фазе дейтеридов титана по данным ЯМР. ЖХХ, 2009, 54, 1365-1369.
9. В.П. Тараков, Г.А. Киракосян, Л.Н. Падурец. Ядерный магнитный резонанс 2H , $^{47,49}Ti$ в у-фазе дейтеридов титана TiD_x . ФТП, 2010, 52, 459-468.
10. С.Э. Хайкин. Физические основы механики, 1962, Москва, Гос. изд. физ.-мат. лит., Гл. 18.
11. J.M. Rowe. A neutron scattering study of the vibrational and diffusional motions of deuterium in the α and β phases of $VD_{0.5}$. Solid State Commun., 1972, 11, 1299-1302.

12. I.G. Ratishvili, N.Z. Namoradze. Temperature dependence of local mode frequencies in ordering metal hydrides. The case of b.c.c. metal lattices. Met. Phys. Novel Technol., 2011, 33, 919-938.

1.1.2.8. On experimental methods for studying of mechanical properties of solids. /M. Galustashvili, D. Driaev/. Nano Studies. – 2019. – #19. – pp. 213-222. – rus.; abs.: rus.

The original experimental methods of studying the mechanical properties and dislocation structure of solids developed at the E. Andronakashvili Institute of Physics are described. Among them are: (a) method of excitation of mechanical vibrations of the ion crystal (for example, LiF) by acting on charged dislocations; (b) method of the vibrating superconductor (for example, the reed of Nb in mixed state) demonstrating the dependence of the effective elastic modulus and dissipation of mechanical energy on the orientation of the Abrikosov vortices relative to the axis of bending of the reed; (c) method for determining the parameters of the relaxation process by measuring the internal friction of the crystal simultaneously at two frequencies of bending vibrations; (d) method of shear deformation in one system of crystallographic planes; and (e) method of bending deformation with dislocations of one mechanical sign. Fig. 9, Ref. 12.

Keywords: solids, mechanical properties, ion crystal, shear deformation, effective elasic modulus

References

1. Д.Г. Дриаев, В.А. Мелик-Шахназаров. ФТТ, 1966, 8, 3280.
2. В.И. Альшиц, Д.Г. Дриаев, В.А. Мелик-Шахназаров. В сб.: Электронные и ионные процессы в твердых телах, 1973, 6, 68.
3. Б.И. Смирнов. Дислокационная структура и упрочнение кристаллов, 1981, Ленинград, Наука.
4. М.В. Галусташвили, Д.Г. Дриаев, З.К. Саралидзе. Авт. свид. СССР 1432383 А1, 1986.
5. М.В. Галусташвили, Д.Г. Дриаев, И.А. Политов, З.К. Саралидзе. ФТТ, 1988, 30, 1533.
6. M.V. Galustashvili, D.G. Driaev, I.A. Politov, Z.K. Saralidze. Phys. Status Solidi A, 1989, 114, 99.
7. Дж.Г. Чигвинадзе. ЖЭТФ, 1974, 67, 2361.
8. Д.Г. Дриаев, Дж.Г. Чигвинадзе. ФНТ, 1976, 2, 1566.
9. P. Esguinazi. J. LowTemp. Phys., 1991, 85, 139.
10. Д.Г. Дриаев, В.А. Мелик-Шахназаров. Электронные и ионные процессы в твердых телах, 1973, 5, 84.
11. Д. Дриаев. Патент Грузии Р4711, 2007.
12. К. Зинер. В сб.: Упругость и неупругость металлов, 1954, Москва, ИИЛ, 74.

1.1.2.9. “Permittivity” of single-layer BN sheet. /L. Chkhartishvili/. Nano Studies. – 2019. – #19. – pp. 291-292. – eng.; abs.: eng.

Nanocapacitors, which are prospective nanoelectronic devices useful, in particular, for electrical energy storage purposes, use nanosheet dielectrics. Effective permittivity, which is an important performance parameter for any nanocapacitor, for the monolayer sheet differs from that of the bulk layered crystal of the same material. Reconstruction characteristic of polar dielectric interface layers seems to be the mechanism affecting apparent permittivity of the corresponding capacitor structure. There is presented theoretical estimate of “permittivity” of the single-layer boron nitride hexagonal sheet. Fig. 2, Ref. 6.

Keywords: nanocacitors, nanoelectronic devices, permittivity, boron nitride, hexagonal sheet

References

1. AZo Quantum. New technology for assembling graphene-based high-quality singleelectron transistors. 2019 March 7, Article ID 6341, 1-2.
2. L. Chkhartishvili, M. Beridze, Sh. Dekanosidze, R. Esiava, I. Kalandadze, N. Mamisashvili, G. Tabatadze. How to calculate nanocapacitance. Am. J. Nano Res. & Appl., 2017, 5, 3-1, 9-12.
3. L. Chkhartishvili. Nanoparticles near-surface electric field. Nanoscale Res. Lett., 2016, 11, 1, 48, 1-4.
4. L. Chkhartishvili. Nanostructure makes crystalline compound physically reactive. Atlas Sci., 2016 April 14, 1-2.

5. L. Chkhartishvili, L. Sartinska, Ts. Ramishvili. Ch. 8: Adsorption selectivity of boron nitride nanostructures designed for environmental protection. In: Advanced Environmental Analysis: Applications of Nanomaterials, 1 (Eds. Ch.M. Hussain, B. Kharisov), 2017, Cambridge, Royal Soc. Chem., 167-192.
6. L. Chkhartishvili, Sh. Dekanosidze, N. Maisuradze, M. Beridze, R. Esiava. Estimation of atomic charges in boron nitrides. Eastern-Eur. J. Ent. Technol., 2015, 3, 5 (75) – Appl. Phys., 50-57 & 63-64.

1.1.2.10. Spectrum of radiation of a bunch of positrons channeled in densely packed nanotubes. /K. Gevorgyan, L. Gevorgian/. Proceedings of NAS RA. Physics. – 2019. – v. 54. – #2. – pp. 172-184. – rus.; abs.: rus., arm., eng.

The problem of the radiation of a relativistic bunch of positrons channeled in densely packed nanotubes was solved. A model parabolic potential was used for the average nanotube field. The channeled positrons oscillate with the same frequency, but with different amplitudes equal to the distance from the point of entry of the positron in the nanotube to its axis. The polarization of the nanotube environment leads to the appearance of a lower energy threshold for the formation of radiation. The boundaries of the frequency interval of radiation also depend on the amplitude of the oscillating positron. For the first time an analytical expression for the spectrum of the total radiation is obtained in the dipole approximation. At a zero angle, both hard (gamma) and soft (X-ray) photons are generated. Directed photon beams have important practical applications. Fig. 1, Tab. 1, Ref. 24.

Keywords: radiation spectrum, positrons, nanotubes, parabolic potential, dipole approximation

References

1. В.Л. Гинзбург. Известия АН СССР, Сер. Физика, 11, 165 (1947).
2. Г. Мотц. Миллиметровые и субмиллиметровые волны. М. ИЛ., 1959, Применение излучения быстрых электронных пучков, 195-209 (1959).
3. R.P. Godwin. Springer Tracts in Modern Physics, 51, 1 (1969).
4. Н.А. Корхмазян. Изв. АН Арм. ССР, Физика, 5, 287; 5 418 (1970).
5. В.Л. Гинзбург. Письма ЖЭТФ, 16, 501 (1972).
6. Л.А. Геворгян, Н.А. Корхмазян. ЖЭТФ, 76, 1226 (1979).
7. Л.А. Геворгян, Н.А. Корхмазян. Научное сообщение ЕФИ, 66, 273 (1977).
8. L.A. Gevorgian, N.A. Korkhmazian. Phys. Lett. A, 75, 453 (1979).
9. Л.А. Геворгян, Н.А. Корхмазян. Авторское свидетельство № 784729, 1970.
10. В. А. Базылев, Н.К. Жеваго. Излучение быстрых частиц в веществе и во внешних полях, М., Наука, 1987.
11. J. Lindhard, Physics Uspekhi, 99, 249 (1969).
12. M.A. Kumakhov. Phys. Lett. A, 57, 17 (1976).
13. V.V. Kaplin, S.V. Plotnikov, S.A. Vorobiev. Zh. Tekh. Fiz., 50, 1079 (1980).
14. A.V. Korol, A.V. Soloviev, W. Greiner. Int. J. Mod. Phys., 8, 49 (1999).
15. L.A. Gevorgian, R.O. Avakian, K.A. Ispirian, R.A. Ispirian. JETP Letters, 68, 467 (1998).
16. L.A. Gevorgian, R.O. Avakian, K.A. Ispirian, A.H. Shamamian. Nuclear Instruments and Methods in Physics Research Section B, 227, 104 (2005).
17. R.O. Avakian, L.A. Gevorgian, K.A. Ispirian, R.K. Ispirian. Pisma Zh. Eksp. Teor. Fiz., 68, 437, (1998); Nucl. Instr. Meth. B, 173, 112 (2001).
18. S. Iijima. Nature, 354, 56 (1991).
19. L.A. Gevorgian, K.A. Ispirian, R.K. Ispirian. Nuclear Instruments and Methods in Physics Research Section B, 145, 155 (1998).
20. L. Gevorgian, L. Hovsepyan. Proc. SPIE, 6634, 663408 (2007).
21. L.A. Gevorgian. Intensive quasi-monochromatic, directed X-ray radiation of planar channeled positron bunch, In book "Charged and Neutral Particles Channeling Phenomena -Channeling 2008", World Scientific, 370-377 (2010).
22. Л.А. Геворгян, Л.А. Овсепян. Изв. НАН Армении, Физика, 42, 131 (2007).
23. Л.А. Овсепян. Изв. НАН Армении, Физика, 42, 138 (2007).
24. J.D. Jackson. Classical Electrodynamics, New York, Wiley, 3rd ed., 1999.

1.1.2.11. Determination of refractive index and thickness of nanosized amorphous carbon films via visible range reflectance spectra. /G. Dabaghyan, L. Matevosyan, K. Avjyan/. Proceedings of NAS RA. Physics. – 2019. – v. 54. – #2. – pp. 249-252. – rus.; abs.: eng.

The values of thickness and refractive index ($1.92 < nf < 2.19$) of amorphous nanosized carbon films obtained on a crystalline silicon substrate by pulsed laser deposition were experimentally determined via an analysis of visible range reflection spectra. Obtained films can be used as single-layer anti-reflective coatings for semiconductors Si and GaAs. Fig. 1, Tab. 1, Ref. 9.

Keywords: refractive index, nanosized amorphous carbon, reflectance spectra, anti-reflective coatings

References

1. R. Hauert. Diamond and Related Materials, 12, 583 (2003).
2. S.C.H. Kwok, J. Wang, P.K. Chu. Diamond and Related Materials, 14, 78 (2005).
3. A.C. Ferrari. Surface and Coatings Technology, 180-181, 190 (2004).
4. F.J. Pern, Zh. Panosyan, A.A. Gippius, et al. NREL/CP-520-37374. February 2005.
5. D. Poelman, P. Smet. J. Phys. D Appl. Phys., 36, 1850 (2003).
6. V.M. Aroutiounian, Kh.S. Martirosyan, P. Soukiassian. Journal of Physics D: Applied Physics, 37, L25 (2004).
7. V.M. Aroutiounian, Kh.S. Martirosyan, P. Soukiassian. Journal of Physics D: Applied Physics, 39, 1623 (2006).
8. A.S. Hovhannisyan. J. Contemp. Phys. (Armenian Ac. Sci.), 43, 136 (2008).
9. D.E. Aspnes, A.A. Studna. Phys. Rev. B, 27, 985 (1983).

1.1.2.12. Interband Absorption and Photoluminescence in Nanospherical InP/InAs/InP Core/Shell/Shell Heterostructure. /V. Harutyunyan, M. Mkrtchyan, E. Kazaryan, D. Hayrapetyan/. Proceedings of NAS RA. Physics. – 2019. – v. 54. – #1. – pp. 44-60. – rus.; abs.: rus., arm., eng.

Single-particle states of charge carriers in a nanospherical InP/InAs/InP heterostructure are theoretically considered in the isotropic effective mass approximation and in the regime of strong size quantization. Results of numerical calculations of the energy levels of charge carriers at different thicknesses of the quantizing InAs layer of the indicated heterophase structure are presented. It is shown that it is possible to achieve the desired value and position of the size quantization levels of charge carriers in the layer by appropriate choice of the layer thickness. Interband optical transitions in the InAs layer are also considered. Values of the effective broadening of the band gap of the InAs layer as a function of the layer thickness are calculated. It is shown by numerical calculations that the absorption has a resonant character and that diagonal transitions dominate in the spectrum of the interband absorption. Values of threshold frequencies and absorption curves for several diagonal transitions involving both light and heavy holes are given. The photoluminescence spectra in the spherical InP/InAs/InP nanoheterostructure were also constructed for various temperatures close to room temperature. Fig. 4, Tab. 9, Ref. 45.

Keywords: interband absorption, nanospherical InP heterostructure, InAs layer, band gap

References

1. N.G. Aghekyan, E.M. Kazaryan, A.A. Kostanyan, H.A. Sarkisyan. Superlattices and Microstructures, 50, 199 (2011).
2. A. Ferron, P. Serra, O. Osenda. Phys. Rev. B, 85, 165322 (2012).
3. В.А. Арутюнян, Д.Б. Айрапетян, Д.А. Багдасарян. Изв. НАН Армении, Физика, 51, 471 (2016).
4. В.А. Арутюнян, Д.Б. Айрапетян, Э.М. Казарян. Изв. НАН Армении, Физика, 53, 65 (2018).
5. V. Harutyunyan. Effect of Static Electric Fields on the Electronic and Optical Properties of Layered Semiconductor Nanostructures, PART I Effect of Static Electric Fields on the Electronic Properties of Layered Semiconductor Nanostructures, Bentham Science, 2015.
6. M. Henini. Handbook of Self Assembled Semiconductor Nanostructures for Novel Devices in Photonics and Electronics, Elsevier, 2011.
7. A.M. El-Toni, M.A. Habila, J.P. Labis, Z.A. Othman, M. Al hoshan, A.A. Elzatahry, F. Zhang. Nanoscale, 8, 2510 (2016).
8. J. Berezowsky, O. Gywat, F. Meier, D. Battaglia, X. Peng, D.D. Awschalom, Nature Physics, 2, 831 (2006).
9. D.T. Cat, A. Pucci, K. Wandlet, Physics and Engineering of New Materials, Springer, Berlin-Heidelberg 2009.

10. S. Brovelli, R.D. Schaller, S.A. Crooker, F. Garcia-Santamaria, Y. Chen, R. Vishvanatha, J.A. Hollingsworth, H. Htoon, V.I. Klimov. *Nature Commun.*, 2, Article Number 280 (2011).
11. A.M. Smith, L.A. Lane, S. Nie. *Nature Commun.*, 5, Article number: 4506 (2014).
12. K. Li, *Nanotechnology*, 1, 482 (2014).
13. C.S. S.R. Kumar, (Ed.), *Semiconductor nanomaterials*. John Wiley & Sons; pp. 393-427 (2010).
14. Micro Systems and Devices for (Bio)chemical Processes, Jr: *Chemical engineering*, vol.38, San Diego, Academic Press, 2010.
15. M. Rai, N. Duran, (Eds.), *Metal Nanoparticles in Microbiology*. Heidelberg-Dordrecht-London-New York: Springer Science & Business Media, 2011.
16. N. Waikopf, R. Rotem, I. Shweky, L. Yedidya, H. Soreq, U. Banin, *BioNanoScience*, 3, 1 (2013).
17. S. Ogli, A. Rostani. *IET Nanobiotechnology*, 7, 140 (2013).
18. J. Li, D. Wang, R. R. LaPierre, *Advances in III-V Semiconductor Nanowires and Nanodevices*, Bentham Science (2011).
19. S. Mokkapati, Ch. Jagadish. *Materials today*, 12, 22 (2009).
20. M. Fang, N. Han, F. Wang, Z-X. Yang, S.P. Yip, G. Dong, J.J. Hou, Y. Chueh, J.C. Ho. *Journal of Nanomaterials*, 2014, Article ID 702859 (2014).
21. PATENT WO 2007020416 A1, 22 Feb., 2007.
22. PATENT CA 2617972 C, 15 July, 2014.
23. K.J. Bachman. *Annual Review of Materials Science*, 11, 441 (1981).
24. I. Gyuro. *III-Vs Review*, 9, 30 (1996).
25. Ch. Ippen, T. Greco, A. Wedel. *Journ. Inf. Displ.*, 13, 91 (2012).
26. L. Froberg. *Growth, Physics, and Device Applications of InAs-based Nanowires*, Lund university, Sweden, 2008, 84 p.
27. R. Contreras-Guerrero, S. Wang, M. Edirisooriya, W. Priyantha, J.S. Rojas-Ramirez, K. Bhuwalka, G. Doornbos, M. Holland, R. Oxland, G. Vellianitis, M. van Dal, B. Duriez, M. Passlack, C.H. Diaz, R. Droopad. *Journ. of Crystal Growth*, 378, 117 (2013).
28. K. Xu, Y. Qi, Z. Gao, J. Li, X. Wang, Y. Zhang, Z. Han, E. Gao. *Integrated Ferroelectrics*, 167, 205 (2015).
29. V.I. Klimov. *Semiconductor and Metal Nanocrystals: Synthesis and Electronic and Optical Properties*, CRC Press, 2003, 500 p.
30. P. Mohan, J. Motohisa, T. Fukui. *Appl. Phys. Lett.*, 88, 013110 (2006).
31. P. Mohan, J. Motohisa, T. Fukui. *Appl. Phys. Lett.*, 88, 133105 (2006).
32. M. Helmi, N. Alouane, N. Chauvin, C. Chevallier. *Nanotechnology*, 22, 405702 (2011).
33. C.L. dos Santos, P. Piquini. *Journ. Appl. Phys.*, 111, 054315 (2012).
34. <http://www.ioffe.ru/SVA/NSM/Semicond/InP/basic.html>
35. <http://www.ioffe.ru/SVA/NSM/Semicond/InP/bandstr.html>
36. W.E. Buhro, V.L. Colvin. *Nature Materials*, 2, 138 (2003).
37. Y. Wang, X. Yang, T.C. He, Y. Gao, H.V. Demir, X.W. Sun, H.D. Sun. *Appl. Phys. Lett.*, 102, 021917 (2013).
38. <http://www.ioffe.ru/SVA/NSM/Semicond/InAs/basic.html>
39. <http://www.ioffe.ru/SVA/NSM/Semicond/InAs/bandstr.html>
40. M.H. Sun, E.S.P. Leong, A.H. Chin, C.Z. Ning, G.E. Cirlin, Yu.B. Samsonenko, V.G. Dubrovskii, L. Chuang, C. Chang-Hasnain. *Nanotechnology*, 21, 335705 (2010).
41. М. Абрамович, И. Стиган. *Справочник по специальным функциям*, М., 1979. 830 с.
42. В.Н. Неверов, А.Н. Титов. *Физика Нанокоррозионных систем*, УГУ им. А.М. Горького, Екатеринбург, 2008, ст. 48-57.
43. Л.Е. Воробьев, Е.Л. Ивченко, Д.А. Фирсов, В.А. Шалыгин. *Оптические свойства наноструктур*, Санкт-Петербург, Наука, 2001, ст. 88-95.
44. Э.М. Казарян, А.А. Костанян, А.А. Саркисян. *Изв. НАН Армении, Физика*, 42, 218 (2007).
45. P.K. Basu. *Theory of Optical Processes in Semiconductors*, Clarendon Press, Oxford, 1997, 464 pp.

1.1.2.13. Selective reflection of laser radiation from ultrathin layers of cesium atomic vapors confined in a nanocell. /A.D. Sargsyan, A.S. Sarkisyan, D. Sarkisyan/. Proceedings of NAS RA. Physics. – 2018. – v. 53. – #4. – pp. 402-416. – rus.; abs.: rus., eng.

The effect of van der Waals interaction of cesium atoms with sapphire windows of a nanocell was experimentally studied using the selective reflection process. The distance between the windows (thickness L) varied in the range 50-2000 nm and the nanocell was filled with vapors of cesium atoms. The C_3 coefficient of van der Waals interaction for Cs atoms and ($6S_{1/2} < 6P_{1/2}$ transition) the sapphire window of the nanocell is measured. It is shown that using a selective reflection spectrum it is possible to determine magnetic fields with a spatial resolution of 70 nm and, consequently, both homogeneous and highly gradient magnetic fields can be measured. Fig. 8, Ref. 24.

Keywords: laser radiation, ultrathin layers, nanocell, Cs atoms, selective reflection, the effect of van der Waals

References

- 1.D. Budker, W. Gawlik, D. Kimball, S.R. Rochester, V.V. Yaschuk, A. Weis. Rev. Mod. Phys. 74, 1153 (2002).
2. M. Auzinsh, D. Budker, S.M. Rochester. Optically Polarized Atoms: Understanding Light Atom Interactions. Oxford University Press, 2010.
3. W. Demtröder. Laser Spectroscopy: Basic Concepts and Instrumentation. Springer- Verlag, 2004.
4. T.A. Vartanyan, D.L. Lin. Phys. Rev. A, 51, 1959 (1995).
5. B. Zambon, G. Nienhuis. Opt. Commun., 143, 308(1997).
6. H. Failache, S. Saltiel, M. Fichet, D. Bloch, M. Ducloy. Phys. Rev. Lett., 83, 5467 (1999).
7. G. Dutier, S. Saltiel, D. Bloch, M. Ducloy. JOSA B, 20, 793(2003).
8. D. Bloch, M. Ducloy. Adv. At. Mol. Opt. Phys., 50, 91(2005).
9. А. Саргсян, Э. Клингер, Е. Пашиян-Леруа, К. Леруа, А. Папоян, Д. Саркисян. Письма в ЖЭТФ, 104, 222 (2016).
10. M. Fichet, G. Dutier, A. Yarovitsky, P. Todorov, I. Hamdi, I. Maurin, S. Saltiel, D. Sarkisyan, M.-P. Gorza, D., Bloch. Europhys. Lett., 77, 54001 (2007).
11. A. Sargsyan, E. Klinger, G. Hakumyan, A. Tonoyan, A. Papoyan, C. Leroy, D. Sarkisyan. JOSA B, 34, 776 (2017).
12. Э. Клингер, А. Саргсян, К. Леруа, Д. Саркисян. ЖЭТФ, 152, 641 (2017).
13. A. Sargsyan, A. Papoyan, I.G. Hughes, Ch. S. Adams, D. Sarkisyan. Optics Letters, 42, 1476 (2017).
14. K.A. Whittaker, J. Keaveney, I.G. Hughes, A. Sargsyan, D. Sarkisyan, C.S. Adams. Phys. Rev. Lett., 112, 253201 (2014).
15. K.A. Whittaker, J. Keaveney, I.G. Hughes, A. Sargsyan, D. Sarkisyan, C.S. Adams. Phys. Rev. A, 92, 052706 (2015).
16. А. Саргсян, А. Амирян, С. Карталева, Д. Саркисян. ЖЭТФ 152, 54 (2017).
17. J. Keaveney. Collective Atom Light Interactions in Dense Atomic Vapours. Springer, 2014.
18. M.A. Zentile, J. Keaveney, L. Weller, D.J. Whiting, C.S. Adams, I.G. Hughes. Comput. Phys. Commun., 189, 162 (2015).
19. B.A. Olsen, B. Patton, Y.Y. Jau, W. Happer. Phys. Rev. A, 84, 063410 (2011).
20. L. Weller, K.S. Kleinbach, M.A. Zentile, S. Knappe, C.S. Adams, I.G. Hughes. J. Phys. B, 45, 215005 (2012).
21. А.Д. Саргсян, Г.Т. Ахумян, А.О. Амирян, К. Леруа, А.С. Саркисян, Д.Г. Саркисян. Известия НАН Арм., Физика, 50, 428 (2015).
22. А. Саргсян, Б. Глушко, Д. Саркисян, ЖЭТФ 147, 668 (2015).
23. A. Sargsyan, A. Tonoyan, G. Hakumyan, C. Leroy, Y. Pashayan-Leroy, D. Sarkisyan. Opt. Comm. 334, 208 (2015).
24. A. Sargsyan, E. Klinger, A. Tonoyan, C. Leroy, D. Sarkisyan. J. Phys. B, 51, 145001 (2018).

1.1.2.14. Physico-mechanical and tribological properties of nanocomposites and their vulcanizates on the basis of molybdenum disulphide and ethylene-propylene block copolymer. /N. Kakhramanov, G. Gasimova, S. Pesetskiy, A. Quliyev, E. Dadasheva, A. Qasanova/. Azerbaijan Chemical Journal. – 2019. – #1. – pp. 39-45. – eng.; abs.:rus., az., eng.

The effect of the concentration and size of molybdenum disulphide fine particles on the breaking stress ultimate tensile stress, elongation at break, melt flow index and wear resistance of composite materials based on ethylene-propylene block copolymer and molybdenum disulphide is considered. The possibility of a significant improvement in the properties of nanocomposites by carrying out the crosslinking reaction in a dynamic mode is shown. Nanocomposites patterns of transition from solid to viscous state are investigated by thermomechanical analysis. The optimal concentrations of the reacting components have been established, at which the highest physicomechanical properties of nanocomposites are achieved. Fig. 3, Tab. 2, Ref. 10.

Keywords: block copolymer, molybdenum sulphide, thermomechanics, nanocomposites, breaking stress

References

- 1.Berlin A.A., Volfson S.A., Oshman V.G. Printcipy sozdaniia kompozitcionnykh materialov. M.: Himia, 1990. 240s.
2. Gorohovskii A.V. Kompozitnye materialy. Saratovskii gos. tekhnicheskii un., 2008. 73 s.
3. Simonov-Emelianov I.D., Kuleznev V.N. Osnovy sozdaniia kompozitcionnykh materialov. M.: MIKHM, 1986. 64 s.
4. Kakhramanov N.T., Kakhramanly Iu.N. Nanotekhnologii v oblasti materialovedeniia. "Uchenye Zapiski. Nacionalnaiia Aviac. Akademiiia Azerbaidzhana. 2009. T.11, № 3. S. 7-15.
5. Ivanchev S.S., Ozerin A.N. Nanostruktury v polimernykh sistemakh. //Vysokomolek. soed. 2006. B. T. 48. № 8. S. 1541-1544.

6. Kodolov V.I., Khokhriakov N.V. Kuznetsov A.P. Perspectives of nanostructure and nanosystems application when creating composites with predicted behavior. In the book: Space challenges in XXI century. V. 3. Novel materials and technologies for space rockets and space development. M.: Torus press. 2007. P. 201-205.
7. Pomogailo A.D. Gibridnye polimer-neorganicheskie nanokompozity // Uspehi himii. 2000. T. 69. № 1. S. 60.
8. Kodolov V.I., Hokhriakov N.V., Kuznetcov A.P. K voprosu o mehanizme vliianiia nanostruktur na strukturno izmeniaiushchiesia s redy pri formirovani "intellektualnykh" kompozitov // Nanotekhnika. 2006. T. 7. № 3. S. 27-35.
9. Okhlopkova T.A., Okhlopkova A.A., Spiridonov A.M., Nikiforov L.A. Ctrukturnye izmeneniiia sverkhvysokomo-tekuliarnogo polietilena pod vozdeistviem keramicheskikh nanodispersii // Voprosy materialovedeniia. 2014. T. 79. № 3. S. 145-153.
10. Okhlopkova A.A., Okhlopkova T.A., Borisova R.V. Upravlenie protcessami strukturoobrazovaniia v polimernykh kompozitcionnykh materialakh na osnove SVMPE // Nauka i obrazovanie. 2015. T. 78. № 2. S. 73-78.

1.1.2.15. Influence of fullerenes on dielectric and conductivity properties of smectic a liquid crystal with negative dielectric anisotropy. /T. Ibragimov, A. Imamaliyev, G. Ganizade/. Azerbaijan Journal of Physics. – 2019. – v. 25. – #1. – pp. 31-34. – eng.; abs.: eng.

Influence of fullerenes on dielectric and conductivity properties of smectic A liquid crystal is investigated. It is shown that the transverse component of the real part of dielectric permittivity increases at the additive of fullerenes while the longitudinal component decreases. At this case, a maximum of dielectric absorption shifts to the high-frequency region. Conductivity increases at low frequencies and decreases at the high ones. Experimental results are explained on the base of the Maier - Meier theory for anisotropic liquids and also participation of fullerenes in ionic conductivity of liquid crystal. Fig. 4, Ref. 13.

Keywords: liquid crystal, smectic A, fullerenes, dielectric permittivity

References

1. T.D. Ibragimov, G.M. Bayramov. Influence of small particles on Carr-Helfrich electrohydrodynamic instability in the liquid crystal. Intern. Journal for Light and Electron Optics, 2013, v. 124, pp. 3004-3006.
2. T.D. Ibragimov, G.M. Bayramov, A.R. Imamaliyev Clark-Lagerwall effect in the small particles-ferroelectric liquid crystal system. International Journal for Light and Electron Optics, 2013, v. 124, pp. 343-346.
3. T.D. Ibragimov, A.R. Imamaliyev, G.M. Bayramov Influence of barium titanate particles on electrooptic characteristics of liquid crystalline mixture H37. Intern. Journal for Light and Electron Optics, 2016, v. 127, pp. 1217-1220.
4. T.D. Ibragimov, A.R. Imamaliyev, G.M. Bayramov Electro-optic properties of the Ba TiO₃-Liquid crystal 5CB colloid. International Journal for Light and Electron Optics, 2016, v.127, pp.2278-2281.
5. T.D. Ibragimov, A.R. Imamaliyev, G.M. Bayramov. Formation of local electric fields in the ferroelectric BaTiO₃ particles-liquid crystal colloids. Ferroelectrics, 2016, v. 495, No 1, pp.60-68.
6. T.D. Ibragimov, A.R. Imamaliyev, G.M. Bayramov. Peculiarities of electro-optic properties of the ferroelectric particles-liquid crystal colloids. AIP Conference Proceedings, 2016, v. 1727, pp.020011-1-8.
7. S.E. San, M. Okutan, O.K. Äöysal, Y. Yerli. Carbon Nanoparticles in Nematic Liquid Crystals. Chin. Phys. Lett., 2008, v. 25, pp. 212-215.
8. S.E. San, O. Köysal, M. Okutan. Laser-induced dielectric anisotropy of a hybrid liquid crystal composite made up of methyl red and fullerene C60. Journal of Non-Crystalline Solids, 2005, v.351, No 33-36, pp. 2798-2801.
9. M. Okutan, S. ESan, E. Basaran, F. Yakuphanoglu. Determination of phase transition from nematic to isotropic state in carbon nanoballs' doped nematic liquid crystals by electrical conductivity -dielectric measurements. Physics Lett. A, 2005, v. 339, No6, pp. 461-465.
10. Liquid Crystals beyond Displays. Chemistry, Physics, and Applications. Edited by Li Q. New Jersey, Hoboken: John Wiley and Sons, Inc. 2012, 573.
11. F. Haragushi, K. Inoue, N. Toshima et al. Reduction of the threshold voltages of nematic liquid crystal electrooptical devices by doping inorganic nanoparticles. Japanese Journal Applied Physics, 2007, v. 46, pp. L796-L797
12. L. Blinov. Structure and properties of liquid crystal. New York: Springer, 2011, 439.
13. L.M. Blinov, V.G. Chiginov. Electrooptic effects in liquid crystal materials. 1994, Springer Verlag, p.488.

1.1.2.16. Structural peculiarities and dielectric properties of high-density polyethylene films containing InP and Ge particles. /S. Rzayeva, Sh. Rashidova, T. Ismailov/. Azerbaijan Journal of Physics. –2019. – v. 25. – #1. – pp. 39-44. – eng.; abs.: eng.

High density polyethylene (HDPE) films of various thickness (100, 200 μm) and content (2-10 vol.%) of InP/Ge particles were obtained by hot pressing a homogeneous mixture of fillers with a polymer matrix under a pressure of 10 MPa and at a temperature $T=413\text{K}$. HDPE films with dispersed InP and Ge particles were characterized by a combination of X-ray fluorescence microscopy (XRFM), X-ray diffractometry (XRD) and infrared Fourier transform spectroscopy (FT-IR) methods. It is shown that the obtained samples are HDPE-based composites with nanostructured InP and Ge-particles with a size of coherent scattering regions (CSR) of 44.1 and 55.7 nm, respectively. By scanning the surface of the films with a spot of a beam with a diameter of 10 μm , local concentrations of InP and Ge particles in the samples were determined. The dependence of the characteristic frequencies of FT-IR for HDPE at 719 and 730 cm^{-1} on the concentration of the filler was established. The temperature (in the range of 290–440K) and frequency (range 25– 10^6 Hz) dependences of the dielectric constant and the angle of the tangent of dielectric losses were measured. The dependence of the values of dielectric characteristics on the content of the filler in the composite is established. Fig. 5, Tab. 2, Ref. 15.

Keywords: InP, Ge, HDPE, X-ray fluorescence microscopy, X-ray diffractometry, infrared Fourier spectroscopy, dielectric properties

Reference:

1. R. Wang, Y.T. Zhang, J.Q. Yao. IEEE Photonics journal, 7, 2015: 2600308
2. M. Pattabi, A.B. Saraswathi. Composite Interfaces, 17, 2010: 103-111.
3. N.P. Gaponik, D.V. Talapin, A.L. Rogach, A. Eychmüller. Journal of Materials Chemistry, 10, 2000: 2163-2166.
4. F.Habelhams, B. Nessark, D. Bounhafs, A. Cheriet, H. Derbal. Synthesis and characterisation of polypyrrole - indium phosphide composite film. Ionics, 2010: 177-184.
5. B. Scruggs, M. Haschke, L. Herczeg, J. Nicolosi. Advances in X-ray Analysis, 42, 2000: 19-25.
6. K. Sugihara, M. Satoh, Y. Hayakawa, A. Saito, T. Sasaki. Advances in X-ray Analysis, 42, 2000: 161- 170.
7. Sh.Sh. Rashidova, M.A. Gusinova, S.M. Rzayeva. Journal of Engineering Physics and Thermophysics, Vol 88, 2015: 781-784.
8. Chung, T.C.M. «Functional polyolefins for energy applications», Macromolecules, 2013, vol. 46, p. 6671 - 6698.
9. Scheidl, K. Global PE Report, Proc. of PE, 99 Polyethylene World Congr., Zurich, September, 1999. P. 1-2.
10. Polyethylene of high pressure. Scientific-technical basis of industrial synthesis. A.V. Polyakov, F.I. Duntov, A.E. Sofiyev et al. L: Chemistry, 1988, p.200 (in Russian)
11. T. Hirtsu and P. Nugroho. Polymeric structure. Journal of Applied Polymer Science, 1997, 66, 6, 1049.
12. A.R Blythe., D. Bloor. Electrical properties of polymers. M.: fizmatlit, 2008. 376 p. (in Russian).
13. R. Kotek Recent advances in polymer fiber. Polymer Reviews Vol. 48, issue 2, April–June, 2008, p. 221-229
14. R. Zbinden. IR of spectroscopy of high polymers. M.: Mir, 1966, p.356 (in Russian).
15. M.I. Aliev, E.M. Gojaev, S.M. Rzaeva and Sh.Sh. Rashidova. Physical Science International Journal 8 (3): XX -XX, 2015, Article no. PSIJ. 18839. p. 1-5.

1.1.2.17. Photoluminescence properties of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ nanopowders. /Sh. Ahmadova, T. Naghiyev, Sh. Aliyeva, A. Sadigova, T. Mehdiyev/. Azerbaijan Journal of Physics. –2019. – v. 25. – #2. – pp. 25-30. – eng.; abs.: eng.

The photoluminescence spectra of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ ferrite nanopowders with different Zn contents were studied. The experiments were carried out at 300K, spectral lines were used to excitation luminescence: Xe-lamp with the wavelength 280 nm, 290 nm, 300 nm, 325 nm, 350 nm, 375 nm, 388 nm, 400 nm, 425 nm, and also YAG Nd laser (532 nm). The obtained spectra were interpreted in the framework of the proposed model in [1] for Fe_3O_4 , a structural analogue of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ ferrites. Fig. 4, Tab. 1, Ref. 15.

Keywords: ferrites, photoluminescence, nanopowders, sublattice

References

1. M.E.Sadat, Masoud Kaveh Baghbador, Andrew W. Dunn, H.P. Wagner, Rodney C. Ewing, Jiaming Zhang, Hong Xu, Giovann, M. Pauletti, David B. Mast, and Donglu Shi. Photoluminescence and photothermal effect of Fe₃O₄ nanoparticles for medical imaging and therapy, *Applied Physics Letter*, 105, 091903-1-091903-5, 2014.
2. E.W. Gorter. *Philips Res. Rep.*, 9, 4, 295, 5, 321, 6, 403, 1954.
3. M.Sorescu, L. Diamandescu, R. Peelamedu, R. Roy, P. Yadoji. Structural and magnetic properties of NiZn ferrites prepared by microwave sintering, *Journal of Magnetism and Magnetic Materials*, 279 (2004), 195–201, doi:10.1016/j.jmmm.01.079, 2004.
4. V. Blanco-Gutiérrez, M.J. Torralvo, Sáez-PucheR. and P.Bonville. Magnetic properties of solvothermally synthesized ZnFe₂O₄ nanoparticles *Journal of Physics: Conference Series* 200,072013, 5 p., 2010.
5. V.G. Kostishyn, B.K. Ostafiychuk, V.V. Moklyak, A.V. Nuriev. *Materials of Electronics Engineering*, 2013, 4, 22-29.
6. A.A.Sadigova, S.A.Ahmadova, Sh.N.Aliyeva, T.R.Mehdiyev. IR Diffuse Reflectance Spectra of Nanopowders of Ni_{1-x}ZnxFe₂O₄ Ferrites, *AJP Fizika* XXIV, p .26, 2018
7. N.N. Scholtz, K.A. Piskarev. *Ferrimagnetic materials for radio-frequencies*, Publishing House Energy, Moscow, 2013.
8. S. Aliyeva, S. Babayev, T. Mehdiyev. Raman spectra of Ni_{1-x}ZnxFe₂O₄ nanopowders, *JRS*; 49 (2), 271, 2018.
9. Y.V. Kasyuk, L.A. Bliznyuk, N.A. Basov, A.K. Fedotov, A.S. Fedotov, I.A. Svito. Structure and electrophysical properties of doped zinc oxide ceramics, International Scientific Conference, November 22-25, Minsk, Vol. 2, pp.84-86, 2016.
10. Michael Hoppe, Sven Doring, Mihaela Gorgoi, Stefan Cramm and Martina Muller. Enhanced ferrimagnetism in auxetic NiFe₂O₄ in the crossover to the ultrathin film limit, *Phys. Rev. B* 91, 054418 – Published 24 February 2015.
11. J. B. Goodenough. *Magnetism and the chemical bond*. (Interscience Publishers, New York, 1963).
12. J. Kanamori. "Superexchange interaction and symmetry properties of electron orbitals", *Journal of Physics and Chemistry of Solids* 10, 87, doi: 10.1016/0022-3697(59)90061-7, 1959.
13. P.W. Anderson. "Chapter 2 - Exchange in Insulators", in G. T. Rado and H. Suhl, *Magnetism*. Academic Press. pp. 25-83, isbn:978-0-12-575301-2, 1963.
14. F. Shahbaz Tehrani, V. Daadmehr, A.T. Rezakhani, R. Hosseini Akbarnejad, S. Gholipour. Structural, magnetic, and optical properties of zinc- and copper- substituted nickel ferrite nano-crystals, *Journal of Superconductivity and Novel Magnetism*, October 2012, Vol. 25, Iss. 7, pp. 2443-2455.
15. I.F. Yusibova, Sh.N. Aliyeva, T.R. Mehdiyev. Verwey transition in Raman scattering spectra of (Ni-Zn) ferrite nanofilms, *AJP Fizika*, Vol. XXIV, №3, Section: Az, 2018, 95-100.

1.1.2.18. Dielectric relaxation in the colloid fullerenes - liquid crystal 5CB. /G. Ganizade/. Azerbaijan Journal of Physics. – 2019. – v. 25. – #3. – pp. 3-6. – eng.; abs.: eng.

Influence of fullerenes C₆₀ on dielectric properties of liquid crystal 5CB (4'-Pentyl-4-biphenylcarbonitrile) is investigated. It is shown that at additive of fullerenes the components of the dielectric permittivity decrease, and the longitudinal component decreases by greater extent than its transverse component. The dielectric anisotropy also decreases. The dielectric absorption is shifted to low frequencies. The experimental results are explained by the Maier - Meier theory for anisotropic liquids. Fig. 4, Ref. 7.

Keywords: Liquid crystal; fullerenes; dielectric permittivity; phase transition

References

1. N.V. Kamanina. Fullerene-dispersed nematic liquid crystal structures: dynamic characteristics and self-organization processes, *Physics- Uspekhi* 48, 419-427, 2005 (in Russian).
2. S.E. San, M. Okutan, O. Koysal, Y. Yerli. Carbon Nanoparticles in Nematic Liquid Crystals Chin. Phys. Lett. 25, 212-215, 2008.
3. S.E. San, O. Koysal, M. Okutan. Laser-induced dielectric anisotropy of a hybrid liquid crystal composite made up of methyl red and fullerene C₆₀, *J. of Non-Crystalline Solids*, 351, 2798-2801, 2005.
4. M. Okutan, S.E. San, E. Basaran, F. Yakuphanoglu. Determination of phase transition from nematic to isotropic state in carbon nano-balls' doped nematic liquid crystals by electrical conductivity-dielectric measurements, *Phys. Lett. A*, 339, 461-465, 2005.
5. Yu. Reznikov. Ferroelectric colloids in liquid crystals, in: Q. Li (Ed.), *Liquid Crystals Beyond Displays: Chemistry, Physics, and Applications*, John Wiley and Sons, Inc., Hoboken, NJ, p.p. 403-426, 2012.
6. M.V. Gorkunov, M.A. Osipov. Mean field theory of a nematic liquid crystal doped with anisotropic nanoparticles, *Soft Matter*, 7, 4348-4356, 2011.

7. W. Maier, G. Meier. A simple theory of the dielectric are some homogeneous criteria oriented liquid crystal phases of nematic type. Z. Naturforsch. 16A, 262–267, 1961.

1.1.2.19. Optical and UV-VIS luminescence spectra of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ ferrite nanopowders. /A. Sadigova, Sh. Aliyeva, Sh. Ahmadova, I. Yusibova, T. Naghiyev, T. Mehdiyev/. Azerbaijan Journal of Physics. – 2019. – v. 25. – #3. – pp. 34-43. – eng.; abs.: eng.

The optical and UV-VIS luminescent spectra of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ ferrite nanopowders with $x = 0; 0,25; 0,4; 0,5; 0,6; 0,75; 1,0$ were investigated in $4000-50 \text{ cm}^{-1}$ and 200-700 nm at room temperature. The features of the diffuse reflectance spectra of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ ferrites were analyzed by the Kramers-Kronig procedure. The agreement with the data of published studies of other authors allowed us to give a hypothetical interpretation of the results. Fig. 5, Tab. 2, Ref. 43.

Keywords: ferrites, nanopowders, IR spectra, hypothetical interpretation, Kramer-Kronig relations

References

- 1.S. Aliyeva, S. Babayev, T. Mehdiyev. Raman spectra of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ nanopowders, JRS 2018; 49 (2), 271.
2. Sh.N. Aliyeva, A.M.Kerimova, R.B. Abdullayev, T.R. Mehdiyev. IR spectra of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ ferrite micropowders, PhSS, 2017, v. 59, no. 3, pp. 528-533.
3. Sh.N. Aliyeva. Magnetic properties of $\text{Ni}_{1-x}\text{Zn}_x\text{Fe}_2\text{O}_4$ micropowders and thin films, Baku, 2017, p. 199.
4. Sh.N. Aliyeva, Y.N. Aliyeva, A.I. Nadjafov, I.S. Hasanov, E.K. Huseynov T.R. Mehdiyev. EPR and SPM studies of Zn-Ni ferrites, Phys. Status Sol. (c), 615, 2015/DOI 10.1002/pssc.201400273.
5. N.N. Scholtz, K.A. Piskarev. Ferrimagnetic materials for radio-frequencies, Publishing House Energy, Moscow, 2013.
6. U.V. Kasyuk, L.A. Bliznyuk, N.A. Basov, A.K.Fedotov, A.S. Fedotov, I.A. Svito .Structure and electro-physical properties of doped ceramics on the base of zinc oxide. The theses of International conference, Minsk, November 22 – 25, 2016, vol. 2, pp. 84-86.
7. J. Nishitani, K. Kozuki, T. Nagashima, M. Hangyo. Appl.Phys. Lett.2010,96,221906-1.
8. F.Sh. Tehrani, V. Daadmehr, A.T. Rezakhani, R.H. Akbarnejad, S. Gholipour. J. Supercond. Novel Magnetism, 2012, 25, 2443.
9. R.C. Sriprya, Ezhil Arasi S. Madhavan. J. Victor Antony Raj M. Synthesis and Characterization studies of ZnFe_2O_4 nanoparticles, MMSE Journal vol.9, iss.1, 2017, p.13-18.
10. K. Kombaiah, J. Judith Vijaya*, L. John Kennedy, K. Kaviyarasu. Catalytic studies of NiFe_2O_4 nanoparticles prepared by conventional and microwave combustion method, Materials Chemistry and Physics 221, 2019, 11-28.
11. Jamil K. Salem, Talaat M. Hammad, Roger R. Harrison. Synthesis, structural and optical properties of Ni-doped ZnO microspheres, J Mater Sci: Mater Electron, 2012, DOI 10.1007/s10854-012-0994-0.
12. P.A. Sheena, K.P. Priyanka, N. Aloysius Sabu, Boby Sabu, Thomas Varghese. Effect of calcination temperature on the structural and optical properties of nickel oxide nanoparticles, Nanosystems: Physics, Chemistry, Mathematics, 2014, 5 (3), p.441-449.
13. M.E. Sadat, Masoud Kaveh Baghbador, Andrew W. Dunn, H.P. Wagner, Rodney C. Ewing, Jiaming Zhang, Hong Xu, Giovanni M. Pauletti, David B. Mast and Donglu Shi. Photoluminescence and photothermal effect of Fe_3O_4 nanoparticles for medical imaging and therapy, Applied Physics Letter, 105, 0919031-091903-5, 2014.
14. Walmir E. Pottker, Rodrigo Ono, Miguel Angel Cobos, Antonio Hernando, Jefferson F.D.F. Araujo, Antonio C.O. Bruno, Sidney A. Lourenço, Elson Longo, Felipe A. La Porta. Influence of order-disorder effects on the magnetic and optical properties of NiFe_2O_4 nanoparticles, 44, 2018, 17290-17297.
15. S. Lakshmi Reddy, Tamio Endo and G. Siva Reddy. Electronic (Absorption) Spectra of 3d Transition Metal Complexes, in book IntechOpen: Advanced Aspects of Spectroscopy, chapter 1, 2012.
16. Veronica D'Ippolito, Giovanni Battista Andreozzi, Ulf Hålenius, Henrik Skogby, Kathrin Hametner, Detlef Günther. Color mechanisms in spinel: cobalt and iron interplay for the blue color, Phys Chem Minerals, 2015, 42:431-439.
17. H.K. Mao and P.M. Bell. Crystal-field effects in spinel: oxidation states of iron and chromium, Geochimica et Cosmochimica Acta, 1975, vol. 39, pp. 869 to 871.
18. F. Shahbaz Tehrani, V. Daadmehr, A.T. Rezakhani, R. Hosseini Akbarnejad, S. Gholipour. Structural, magnetic, and optical properties of zinc- and copper-substituted nickel ferrite nanocrystalsJournal of Superconductivity and Novel Magnetism, 2012, vol. 25, issue 7, pp 2443-2455, DOI:10.1007/s10948-012-1655-5.
19. David M. Sherman and T. David Waite. Electronic spectra of Fe^{3+} oxides and oxide hydroxides in the near IR to near UV, American Mineralogist, vol. 70, p. 1262-1269, 1985.

20. The Infrared Spectra of Minerals, 539 pp., Mineral.Soc., London, 1974.
21. *J.T. Keiser, C.W. Brown, R.H. Heidersbach.* Infrared spectra of magnetite nanoparticles, J. Electrochim. Soc. 1982, vol. 129, p. 2686.
22. *Ma M., Zhang Yu., Wei Yu. et. al.* Preparation and characterization of magnetite nanoparticles coated by amino silane, Colloids and Surfaces: Physicochem. Eng. Aspects. 2003, vol.212, p.219-226.
23. *F. Shahbaz Tehrani, V. Daadmehr, A.T. Rezakhani, R. Hosseini Akbarnejad, S. Gholipour.* Structural, magnetic, and optical properties of zinc- and copper- substituted nickel ferrite nanocrystals Journal of Superconductivity and Novel Magnetism, 2012, vol. 25, issue 7, pp 2443-2455, DOI:10.1007/s10948-012-1655-5.
24. *Jiaqi Wan, Xuehui Jiang, Hui Li and Kezheng Chen.* Facile synthesis of zinc ferrite nanoparticles as non-lanthanide T1 MRI contrast agents, Journal of Material Chemistry, 2012, 22, 13500-13505.
25. *Zahra Rezay Marand, Mitra Helmi Rashid Farimani, Nasser Shahtahmasebi.* Study of magnetic and structural and optical properties of Zn doped Fe₃O₄ nanoparticles synthesized by co-precipitation method for biomedical application, Nanomedicine Journal, v. 1, no. 4, 2014, p. 238-247.
26. *Santosh S. Jadhav, Sagar E. Shirsath, B.G. Toksha, S. J. Shukla, K. M. Jadhav.* Effect of Cation Proportion on the Structural and Magnetic Properties of Ni-Zn Ferrites Nano- Size Particles Prepared By Co-Precipitation Technique, Chinese Journal of Chemical Physics, vol. 21, N4, 2008, p.381-386.
27. *F. Shahbaz Tehrani, V. Daadmehr, A.T. Rezakhani, R. Hosseini Akbarnejad, S. Gholipour.* Structural, magnetic, and optical properties of zinc- and copper- substituted nickel ferrite nanocrystals Journal of Superconductivity and Novel Magnetism, 2012, vol. 25, issue 7, pp 2443-2455, DOI:10.1007/s10948-012-1655-5.
28. *K.P.Belov.* Ferrimagnets with "weak" magnetic sublattice, UFN, 1996, June, vol. 166, no. 6, pp. 669-681.
29. *C.R. Becker, Ph. LAU, R. Geick and V. Wagner.* Antiferromagnetic Resonance in NiO:Co²⁺ and NiO:Fe²⁺, Phys. Stat. Sol. (b) 67, 653-663 1975.
30. *Vidales J.L. M., A.L. Delgado, E. Vila and F.A. Lopez.* The effect of the starting solution on the physico-chemical properties of zinc ferrite synthesized at low temperature, J. Alloys Comp., 1999, 287, 276.
31. *Junichi Nishitani, Kohji Kozuki, Takeshi Nagashima and Masanori Hangyo.* Terahertz radiation from coherent antiferromagnetic magnons excited by femtosecond laser pulses, Appl. Phys. Lett., 96, p. 221906-1-221906-3, 2010.
32. *Shin G. Chou, Paul E. Stutzman, Shuangzhen Wang, Edward J. Garboczi, William F. Egelhoff, and David F. Plusquellic.* High- Resolution Terahertz Optical Absorption Study of the Antiferromagnetic Resonance Transition in Hematite (α -Fe₂O₃), J. Phys. Chem. C, 2012, 116 (30), p.16161-16166.
33. *A.N. Kislov, V.G. Mazurenko, A.N. Varaksin.* Analysis of vibronic structure of optical spectra in ZnO:Ni⁺³ crystals on the base of localized oscillation modelling, Phys. Stat. Sol, 1999, vol. 4, pp. 618-622.
34. *A.B.P. Lever.* Inorganic Electronic Spectroscopy. Amsterdam: Elsevier, 1968; *D.M. Sherman, T.D. Waite.* Electronic spectra of Fe³⁺ oxides and oxide hydroxides in the near IR to near UV, Amer. Mineralogist. 1985, 70, p. 1262.
35. *A.V. Dijken, E.A. Meulenkamp, D. Vanmaelbergh, Meijerink.* The Kinetics of the radiative and nonradiative processes in nanocrystalline ZnO particles upon photoexcitation, J. Phys. Chem. B 104, 2000, 1715-1723.
36. *L.Jing, Y. Qu, B. Wang, S. Li, B. Jiang, L.Yang, W.Fu, H.Fu.* Review of photoluminescence performance of nano-sized semiconductor materials and its relationships with photocatalytic activity, J. Sun, Sol. Energy Mater. Sol. Cells 90, 2006, 1773-1787.
37. *R.K. Sendi, S. Mahmudm,* Quantum size effect on ZnO nanoparticle-based discs synthesized by mechanical milling, Appl. Surf Sci. 258, 2012, 8026-8031.
38. *J. Becker, K.R. Raghupathi, J. St Pierre, D. Zhao, R.T. Koodal.* Tuning of the crystallite and particle sizes of ZnO nanocrystalline materials in solvothermal synthesis and their photocatalytic activity for dye degradation, J. Phys. Chem. C 115, 2011, 13844-13850
39. *D.I. Khomskii.* Transition metal oxides. Cambridge University Press, Cambridge, 2014. ISBN: 978-1-107-02017-7.
40. *Burns, Roger G. Mineralogical. Applications of Crystal Field Theory, ISBN 10:0521076102/ISBN13: 9780521076104,* Published by Cambridge University Press, Cambridge, 1970.
41. *E. Pavarini, E.Koch, F.Anders.* Correlated Electrons: From Models to Materials, Crystal-field Theory, Tight-binding Method, and Jahn-Teller Effect Julich 2012, ISBN 978-3-89336- 796-2.
42. *D.I. Khomskii.* Transition metal oxides. Cambridge University Press, Cambridge, 2014. ISBN: 978-1-107-02017-7.
43. *I.B. Bersuker.* The Jahn-Teller Effect. Cambridge University Press, Cambridge, 2006. ISBN 139780521822121.

2. NANO CHEMISTRY

2.1. Inorganic Materials

1.2.1.1. Nanochemical processes under conditions of geological repository: 1. Evolution of bentonite buffer. /O. Lavrynenko, B. Shabalin, O. Pavlenko/. Nano Studies. – 2019. – #19. – pp. 223-238. – eng.; abs.: eng.

In the review, an analysis of the bentonite buffer evolution under the geological repository conditions of a high-level waste and spent nuclear fuel was carried out. The analysis of the results of modern laboratory, field and model studies of the evolutionary changes in clay affecting the stability of the buffer and the safety of the storage has been carried out. A number of aspects of the bentonite transformation are highlighted: low-temperature changes taking into account potential cementing, colloid-chemical changes during the bentonite interaction with water, the effect of swelling pressure is shown. The analysis of the processes at the phase-interfaces cement-bentonite and iron(steel)-bentonite was carried out. The effect of mechanical shifts/displacements of the host rocks and erosion of bentonite on the mass redistribution under storage conditions was shown there. Ref. 77.

Keywords: nanochemical processes, spent nuclear fuel

References

1. Design and production of the KBS–3 repository (Technical Report TR–10–12), 2010, SKB.
2. В.М. Шестopalов, Ю.Ф. Руденко, Э.В. Соботович. Изоляция радиоактивных отходов в недрах (Проблемы и возможные решения), 2006, Киев, НАНУ.
3. Н.П. Лаверов, Б.И. Омельяненко, С.В. Юдинцев. Изоляционные свойства бентонитового буфера в условиях подземного хранилища высокоактивных отходов. Геология рудных месторождений, 2004, 46, 1, 27–42.
4. Geological Disposal. A review of the development of bentonite barriers in the KBS–3V disposal concept. NDA Technical Note No. 21665941, 2014.
5. H. Laine. Long-Term Stability of Bentonite: A Literature Review (Working Report 2010-53), 2010, Olkiluoto, Posiva OY.
6. D. Savage. Prospects for coupled modeling. STUK–TR 13, 2012, Helsinki.
7. J. McMurry, D.A. Dixon, J.D. Garroni, B.M. Ikeda, S. Stroes-Gascoyne, P. Baumgartner, T.W. Melnyk. Evolution of a Canadian deep geologic repository: Base scenario, Ontario Power Generation Nuclear Waste Management Division Report No 06819–REP–01200–10092–R00, 2003.
8. R. Pusch, O. Karnland. Hydrothermal effect on montmorillonite. A preliminary study. SKB Report TR–88–15, 1988, Swedish Nuclear Fuel and Waste Management Company, Stockholm.
9. R. Pusch, H. Takase, S. Benbow. Chemical processes causing cementation in heat affected smectite – the Kinnekulle bentonite. SKB TR–98–25. 1998, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
10. J. Rutqvist, J. Noorishad, C.F. Tsang. Coupled thermohydromechanical analysis of a heater test in unsaturated clay and fractured rock at Kamaishi Mine. SKI Report 99:50, 1999, Stockholm, Swedish Nuclear Inspectorate (SKI).
11. J.T. Birkholzer, J. Rutqvist, E.L. Sonnenthal, D. Barr. DECOVALEX–THMC Task D: Long-Term Permeability / Porosity Changes in the EDZ and Near Field due to THM and THC Processes in Volcanic and Crystalline-Bentonite Systems (LBNL Report 59122), 2005, Berkeley, Lawrence Berkeley National Laboratory.
12. A. Bond, C. Watson. The use of QPAC–EBS for Project THERESA benchmarking studies (Quintessa Report QRS–3009A–1), 2008, Henley-on-Thames, Quintessa Limited.
13. L. Zheng, J. Samper. A coupled THMC model of FEBEX mock-up test. Phys. & Chem. Earth, 2008, 33, S486–S498.
14. L. Zheng, J. Samper, L. Montenegro, A.M. Fernandez. A coupled THMC model of heating and hydration laboratory experiment in unsaturated compacted FEBEX bentonite. J. Hydrology, 2010, 386, 80–94.
15. A. Hedin. Integrated near-field evolution model for a KBS–3 repository (SKB Technical Report TR–06–11), 2006, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
16. O. Karnland, M. Birgersson. Montmorillonite stability with special respect to KBS–3 conditions (SKB Technical Report TR–06–11), 2006, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
17. A. Hedin. Integrated near-field evolution model for a KBS–3 repository (SKB Technical Report R–04–36), 2004, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
18. D. Arcos, F. Grandia, C. Domenech. Geochemical evolution of the near field of a KBS–3 repository (SKB Technical Report TR–06–16), 2006, Stockholm, Swedish Nuclear Fuel and Waste Management Company.

19. M. Birgersson, O. Karnland. Ion equilibrium between montmorillonite interlayer space and an external solution. *Geochimica et Cosmochimica Acta*, 2009, 73, 1908-1923.
20. M. Ochs, C. Talerico, P. Sellin, A. Hedin. MX-80 bentonite. *Phys. & Chem. Earth*, 2006, 31, 600-609.
21. I. Neretnieks, L. Liu, L. Moreno. SKB Report TR-09-35, 2010, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
22. M.H. Bradbury, B. Baeyens. Porewater chemistry in compacted re-saturated MX-80 bentonite. *J. Contaminant Hydrology*, 2003, 61, 329-338.
23. A. Muurinen, T. Carlsson. Development of methods for on-line measurements of chemical conditions in compacted bentonite. *Phys. & Chem. Earth*, 2007, 32, 241-246.
24. M.A. Glaus, B. Baeyens, M.H. Bradbury, A. Jakob, L.R. van Loon, A. Yaroshchuck. Diffusion of ^{22}Na and ^{85}Sr in montmorillonite: Evidence of interlayer diffusion being the dominant pathway at high compaction. *Environ. Sci. & Technol.*, 2007, 41, 478-485.
25. O. Karnland, A. Muurinen, F. Karlsson. Bentonite pressure in NaCl solutions – experimentally determined data and model calculations. In: Symp. Large-Scale Field Tests in Granite, 2002, Sitges.
26. O. Karnland, A. Muurinen, F. Karlsson. Bentonite swelling pressure in NaCl solutions. In: Advances in Understanding Engineered Clay Barriers, 2005, London, Taylor & Francis Group.
27. M. Birgersson, O. Karnland, U. Nilsson. Freezing in saturated bentonite – A thermodynamic approach. *Phys. & Chem. Earth*, 2008, 33, S527- S530.
28. L. Duro, M. Grive, E. Cera, X. Gaona, C. Dominech. Bibliographer (SKB Technical Report TR-06-2), 2006, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
29. I. Neretnieks, L. Liu, L. Moreno. SKB Report TR-09-35, 2010, Stockholm, Swedish Nuclear Fuel Company.
30. Spent nuclear fuel spent on nuclear fuel at Forsmark. Main Report of the SR-Site Project, Volumes I – III (SKB Report TR-11-01), 2011, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
31. M.H. Bradbury, B. Baeyens. Porewater chemistry in compacted re-saturated MX-80 bentonite. *J. Contaminant Hydrology*, 2003, 61, 329-338.
32. D. Savage, R. Arthur, C. Watson, J. Wilson. SSM Technical Report 2010–12, 2010, Stockholm, Swedish Radiation Safety Authority.
33. D. Savage, R. Arthur, C. Watson, J. Wilson, B. Strumberg. Testing geological analysis of experimental data. *Phys. & Chem. Earth*, 2019. –in press.
34. D. Savage, S. Benbow, C. Watson, H. Takase, K. Ono, C. Oda, A. Honda. Natural systems evidence for the alteration of clay under alkaline conditions: An example from Searles Lake, California. *Appl. Clay Sci.*, 2010, 47, 72-81.
35. D. Savage. The effects of high salinity groundwater on the performance of clay barriers, SKB Report 2005:54, 2005, Stockholm, Swedish Nuclear Power Inspectorate.
36. S.S. Agus, T. Schanz. A method for predicting swelling pressure of compacted bentonites. *Acta Geotechnica*, 2008, 3, 125-137.
37. R. Arthur. Handling of hydrogeochemical relations in erosion and swelling pressure models for the buffer and backfill (STUK Technical Report STUK-TR 10), 2011, Helsinki, Finnish Radiation and Nuclear Safety Authority (STUK).
38. P.F. Low. The swelling of clay. II: Montmorillonites. *Soil Sci. Soc. Am. J.*, 1980, 44, 667-676.
39. P.F. Low. Structural component of the swelling pressure of clays. *Langmuir*, 1987, 3, 18-25.
40. R. Grauer. Bentonite as a Chemical Method (Nagra Technical Report 86-12E), 1986, Baden, Nagra.
41. G. Sposito. Thermodynamics of swelling clay–water systems. *Soil Sci.*, 1972, 114, 243-249.
42. A. Hedin. KBS-3 repository (SKB Research Report R-04-36), 2004, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
43. P. Gribs, L.H. Johnson, D. Suter, P.A. Smith, B. Pastina, M. Snellman. Safety assessment for a KBS-3H spent nuclear fuel repository at Olkiluoto: Process report (SKB Report R-08-36), 2008, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
44. C.I. Steefel, P. van Cappellen. A new kinetic approach to modeling water–rock interaction: The role of nucleation, precursors, and Ostwald ripening. *Geochimica et Cosmochimica Acta*, 1990, 54, 2657-2677.
45. B. Fritz, A. Clement, Y. Amal, C. Noguera. Simulation of the nucleation and growth of simple clay minerals in weathering processes: The NANOKIN code. *Geochimica et Cosmochimica Acta*, 2009, 73, 1340-1358.
46. D. Savage, C. Walker, R.C. Arthur, C.A. Rochelle, C. Oda, H. Takase. Alteration of bentonite by hyperalkalineuids: A review of the role of secondary minerals. *Phys. & Chem. Earth*, 2007, 32, 287-297.
47. P. Wersin, M. Birgersson, S. Olsson, O. Karnland, M. Snellman. Impact of corrosion-derived iron on the bentonite buffer within the KBS-3H disposal concept: The Olkiluoto site as case study (SKB Report R-08-34), 2008, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
48. S. Lantenois, B. Lanson, F. Muller, A. Bauer, M. Jullien, A. Planzon. Experimental study of smectite interaction with metal Fe at low temperature: 1. Smectite destabilization. *Clays & Clay Minerals*, 2005, 53, 597-612.

49. L. Carlson, O. Karnland, V.M. Oversby, A. Rance, N. Smart, M. Snellman, M. Vdhnden, L.O. Werme. Experimental studies of corroding iron and bentonite. *Phys. & Chem. Earth*, 2007, 32, 334-345.
50. D. Savage, C. Watson, S. Benbow, J. Wilson. Modeling iron–bentonite interactions. *Appl. Clay Sci.*, 2010, 47, 91-98.
51. D. Savage. Equilibrium activity diagrams relevant to smectite clay stability (Report), 1986, Baden, Nagra. – unpublished
52. R. Grauer. The chemical behavior of montmorillonite in a repository backfill: Selected aspects (Technical Report 88-24E), 1990, Wettingen, Nagra.
53. J. Wilson, G. Cressey, B. Cressey, J. Cuadros, K.V. Ragnarsdttir, D. Savage, M. Shibata. The effect of iron on montmorillonite stability. (II) Experimental investigation. *Geochimica et Cosmochimica Acta*, 2006, 70, 323-336.
54. T. Ishidera, K. Ueno, S. Kurosawa, T. Suyama. Investigation of montmorillonite alteration and form of iron corrosion products in compacted bentonite in contact with carbon steel for ten years. *Phys. & Chem. Earth*, 2008, 33, S269-S275.
55. R. Mosser-Ruck, M. Cathelineau, D. Guillaume, D. Charpentier, D. Rousset, O. Barres, N. Michau. Effects of temperature, pH, and iron / clay and liquid / clay ratios on experimental conversion of dioctahedral smectite to berthierine, chlorite, vermiculite, or saponite. *Clays & Clay Minerals*, 2010, 58, 280-291.
56. L. Carlson, O. Karnland, S. Olsson, A. Rance, N.R. Smart. Experimental studies on the interactions between anaerobically corroding iron and bentonite (SKB Report R-08-28), 2008, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
57. A.E. Milodowski, M.R. Cave, S.J. Kemp, H. Taylor, B. Vickers, K. Green, C.L. Williams, R.A. Shaw. Mineralogical investigations of the interaction between iron corrosion products and bentonite from the NF-PRO Experiments, Phase 1 (SKB Technical Report TR-09-02), 2009, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
58. O. Bildstein, L. Trotignon, M. Perronet, M. Jullien. Modelling iron–clay interactions in deep geological disposal. *Phys. & Chem. Earth*, 2006, 31, 618-625.
59. J. Samper, C. Lu, L. Montenegro. Reactive transport model of interactions of corrosion products and bentonite. *Phys. & Chem. Earth*, 2008, 33, S306-S316.
60. C. Lu, J. Samper, B. Fritz, A. Clement, L. Montenegro. Interactions of corrosion products and bentonite: An Extended multicomponent reactive transport model. *Physics and Chemistry of the Earth*, 2011.
61. H.G. Montes, N. Marty, B. Fritz, A. Clement, N. Michau. Modeling of long-term diffusion reaction in a bentonite barrier for radioactive waste containment. *Appl. Clay Sci.*, 2005, 30, 181-198.
62. F.M.I. Hunter, F. Bate, T.G. Heath, A.R. Hoch. Geochemical investigation of iron transport into bentonite as steel corrodes (SKB Technical Report TR-07-09), 2007, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
63. P.S. Wrestling, M.C. Snellman. Reporting R-08-45, 2008, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
64. A.E. Milodowski, M.R. Cave, S.J. Kemp, H. Taylor, K. Green, C.L. Williams, R.A. Shaw. Gowing, CJB, and NF-PRO Experiments, Phase 2 (SKB Report TR-09-03), 2009, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
65. M.E. Tucker. Sedimentary Petrology. An Introduction to the Sedimentary Rocks, 1991, Oxford, Blackwell Sci. Publ.
66. G.S. Odin. Clay mineral formation at the continent–ocean boundary: The Verdine Facies. *Clay Minerals*, 1990, 25, 477-483.
67. L. Truche, G. Berger, C. Destigneville, D. Guillaume, E. Giffaut. Redox reactions. *Geochimica et Cosmochimica Acta*, 2009, Goldschmidt Conference Abstracts: A1348.
68. Posiva Report TKS-2009. Nuclear Waste Management at Olkiluoto and Loviisa Power Plants, 2009, Olkiluoto, Posiva Oy.
69. L. Borgesson, J. Hernelind. Earthquake induced rocks shear through a deposition hole. Modeling of three model tests scaled 1:10 (SKB Report TR-10-33), 2010, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
70. J. Hernelind. Modeling and analysis of earthquake for canalizing external and earthquake (SKB Report TR-10-34), 2010, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
71. E.E. Alonso, A. Gens, A. Josa. A constitutive model for partially saturated soils. *Geotechnique*, 1990, 40, 405-430.
72. E.E. Alonso, J. Vaunat, A. Gens. Modeling the mechanical behavior of expansive clays. *Eng. Geology*, 1999, 54, 173-183.
73. B. Miller, N. Marcos. Process report—FEPs and scenarios for a spent fuel repository at Olkiluoto (Posiva Report 2007-12), 2007, Olkiluoto, Posiva Oy.
74. M. Birgersson, L. Borgesson, M. Hedström, O. Karnland, U. Nilsson. Bentonite erosion. Final Report (Technical Report TR-09-34), 2009, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
75. I. Neretnieks, L. Liu, L. Moreno. SKB Report TR-09-35, 2010, Stockholm, Swedish Nuclear Fuel and Waste Management Company.

76. L. Moreno, I. Neretnieks, L. Liu. Modeling of erosion (SKB Report TR-10-64), 2011, Stockholm, Swedish Nuclear Fuel and Waste Management Company.
77. L. Moreno, I. Neretnieks, L. Liu. SKB Report TR-10-64, 2011, Stockholm, Swedish Nuclear Fuel and Waste Management Company.

1.2.1.2. Nanochemical processes under conditions of geological repository: 2. Steel-bentonite interface. /O. Lavrynenko, B. Shabalov, O. Pavlenko/. Nano Studies. – 2019. – #19. – pp. 239-258. – eng.; abs.: eng.

In the second part of the review, a description of the processes proceeding on the steel-bentonite interface is given. The changes that the unsaturated groundwater undergoes a buffer upon contact with the surface of the steel container are considered. The results of modeling the corrosion process under conditions of geological storage are presented, and the contribution of the anionic component and microorganisms is highlighted. The main provisions that are taken into account in conducting geochemical modeling of corrosion processes in the conditions of geological storage of radioactive waste are presented. Ref. 85.

Keywords: steel-bentonite interface, unsaturated groundwater, buffer, geological repository, radioactive waste

References

1. K. Idemitsu, S. Yano, X. Xia, Y. Kikuchi, Y. Inagaki, T. Arima. Migration behavior of iron in compacted bentonite under reducing condition using electromigration. Mater. Res. Soc. Symp. Proc., 2003, 757, II3. 7.1-8.
2. G. Kamei, C. Oda, S. Mitsui, M. Shibata, T. Shinozaki. Fe(II)-Na ion exchange at interlayers of smectite: adsorption-desorption experiments and a natural analogue. Eng. Geology, 1999, 54, 15-20.
3. J.W. Stuki, P.F. Low, C.B. Roth, D.C. Golden. Effects of oxidation state on octahedral iron on clay swelling. Clays & Clay Minerals, 1984, 32, 357-362.
4. J.E. Kostka, J. Wu, K.H. Nealson, J. Stuki. The impact of structural Fe(III) reduction by bacteria on the surface chemistry of smectitic clay minerals. Geochimica et Cosmochimica Acta, 1999, 63, 3705-3713.
5. J. Wilson, D. Savage, J. Cuadros, M. Shibata, K.V. Ragnarsdottir. The effect of iron on montmorillonite stability. (I) Background and thermodynamic consideration. Geochimica et Cosmochimica Acta, 2006, 70, 306-322.
6. J. Wilson, G. Cressey, B. Cressey, J. Cuadros, K.V. Ragnarsdottir, D. Savage, M. Shibata. The effect of iron on montmorillonite stability. (II) Experimental investigations. Geochimica et Cosmochimica Acta, 2006, 70, 323-336.
7. S. Necib, N. Diomidis, P. Keech, M. Nakayama. Corrosion of carbon steel in clay environments relevant to radioactive waste geological disposals, Mont Terri rock laboratory (Switzerland). Swiss J. Geo Sci., 2017, 110, 329-342.
8. F.T. Madsen. Clay mineralogical investigations related to nuclear waste disposal. Clay Minerals, 1998, 33, 109-129.
9. A.E. Milodowski, M.R. Cave, S.J. Kemp, H. Taylor, B.P. Vickers, K. Green, C.L. Williams, R.A. Shaw. Mineralogical investigations of the interaction between iron corrosion products and bentonite from the NF-PRO Experiments, Phase 1. SKB TR-09-02, 2009, Svensk Karnbranslehantering AB.
10. P. Sellin, O.X. Leupin. The use of clay as an engineered barrier in radioactive-waste management – A review. Clays & Clay Minerals, 2013, 61, 477-498.
11. F. Arbaoui, S.A. Amzert, M.C.N. Boucherit. Corrosion of aqueous solution. J. Fund. & Appl Sci., 2017, 9, 1300-1319.
12. D. Svensson, S. Hansen. Environmental Recycling. Clays & Clay Minerals, 2013, 61, 566-579.
13. N.R. Smart, L. Carlson, F.M.I. Hunter, O. Karland, A.M. Pritchard, A.P. Rance, L.O. Werme. Interactions between iron products and bentonite. Serco Assurance Report SA/EIG/12156/C001, 2006.
14. M.L. Schlegel, Ch. Bataillon, K. Benhamida, C. Blanc, D. Menut, J.L. Lacour. Metal corrosion and argillite transformation at the water-saturated, high-temperature iron-clay interface: A microscopic-scale study. Appl. Geo Chem., 2008, 23, 2619-2633.
15. S. Kaufhold, A.W. Hassel, D. Sanders, R. Dohrmann. Corrosion of high-level radioactive waste ironing canisters. J. Hazard Mater., 2015, 285, 464-473.
16. E. Torres, M.J. Turrero, P.L. Martin. Geochemical Processes at the carbon steel / bentonite interface in repository conditions. Mater. Res. Soc. Symp. Proc., 2007, 985, 0985-NN11-14.
17. H.J. Herbert, H.C. Moog. Cation exchange, interlayer spacing, MX-80 bentonite in high molar saline solutions. Eng. Geology, 1999, 54, 55-65.
18. M.H. Bradbury, B. Baeyens. Porewater chemistry in compacted re-saturated MX-80 bentonite. J. Contaminant Hydrology, 2003, 61, 329-338.

19. L. Carlson, O. Karnland, V.M. Oversby, A.P. Rance, N.R. Smart, M. Snellman, M. Vahanen, L.O. Werme. Experimental studies of interactions between anaerobically corroding iron and bentonite. *Phys. & Chem. Earth*, 2006, 32, 334-345.
20. K. Ueno, T. Suyama, T. Ishidera, T. Sato. Ten-year interaction of carbon steel coupon and compacted bentonites. *Clay Sci.*, 2011, 15, 89-102.
21. P. Wersin. Impact of corrosion-derived iron on the bentonite buffer within the KBS-3 concept. Posiva SKB Report: R-08-34, 2007.
22. Ch.W. Davies, C.T. Davie, E.A. Charles, M.L. White. Physicochemical and geotechnical alterations to the MX-80 bentonite at the waste canister interface in an engineered barrier system. *Geo Sci.*, 2017, 7, 69, 1-24.
23. F.A. Martin, C. Baillaon, M.L. Schlegel. Corrosion of iron and low alloyed steel within a water saturated brick of clay under anaerobic deep geological disposal conditions: An integrated experiment. *J. Nucl. Mater.*, 2008, 379, 80-90.
24. Bentonite–Iron Interactions in Natural Occurrences and in Laboratory. The effects of the interactions on the properties of Bentonite: A literature survey. Nuria Marcos Working Report 2003–55, 2003.
25. Bentonite–Iron Interactions in Natural Occurrences and in Laboratory. Bentonite–iron interactive surveillance survey. Nuria Marcos Working Report 2003–55, 2003.
26. F. King. Nuclear Waste Management Organization Report No. NWMO TR–2007–01, 2007. 27. D. W. Shoesmith. Assessing the corrosion performance of high-level nuclear waste containers. *Corrosion*, 2006, 62, 703-722.
28. J. McMurry, B.M. Ikeda, S. Stroes–Gascoyne, D.A. Dixon. Evolution of a Canadian geologic repository: Defective container scenario. Atomic Energy of Canada Ltd. Report No. 06819–REP–01200–10127–R00, 2004.
29. N.R. Smart, D.J. Blackwood, L. Werme. Part 1 – Electrochem. Aspects. *Corrosion*, 2002, 58, 547-559.
30. M.L. Schlegel, S. Necib, S. Daumas, M. Labat, C. Blanc, E. Foy, Y. Linard. Corrosion at the carbon steel–clay borehole water interface under anoxic alkaline and fluctuating temperature conditions. *Corrosion Sci.*, 2018, 136, 70-90.
31. J. Stouli, J. Kanoka, M. Kouril, H. Parschov, P. Novk. Impudence of temperature on corrosion rate and porosity of corrosion products of carbon steel in anoxic bentonite environment. *J. Nucl. Mater.*, 2013, 443, 20-25.
32. A.M. El-Shamy, M.F. Shehata, A.I.M. Ismai. Effect of moisture contents of bentonitic clay on the corrosion behavior of steel pipelines. *Appl. Clay Sci.*, 2015, 114, 461-466.
33. N.R. Smart, B. Reddy, A.P. Rance, D.J. Nixon, N. Diomidis. The anaerobic corrosion of carbon steel in saturated compacted bentonite in the Swiss repository concept. *Corrosion Eng. Sci. & Technol.*, 2017, 52, 113-126.
34. F. King, M. Kolar. Simulation of the anaerobic corrosion of carbon steel used fuel containers using the steel corrosion model version 1.0 (SCMV1.0). *Nucl. Waste Manag. Org. Report No. NWMO TR–2012–07*, 2012.
35. K. Mabuchi, Y. Horii, H. Takahashi, M. Nagayama. Effect of temperature and dissolved oxygen on the corrosion behavior of carbon steel in high-temperature water. *Corrosion*, 1991, 47, 500-508.
36. T. Ishikawa, Y. Kondo, A. Yasukawa, K. Kandori. Formation of magnetite in the presence of ferric oxyhydroxides. *Corrosion Sci.*, 1998, 40, 1239-1251.
37. F. King, S. Stroes, G. Squot. The Ontario Power Generation Report No. 1, 06819–REP–01200–1002–R00, 2000.
38. E. Torres, M.J. Turrero, A. Escribano, P.L. Martin. Long-term Performance of Engineered Barrier Systems PEBS. Grant Agreement No: FP7 249681, 2007.
39. K. Porsch, A. Kappler. Fe(II) oxidation by molecular O₂ during HCl extraction. *Environ. Chem.*, 2011, 8, 190-197.
40. P. Refait, J.M.R. Genin. The oxidation of ferrous hydroxide in chloride-containing aqueous media and Pourbaix diagrams of green rust one. *Corrosion Sci.*, 1993, 34, 797-819.
41. F.M. Al-Kharafi, B.G. Ateya, R.M. Abdallah. Electrochemical behavior of low carbon steel in concentrated carbonate chloride brines. *J. Appl. Electrochem.*, 2002, 32, 1363-1370.
42. S. El-Egamy, W.A. Badaway. Passivity and passivity breakdown of 304 stainless steel in alkaline sodium sulphate solutions. *J. Appl. Electrochem.*, 2004, 34, 1153-1158.
43. N.J. Laycock, R.C. Newman. Localised dissolution kinetics, salt films and pitting potentials. *Corrosion Sci.*, 1997, 39, 1771-1790.
44. Y. Kurimura, R. Ochiai, N. Matsuura. Oxygen oxidation of ferrous ions induced by chelation. *Bull. Chem. Soc. Jpn.*, 1968, 41, 2234-2239.
45. J.J. Ulrich, F.C. Anson. Ligand bridging by halide in electrochemical oxidation of chromium(II) at mercury electrodes. *Inorg. Chem.*, 1969, 8, 195-200.
46. R.M. Taylor. Influence of chloride on the formation of iron-oxides from Fe(II) chloride. II. Effect of [Cl] on the formation of lepidocrocite and its crystallinity. *Clays & Clay Minerals*, 1984, 32, 175-180.
47. P. Refait, J.M.R. Genin. The mechanisms of oxidation of ferrous hydroxychloride beta-Fe₂(OH)₃Cl in aqueous solution: The formation of akaganeite vs goethite. *Corrosion Sci.*, 1997, 39, 539-553.
48. U. Schwertmann, H. Fechter. The formation of green rust and its transformation to lepidocrocite. *Clay Minerals*, 1994, 29, 87-92.
49. D. Rezel, P. Bauer, J.M.R. Genin. Superparamagnetic behavior and hyperfine interactions in ferrous hydroxide II and green rust I. *Hyperfine Interactions*, 1988, 42, 1075-1078.

50. A.A. Olowe, J.M.R. Genin. The mechanism of oxidation of ferrous hydroxide in sulfated aqueous media: Importance of the initial ratio of the reactants. *Corrosion Sci.*, 1991, 32, 965-984.
51. R.M. Taylor. Influence of chloride on the formation of iron-oxides from Fe(II)chloride. I. Effect of [Cl] / [Fe] on the formation of magnetite. *Clays & Clay Minerals*, 1984, 32, 167-174.
52. C.T. Lee, M.S. Odziemkowski, D.W. Shoesmith. An *in situ* Raman-electrochemical investigation of carbon steel corrosion in Na₂CO₃ / NaHCO₃, Na₂SO₄, and NaCl solutions. *J. Electrochem. Soc.*, 2006, 153, B33-B41.
53. C.T. Lee, Z. Qin, M. Odziemkowski, D.W. Shoesmith. The influence of groundwater anions on the impedance behavior of carbon steel corroding under anoxic conditions. *Electrochimica Acta*, 2006, 51, 1558-1568.
53. P. Refait, J.A. Bourdoiseau, M. Jeannin, D.D. Nguyen, A. Romaine, R. Sabot. Electrochemical formation of carbonated corrosion products on carbon steel in deaerated solutions. *Electrochimica Acta*, 2012, 79, 210-217.
55. E.B. Castro, C.R. Valentini, C.A. Moina, J.R. Vilche, A.J. Arvia. The influence of ionic composition on the electrodissolution and passivation of iron electrodes in potassium carbonate-bicarbonate solutions in the 8.4 – 10.5 pH range at 25°C. *Corrosion Sci.*, 1986, 26, 781.
56. E.B. Castro, J.R. Vilche, A.J. Arvia. Iron dissolution and passivation in K₂CO₃-KHCO₃ solutions. Rotating-ring-disk electrode and XPS studies. *Corrosion Sci.*, 1991, 32, 37-50.
57. B.W.A. Sherar, P.G. Keech, Z. Qin, F. King, D.W. Shoesmith. Nominally anaerobic corrosion of carbon steel in near-neutral pH saline environments. *Corrosion*, 2010, 66, 045001, 1-11.
58. K. Videm, A.M. Koren. Corrosion, passivity, and pitting of carbon steel in aqueous solutions of HCO₃⁻, CO₂, and Cl⁻. *Corrosion*, 1993, 49, 746-754.
59. L. Legrand, S. Savoye, A. Chausse, R. Messina. Study of oxidation products formed on iron in solutions containing bicarbonate / carbonate. *Electrochimica Acta*, 2000, 46, 111-117.
60. M. Saheb, D. Neff, P. Dillmann, H. Matthiesen, E. Foy. Long term corrosion behavior of low-carbon steel in anoxic environment: Characterization of archaeological artifacts. *J. Nucl. Mater.*, 2008, 379, 118-123.
61. M. Saheb, D. Neff, L. Bellot-Gurlet, P. Dillmann. Raman study of a deuterated iron hydroxycarbonate to assess long-term corrosion mechanisms in anoxic soils. *J. Raman Spectroscopy*, 2011, 42, 1100-1108.
62. M. Saheb, D. Neff, P. Dillmann, H. Matthiesen, E. Foy, L. Bellot-Gurlet. Multisecular corrosion behavior of low carbon steel in anoxic soils: Characterization of corrosion system on archaeological artifacts. *Mater. & Corrosion*, 2009, 60, 99-105.
63. I. Azoulay, C. Remazeilles, P. Refait. Determination of standard Gibbs free energy of formation of chukanovite and Pourbaix diagrams of iron in carbonated media. *Corrosion Sci.*, 2012, 58, 229-236.
64. V. Pandarinathan, K. Lepkova, W. van Bronswijk. Chukanovite (Fe₂(OH)₂CO₃) identified as a corrosion product at sand-deposited carbon steel in CO₂-saturated brine. *Corrosion Sci.*, 2014, 85, 26-32.
65. C. Remazeilles, P. Refait. Fe(II) hydroxycarbonate Fe₂(OH)₂CO₃ (chukanovite) as iron corrosion product: Synthesis and study by Fourier transform infrared spectroscopy. *Polyhedron*, 2009, 28, 749-756.
66. M.L. Schlegel, C. Bataillon, C. Blanc, D. Pret, E. Foy. Anodic activation of iron corrosion in clay media under water-saturated conditions at 90 degrees C: Characterization of the corrosion interface. *Environ. Sci. & Technol.*, 2010, 44, 1503-1508.
67. M.L. Schlegel, C. Bataillon, F. Brucker, C. Blanc, D. Pret, E. Foy, M. Chorro. Corrosion of metal iron in contact with anoxic clay at 90 degrees C: Characterization of the corrosion products after two years of interaction. *Appl. Geo Chem.*, 2014, 51, 1-14.
68. S. Savoye, L. Legrand, G. Sagon, S. Lecomte, A. Chausse, R. Messina, P. Toulhoat. Experimental investigations on iron corrosion products formed in bicarbonate / carbonate containing solutions at 90 degrees C. *Corrosion Sci.*, 2001, 43, 2049-2064.
69. Y. Leon, M. Saheb, E. Drouet, D. Neff, E. Foy, E. Leroy, J.J. Dynes, P. Dillmann. Interfacial layer on archaeological mild steel corroded in carbonated anoxic environments studied with coupled micro and nano probes. *Corrosion Sci.*, 2014, 88, 23-35.
70. K. Azumi, M. Seo. Corrosion behavior of titanium-clad carbon steel in weakly alkaline solutions. *Corrosion Sci.*, 2003, 45, 413-426.
71. Y. Benamor, L. Bousselmi, H. Takenouti, E. Triki. Influence of sulphate ions on corrosion mechanism of carbon steel in calcareous media. *Corrosion Eng. Sci. & Technol.*, 2005, 40, 129-136.
72. E.M. Rus, I. Baldea, C. Calin. OL-T-35 low carbon steel corrosion in presence of sulfate ions. *Studia Universitatis Babes-Bolyai Chemia*, 2007, 52, 137-146.
73. G. Vatankhah, M. Drogowska, H. Menard, L. Brossard. Electrodissolution of iron in sodium sulfate and sodium bicarbonate solutions at pH 8. *J. Appl. Electrochem.*, 1998, 28, 173-183.
74. M.S. Rahman, S. Divi, D. Chandra, J. Daemen. Effect of different salts on the corrosion properties of friction type A607 steel rock bolt in simulated concentrated water. *Tunneling & Underground Space Technol.*, 2008, 23, 665-673.
75. P. Smith, S. Roy, D. Swailes, S. Maxwell, D. Page, J. Lawson. A model for the corrosion of steel subjected to synthetic produced water containing sulfate, chloride and hydrogen sulfide. *Chem. Eng. Sci.*, 2011, 66, 5775-5790.

76. K. Premlall, J.H. Potgieter, S. Sanja. Laser surface treatment to inhibit observed corrosion of reinforcing steel in sulphate: Alkaline media. *Anti-Corrosion Methods & Materials*, 2011, 58, 267-284.
77. Z. Zhu, X. Jiao, X. Tang, H. Lu. Effects of SO₄²⁻-concentration on corrosion behavior of carbon steels. *Anti-Corrosion Methods & Materials*, 2015, 62, 322-326.
78. G. Singh. A survey of underground mine waters from Indian coal mines. *Int. J. Mine Waters*, 2006, 5, 21-32.
79. P. Wersin, M. Snellman. Impact of iron on the performance of clay barriers in waste disposal systems. SKB Report R-08-45 on the status of research and development, 2008.
80. M. Birgersson, P. Wersin. KBS-3H reactive transport modeling of iron-bentonite interactions, an update for the report. Olkilu Working Report 2013-02, 2013.
81. D. Savage, C. Watson, S. Benbow, J. Wilson. Modeling iron-bentonite interactions. *Appl. Clay Sci.*, 2010, 47, 91-98.
82. F.M.I. Hunter, F. Bate, T.G. Heath, A. Hoch, L.O. Werme. Investigation of iron transport into bentonite from anaerobically corroding steel: A geochemical modeling study. *Int. Meeting Clays in natural and Engineering Barriers for Radioactive Waste Confinements*, 2007, Lille, 29-30.
83. D.L. Parkhurst, C.A.J. Appelo. User's guide to PHREEQC (version 2) – A computer program for speciation, batch-reaction, one-dimensional transport, and inverse geochemical calculations. US Geological Survey Water - Resources Investigations Report 99-4259, 1999.
84. K.A. Bond, T.G. Heath, C.J. Tweed. HATCHES: A referenced thermodynamic database for chemical equilibrium studies. Nirex Report NSS/R379, 1997.
85. Ch. Lu, J. Samper, B. Fritz, A. Clement, L. Montenegro. Interactions of corrosion products and bentonite: An extended multi component reactive transport model. *Phys. & Chem. Earth*, 2011, 36, 1661-1668.

1.2.1.3. Regio-controlled synthesis of double condensed oligo-, poly- and cyclo-phosphates, their characterization and possible solid-state applications. /M. Avaliani, E. Shapakidze, N. Barnovi, D. Dzanašvili, G. Todradze, V. Kveselava, N. Gongadze/. *Nano Studies*. – 2019. – #19. – pp. 273-284. – eng.; abs.: eng.

The importance of the efficient and resource-saving technologies for synthesis and applications of new inorganic polymers - condensed compounds are out of any doubt which explains the relevance of authors' works thanks to original technological methods and experience that been cultivated over many years. Multilateral spheres of application of condensed phosphates are very diverse: ion-exchange materials, nanomaterials, efficient applying fertilizers, detergents, cement substances, catalytic agents, raw materials for phosphates glasses, thermo-resistant substances, and also as food additive composites. Besides, the phosphate's binding agents, phosphate-binders and laser materials are supplanted/replaced by biomaterials, on the base of polyphosphates and hydroxyapatite. Present work reviews information about numerous crystals and/or powder of inorganic polymeric compounds, concretely, diverse new groups of condensed phosphates, region-controlled synthesized by the authors during investigations of the systems M^I₂O-M^{III}₂O₃-P₂O₅-H₂O at 400-850 K (where M^I are various monovalent metals, including Ag, and M^{III} - diverse trivalent metals). Numerous, specifically 85, new, formerly unknown, double condensed phosphates - so-called inorganic polymers have been obtained. The possibilities of their application are investigated. For example, synthesized cyclooctaphosphate K₂Ga₂P₈O₂₄ is studied for the catalytic activity. The results showed that it has unique properties as an inorganic polymer and can be used as a best catalyst during organic synthesis reactions, specifically, for the preparation of low molecular weight dienic oléfines. Accordingly to the experiments we established that synthesised acidic triphosphates of gallium and/or indium M^{III}H₂P₃O₁₀'(1-2)H₂O are the best ion-exchange materials. Tab. 1, Fig. 5, Ref. 38.

Keywords: regio-controlled synthesis, resource-saving technologies, polymeric compositions, biomaterials, catalytic activity

References

1. A. Durif. *Crystal Chemistry of Condensed Phosphates*. 2013, Dordrecht, Springer Sci.
2. D. Ribero, W.M. Kriven. Synthesis of LiFePO₄ powder by the organic-inorganic steric entrapment method. *J. Mater. Res.*, 2015, 30, 2133.
3. V.M. Castano, J. Quezada, A.K. De la Borbolla. New inorganic polymers: condensed phosphates and diverse spheres of their applications. In: Abs. World Forum Adv. Mater. "Polychar 26", 1, 2018, Tbilisi.

4. M.A. Avaliani. Synthesis and Characterization of Gallium Indium Condensed Phosphates (PhD Thesis), 1982, Moscow, N. Kurnakov Inst. General & Inorg. Chem.
5. E.V. Murashova, N.N. Chudinova. Double condensed phosphates of caesium–indium. *Inorg. Mater.*, 2001, 37, 12, 1521.
6. L.N. Komissarova. Inorganic and Analytical Chemistry of Scandium, 2006, Moscow.
7. N.P. Vassel, S.S. Vassel, I.V. Pavlova, A.V. Kaklyugin. Thermal study of systems involving the metaphosphates of trivalent metals and silver. *Proc. Inst. High. Educ. (Phys.)*, 2011, 1/2, 121.
8. N.Y. Strutinska, I.V. Zatovsky, I.V. Ogorodnyk, N.S. Slobodyannik. Rietveld refinement of $\text{AgCa}_{10}(\text{PO}_4)_7$ from X-ray powder data. *Acta Crystallographica E*, 2013, 69, 5, 23.
9. S.A. Suthanthiraraj, R. Sarumathi. A new silver ion conducting $\text{SbI}_3\text{--Ag}_4\text{P}_2\text{O}_7$ nano composite solid electrolyte. *J. Appl. Nanosci.*, 2013, 3, 501.
10. I. Grunze. Energy citations database. *Inorg. Mater.*, 2009, 23, 539.
11. M.A. Avaliani, I.V. Tananaev, M.K. Gvelesiani. Synthesis and investigation of double condensed phosphates of scandium and alkali metals. In: *Abs. Synth. Chem. Phosphorus, Sulfur, Silicon & Rel. Elements*, 2003 – 2011, Berlin, Fiz Chemie.
12. M. Avaliani, M. Gvelesiani. Areas of crystallization of condensed scandium and caesium phosphates and regularities of their formation. *Proc. Georgian Acad. Sci. (Chem. Ser.)*, 2006, 32, 52.
13. I.V. Tananaev, X. Grunze, N.N. Chudinova. Prior directions and results in the domain on condensed phosphates' chemistry. *Inorg. Mater.*, 1984, 20, 887.
14. M. Avaliani, M. Gvelesiani, N. Barnovi, B. Purtseladze, D. Dzanashvili. New investigations of poly-component systems. *Proc. Georgian Acad. Sci. (Chem. Ser.)*, 2016, 42, 308.
15. M. Avaliani. Main Types of condensed phosphates synthesized in open systems from solution–melts of phosphoric acids. In: *Abs. 4th Int. Conf. "Nanotechnologies"*, 2016, Tbilisi, GTU, 51.
16. T.P. Marsh. Studies into the Ion Exchange and Intercalation Properties of $\text{AlH}_2\text{P}_3\text{O}_{10}\cdot 2\text{H}_2\text{O}$ (PhD Thesis), 2011, Birmingham, Univ. Birmingham.
17. T. Yokoi, M. Kawashita, G. Kawachi, K. Kikuta, Ch. Ohtsuki. Synthesis of calcium phosphate crystals in a silica hydrogel containing phosphate ions. *J. Mater. Res.*, 2011, 24, 6, 2154.
18. M. Avaliani, M. Gvelesiani, B. Purtseladze, V. Kveselava, P. Nikoleishvili, G. Gorelishvili, R. Chagelishvili, N. Barnovi. Inorganic polymers double condensed phosphates of monoand polyvalent metals – General overview. In: *Abs. Int. Conf. "Innovative Technologies in Metallurgy and Materials Science"*, 2015, Tbilisi, GTU, 124.
19. M. Avaliani, D. Dzanashvili, M. Gvelesiani, N. Barnovi, E. Shapakidze. About new inorganic polymers-double condensed phosphates of silver and trivalent metals. *J. Chem. & Chem. Eng.*, 2017, 11, 60.
20. M. Avaliani. Investigation and thermal behaviour of double condensed compounds of gallium, scandium and silver. *Nano Studies*, 2018, 17/18, 21.
21. R. Dersch, M. Steinhart, U. Boudriot, A. Greiner, J.H. Wendorff. Nanoprocessing of polymers: Applications in medicine, sensors, catalysis, photonics. *Polymers Adv. Technol.*, 2005, 16, 276.
22. Sh. Suzuki., M.R. Whittaker, L. Grondahl, J. Monteiro, E. Wentrup–Byrne. Synthesis of soluble phosphate polymers by RAFT and their *in vitro* mineralization. *J. Biomacromolecules*, 2006, 7, 11, 3178.
23. Ed.P. Zanello. Chains, Clusters, Inclusion Compounds, Paramagnetic Labels. 2012, Amsterdam, Elsevier.
24. M. Avaliani, B. Purtseladze, E. Shohashvili, N. Barnovi. Apropos of inorganic polymerscondensed phosphates and spheres of their applications. *Proc. Georgian Acad. Sci. (Chem. Ser.)*, 2015, 41, 227.
25. W. Brostow, H.E. Hagg Lobland. Materials: Introduction and Applications, 2017, Wiley & Sons.
26. N.T. Kusnetsov, N.N. Chudinova, I.A. Rozanov. Analysis and synthesis, harmony and contrepoin. *Proc. Russ. Acad. Sci.*, 2004, 74, 460.
27. M. Avaliani, N. Barnovi, M. Gvelesiani, N. Esakia. Topics in double condensed phosphates' chemistry of mono- and trivalent metals. *Georg. Chem. J.* 2017, 1, 23.
28. J.E. Mark, H.R. Allcock, H.R. Allcock, R. West. Inorganic Polymers, 2005, Oxford, Oxford Univ. Press.
29. L.E. Lampila. Applications and functions of food-grade phosphates. *Ann. New York Acad. Sci.*, 2013, 1301, 37.
30. W. Brostow. Flexibility in relation to other properties of polymers. In: *Abs. World Forum Adv. Mater. "Polychar 26"*, 1, 2018, Tbilisi.
31. G. Kickelbick. Concepts for the incorporation of inorganic building blocks into organic polymers on a nanoscale. *Prog. Polymer Sci.*, 2003, 28, 83.
32. C. Trobajo, M.L. Rodriguez, M. Suarez, J.R. Garcia, J. Rodriguez, J.B. Parra, M.A. Salvado, P. Pertier, S. Garcia–Granda. Layered mixed tin-titanium phosphates. *J. Mater. Res.*, 2011, 13, 754.
33. D. Yang, R.F. Frindt. Structure of polymer intercalated MnPS_3 and CdPS_3 . *J. Mater. Res.*, 2000, 15, 2408.
34. H. Cheng. Inorganic dissolvable electronics: materials and devices for biomedicine and environment. *J. Mater. Res.*, 2016, 31, 2549.
35. M. Avaliani, V. Chagelishvili, E. Shapakidze, N. Esakia. Synthesis, characterization and applications of oligomeric, polymeric and cyclic condensed phosphates. *Res. J. Life Sci., Bioinformatics, Pharmaceutical & Chem. Sci.*, 2019, 5, 47.

36. T. Janssen. Icosahedral crystals, quasi-crystals: new forms of incommensurate crystal phases. *J. Phys. Colloques*, 1986, 47, 85.
37. M. Avaliani, E. Shapakidze. Areas of crystallization of double condensed phosphates of Ag and trivalent metals and regularities of their formation. *J. Chem. Sci.*, 2018, 9, 63.
38. K.K. Palkina, S.L. Maksimova, N.T. Chibiskova. Structure of crystals LiGa(PO₃)₄. *Inorg. Mater.*, 1981, 17, 95.

1.2.1.4. Photolysis of the phenolic solution in the presence of TiO₂ nanoparticles. /E. Gadirova, I. Hasanova, G. Aliyeva, G. Eyvazova, M. Aghayev, Z. Gakhramanova, A. Aliyev, H. Mahmudov, M. Akhundova, F. Muradova/. Azerbaijan Chemical Journal. – 2019. – #2. – pp. 52-57. – eng.; abs.: rus., az., eng. <https://doi.org/10.32737/0005-2531-2019-2-52-57>

TiO₂ nanoparticles have been prepared using sol-gel method from titanium(IV), isopropoxide, isopropyl alcohol, ammonia hydroxide as precursors at mild condition without any post heat treatment for crystallization. The nanocrystalline TiO₂ powders were composed of rutile form TiO₂ by X-ray diffraction. The size of nanoparticles was examined by transmission electron microscopy and is about 10⁻¹⁵ nm, and the Brunauer-Emmett-Teller specific surface area of the rutile nanopowder makes up 159.6 m²/g. Photochemical degradation of phenol solution under ultraviolet irradiation of the synthesized TiO₂ nanoparticles has been carried out. Fig. 4, Ref. 19.

Keywords: nanoTiO₂, phenol, photocatalytic process, X-ray diffraction, transmission electron microscopy

References

1. Yue D., Qian X., Zhao Y. Photocatalytic remediation of ionic pollutant. *Science Bulletin*. 2015. V. 60. No 21. P. 1791-1806.
2. Liu RL, Ye HY, Xiong XP, Liu HQ. Fabrication of TiO₂/ZnO composite nanofibers by electrospinning and their photocatalytic property. *Mater. Chem. Phys.* 2010. V. 121. No 3. P. 432-439.
3. Kim Y, Yang S, Jeon EH, Baik J, Kim N, Kim HS, Lee H. Enhancement of photo-oxidation activities depending on structural distortion of Fedoped TiO₂ nanoparticles. *Nanoscale Res. Lett.* 2016. 11. P. 41-50.
4. Yu H.J., Zhao Y.F., Zhou C., Shang L., Peng Y., Cao Y.H., Wu L.Z., Tung C.H., Zhang T.R. Carbon quantum dots/TiO₂ composites for efficient photocatalytic hydrogen evolution. *J. Mater. Chem. A*, 2013. V. 2. No 10. P. 3344-3351.
5. Fujishima A., Honda K. Electrochemical photolysis of water at a semiconductor electrode. *Nature*. 1972. V. 238. P. 37-38.
6. Turchi CS, Ollis DF. Photocatalytic degradation of organic water contaminants: mechanisms involving hydroxyl radical attack. *J. Catal.* 1990. V. 122. No 1. P. 178-192.
7. Zhao Y.F., Chen G.B., Bian T., Zhou C., Water-house GIN, Wu L.Z., Tung C.H., Smith L.J., O'Hare D., Zhang T.R. Defectrich ultrathin Zn/Al-layered double hydroxide nanosheets for efficient photoreduction of CO₂ to CO with water. *Adv Mater.* 2015. V. 27. No 47. P. 7824-7831.
8. Zhao Y.F., Zhao B., Liu J.J., Chen G.B., Gao R., Yao S.Y., Li M.Z., Zhang Q.H., Gu L., Xie J.L., Wen X.D., Wu L.Z., Tung C.H., Ma D., Zhang T.R. Oxide-modified nickel photocatalysts for the production of hydrocarbons in visible light. *Angew. Chem. Int. Ed.* 2016. V. 55. No 13. P. 4215-4219.
9. Kamat P.V., Meisel D. Nanoparticles in advanced oxidation processes. *Curr. Opin. Colloid Interface Sci.* 2002. V. 7. No 5-6. P. 282-287.
10. Amin S.A., Pazouki M., Hosseinnia A. Synthesis of TiO₂-Ag nanocomposite with solgel method and investigation of its antibacterial activity against *E. coli*. *Powder Technol.* 2009. V. 196. No 3. P. 241-245.
11. Tanaka K., Abe K., Hisanaga T. Photocatalytic water treatment on immobilized TiO₂ combined with ozonation. *J. Photochem Photobiol A Chem.* 1996. V. 101. No 1. P. 85-87.
12. Guan H.N., Chi D.F., Yu J., Li X.C. A novel photodegradable insecticide: preparation, characterization and properties evaluation of nanoimidacloprid. *Pesticide Biochemistry and Physiology*. 2008. V. 92. No 2. P. 83-91.
13. Okawa K., Suzuki K., Takeshita T., Nakano K. Degradation of chemical substances using wet peroxide oxidation under mild conditions. *J. Hazard. Mater.* 2005. V. 127. No 1-3. P. 68-72.
14. Zhang H., Chen G., Bahnemann D.W. Photoelectrocatalytic materials for environmental applications. *J. Mater. Chem.* 2009. V. 19. No 29. P. 5089-5121.
15. Roy P., Berger S., Schmuki P. TiO₂ nanotubes: synthesis and applications. *Angew Chem. Int. Ed.* 2011. V. 50. P. 2904-2939.
16. Gogos A., Knauer K., Bucheli T.D. Nanomaterials in plant protection and fertilization: current state, foreseen applications, and research priorities. *J. Agric. Food. Chem.* 2012. V. 60. No 39. P. 9781-9792.

17. Kumar S.G., Devi L.G. Review on modified TiO₂ photocatalysis under UV/visible light: selected results and related mechanisms on interfacial charge carrier transfer dynamics. *J. Phys. Chem. A.* 2011. V. 115. No 46. P. 13211-13241.
18. Chen X.B., Mao S.S. Titanium dioxide nanomaterials: synthesis, properties, modifications, and applications. *Chem. Rev.* 2007. V. 107. No 7. P. 2891-2959.
19. Kouloumbos V.N., Tsipi D.F., Hiskia A.E., Nikolic D., Breemen R.B. Identification of photocatalytic degradation products of diazinon in TiO₂ aqueous suspensions using GC/MS/MS and LC/MS with quadrupole time-of-flight mass spectrometry. *J. Am. Soc. Mass. Spectrom.* 2003. V. 14. No 8. P. 803-817.

1.2.1.5. Synthesis and physico-chemical study of platinum nanocomposite on mesoporous carbon nitride. /V. Akhmedov, N. Melnikova, A. Babayeva, G. Nurullayev, Z. Aliyeva, D. Tagiyev/. *Azerbaijan Chemical Journal.* – 2019. – #3. – pp. 6-14. – eng.; abs.: rus., az., eng.

New platinum nanocomposites were synthesized using mesoporous carbon nitride as a stabilizing matrix and carrier for nanoparticles by the method of reduction of H₂PtCl₆·6H₂O by methanol–water mixture. The textural, morphological, optical properties, structural and phase compositions of composites with different contents of loaded platinum particles were studied. The obtained composites have prospects for use as effective heterogeneous catalysts in chemical and photo-, -electrochemical processes. Fig. 7, Tab. 1, Ref. 38.

Keywords: mesoporous carbon nitride, nanocomposite, platinum, physicochemical methods of analysis, heterogeneous catalysis

References

1. Thomas A., Fischer A., Goettmann F., Antonietti M., Müller J.O., Schlögl R., Carlsson J.M. Graphitic carbon nitride materials: variation of structure and morphology and their use as metal-free catalysts. *J. Mater. Chem.* 2008. V. 18. P. 4893-4908.
2. Su D.Sh., Zhang J., Frank B., Thomas A., Wang X., Paraknowitsch J., Schlögl R. Metal-free heterogenous catalysis for sustainable chemistry// *ChemSusChem.* 2010. V. 3. P. 169-180.
3. Zhu J., Xiao P., Li H., and C. Carabineiro S.A. Graphitic carbon nitride: synthesis, properties and applications in catalysis. *ACS Appl. Mater. Inter-faces.* 2014. V. 6 . P. 16449-16465.
4. Dong G., Zhang Y., Pan Q., Qiu J. A fantastic graphitic carbon nitride (g-C₃N₄) material: Electronic structure, photocatalytic and photoelectronic properties. *J. Photochem. Photobiology C: Photo-chem. Reviews.* 2014. V. 20. P. 33-50.
5. Ida S. and Ishihara T. Recent progress in two-dimensional oxide photocatalysts for water splitting. *J. Phys. Chem. Lett.* 2014. V. 5. P. 2533-2542.
6. Wen J., Xie J., Chen X., Li X. A review on g-C₃N₄-based photocatalysts. *Appl. Surface Sci.* 2017. V. 391. P. 72-123.
7. Wang Y., Wang X., Antonietti M. Polymeric graphitic carbon nitride as a heterogeneous organocatalyst: from photochemistry to multipurpose catalysis to sustainable chemistry. *Angewandte Chemie International Edition.* 2012. V. 51. P. 68-89.
8. Zheng Y., Liu J., Liang J., Jaroniec M., Qiao Sh.Z. Graphitic carbon nitride materials: controllable synthesis and applications in fuel cells and photocatalysis. *Energy Environ. Sci.* 2012. Issue 5. P. 6717-6731.
9. Sun Sh. and Liang Sh. Recent advances in functional mesoporous graphitic carbon nitride (mpg-C₃N₄) polymers. *Nanoscale.* 2017. V. 9. P. 10544-10578.
10. McCusker L.B., Liebau F., Engelhardt G. Nomenclature of structural and compositional characteristics of ordered microporous and mesoporous materials with inorganic hosts (IUPAC Recommendations 2001). *Pure Appl. Chem.* 2001. V. 73. No 2. P. 381-384.
11. Akhmedov V.M., Melnikova N.E., Akhmedov I.D. Sintez, svoistva i primenie polimernykh nitridov ugleroda. *Izv. AN Ser. him.* 2017. № 5. S. 782-807. Akhmedov V.M., Melnikova N.E., Akhmedov I.D. Synthesis, properties, and application of polymeric carbon nitrides Russian Chem. Bulletin. In. Edition. 2017. № 5. P. 782-807.
12. Wang L., Wang Ch., Hu X., H.Xue, Pang H. Metal-/Graphitic carbon nitride composites: synthesis, structures, and applications. *Chem. As. J.* 2016. V. 11. P. 3305-3328.
13. Wang X., Chen X., Thomas A., Fu X. and Antonietti M. Metal-containing carbon nitride compounds: A new functional organic-metal hybrid material. *Adv. Mater.* 2009. V. 21. P. 1609-1612.
14. Zhao Z., Sun Y., Dong F. Graphitic carbon nitride based nanocomposites: A review *Nanoscale.* 2015. V. 7. P. 15-37.
15. Poling L., Poling P. Obshchaya himiya. M.: Mir, 1978. 674 s.
16. Jun J.S., Lee E.Z., Wang X., Hong W.H., Stucky G.D., Thomas A. From melamine-cyanuric acid supramolecular aggregates to carbon nitride hollow spheres. *Adv. Funct. Mater.* 2013. V. 23. P. 3661-3667.

17. Van Rheenen P.R., McKelvy M.J., Glaunsinger W.S. Synthesis and characterization of small platinum particles formed by the chemical reduction of chloroplatinic acid. *J. Solid State Chem.* 1987. V. 67. P. 151-169.
18. White R.J., Luque R., Budarin V.L., Clarkand J.H., Macquarrie D.J. Supported metal nanoparticles on porous materials. *Methods and applications. Chem. Soc. Rev.* 2009. V. 38. P. 481-494.
19. Kosuge K., Sato T., Kikukawa N., Takemori M. Morphological control of rod and fiberlike SBA-15 type mesoporous silica using water-soluble sodium silicate. *Chem. Mater.* 2004. V. 16. P. 899-905.
20. Vinu A., Ariga K., Mori T., Nakanishi T., Hishita S., Golberg D., Bando Y. Preparation and characterization of well-ordered hexagonal mesoporous carbon nitride. *Adv. Mater.* 2005. V. 17. P. 1648-1652.
21. Li X.H., Antonietti M. Metal nanoparticles at mesoporous N-doped carbons and carbon nitrides: functional Mott-Schottky heterojunctions for catalysis. *Chem. Soc. Rev.* 2013. V. 42. P. 6593-6604.
22. Yu J., Wang K., Xiao W., Cheng B. Photocatalytic reduction of CO₂ into hydrocarbon solar fuels over g-C₃N₄-Pt nanocomposite photocatalysts. *Phys. Chem. Chem. Phys.* 2014. V. 16. P. 11492-11501.
23. Hirai H., Nakao Y., Toshima N. Preparation of colloidal transition metals in polymers by reduction with alcohols and ethers. *J. Macromol. Sci.: Part A-Chemistry.* 1979. V. 13. P. 727-750.
24. Hirai H. Formation and catalytic functionality of synthetic polymernoble metal colloid. *J. Macromol. Sci.-Chem.* 1979. A 13. P. 633-649.
25. Chen Sh., Yang Q., Wang H., Zhang Sh., Li J., Wang Y., Chu W., Ye Q., Song L. Initial reaction mechanism of platinum nanoparticle in methanol - water system and the anomalous catalytic effect of water. *Nano Lett.* 2015. V. 15. P. 5961-5968.
26. Lin C.S., Khan M.R., Lin S.D. The preparation of Pt nanoparticles by methanol and citrate. *J. Colloid Interface Sci.* 2006. V. 299. P. 678-685.
27. Sing W. K.S., Everett D.H., Haul W.R.A., Moscou L., Pierotti R.A., Rouquerol J. and Siemieniewska T. Reporting physisorption data for gas/solid systems - with special reference to the determination of surface area and porosity. *Pure Appl. Chem.* 1985. V. 57. P. 603-619.
28. Wang X., Maeda K., Chen X., Takanabe K., Domen K., Hou Y., Fu X. and Antonietti M. Polymer semiconductors for artificial photosynthesis: hydrogen evolution by mesoporous graphitic carbon nitride with visible light. *J. Am. Chem. Soc.* 2009. V. 131. P. 1680-1681.
29. Ge L. Synthesis and photocatalytic performance of novel metalfree g-C₃N₄ photocatalysts. *Materials Lett.* 2011. V. 65. P. 2652-2654.
30. Sun B.-w., Yu H.-y., Yang Y.-j., Li H.-j., Zhai Ch.-y., Qian D.-J. and Chen M. New complete assignment of X-ray powder diffraction patterns in graphitic carbon nitride using discrete Fourier transform and direct experimental evidence. *Phys. Chem. Chem. Phys.* 2017. V. 19. P. 26072-26084.
31. Chen M.-L., Park Ch.-Y., Choi J.-G., Oh W.-Ch. Synthesis and characterization of metal (Pt, Pd and Fe) - graphene composites. *J. Korean Ceramic Soc.* 2011. V. 48. P. 147-151.
32. Ong W.-J., Tan L.-L., Chai S.-P., Yong S.-T. Hetero-junction engineering of graphitic carbon nitride (g-C₃N₄) via Pt loading with improved daylight-induced photocatalytic reduction of carbon dioxide to methane. *Dalton Trans.* 2015. V. 44. P. 1249-1257.
33. Harold P.K., Leroy E.A. *X-Ray diffraction procedures: for polycrystalline and amorphous materials (2nd Edition).* New York: John Wiley & Sons, 1974. 992 p.
34. Shalom M., Inal S., Fettkenhauer C., Neher D., An-tonietti M. Improving carbon nitride photocatalysis by supramolecular preorganization of monomers. *J. Am. Chem. Soc.* 2013. V. 135. P. 7118-7121.
35. Liu J.H., Zhang T.K., Wang Z.C., Dawson G., Chen W. Simple pyrolysis of urea into graphitic carbon nitride with recyclable adsorption and photocatalytic activity. *J. Mater. Chem.* 2011. V.21. P. 14398-14401.
36. Zhang G.G., Zhang J.S., Zhang M.W., Wang X.C. Polycondensation of thiourea into carbon nitride sem-iconductors as visible light photocatalysts. *J. Mater. Chem.* 2012. V. 22. P. 8083-8091.
37. Bojdys M.J., Müller J.O., Antonietti M., Thomas A. Ionothermal synthesis of crystalline, condensed, graphitic carbon nitride. *J. Chem. Eur.* 2008. V. 14. P. 8177-8182.
38. Bai X., Zong R., Li C., Liu D., Liu Y., Zhu Y. Enhancement of visible photocatalytic activity via Ag@C₃N₄ core-shell plasmonic composite. *Appl. Catalysis B: Environmental.* 2014. V. 147. P. 82-91.

1.2.1.6. The process of Ostwald maturation on TiGaTe₂ crystal surface. /K. Khalilova, N. Abdullayev, K. Kagramanov/. Azerbaijan Journal of Physics. –2019. – v. 25. – #1. – pp. 18-20. – eng.; abs.: eng.

It is shown that nano-island crystallization on TiGaTe₂ crystal surface is accompanied by their coalescence with process transition into Ostwald maturation. The analysis of conditions of Ostwald maturation (*OM*) stage beginning is carried out in the work. The growth mechanism on isotropic substrate on principle

doesn't differ from the one on crystal substrate. The process of condensate formation is divided on following parts: germformation, growth and coalescence of germs, channel formation. Fig. 5, Ref. 3.

Keywords: coalescence, maturation process, fluctuation, Ostwald maturation

References

1. R.M. Sardarli, O.A. Samedov, A.I. Nadjafov, A.P. Abdullayev, E.A. Zeynalova, D.G. Jabbarov. Azerbaijan National Academy of Sciences Transaction, FRTE ser, physics and astronomy, 2005, № 5, c. 45-50 (in Russian).
2. V.G. Dubrovskiy. Letters JTF, 2010, t.36, vip.5, p.53-61 (in Russian).
3. F.M. Cooney, A.K. Shekin, A.P Grinin. UFN, 2001. T. 171, p. 345-385 (in Russian).

2.2. Organic Materials

1.2.2.1. Removal of Fe(III) from aqueous solution by bio adsorption. /M. Azadi, Y. Mirzaie, Z. Shams Ghahfarokhi, A. Ebrahimitalab, A. Teimouri/. Nano Studies. – 2019. – #19. – pp. 131-144. – eng.; abs.: eng. In the present study, chitosan (CS) was modified by nano-diopside (nDP) and nano-hydroxyapatite (nHAp) using the facile freeze-drying method. The prepared biodegradable nano-composite displays high efficiency in adsorption of Fe(III) ion in aqueous solution. The chemical and morphological structures of nano-composite were investigated by scanning electron microscopy (SEM), thermogravimetric analysis (TGA), X-ray diffraction (XRD), and Fourier-transform infrared spectroscopy (FT-IR). Moreover, effects of pH, contact time, adsorbent concentration, and temperature on the adsorption process were studied. The best adsorption efficiency (27.7mg/g) was obtained with 0.01g of CS/nDP/nHAp, at the pH of 7 and with the contact time of 30 min at 37°C. Tab. 1, Fig. 9, Ref. 47.

Keywords: nano-diopside (nDP), chitosan (CS), biodegradable nanocomposite, scanning electron microscopy (SEM), Fourier – transform infrared spectroscopy (FT-IT)

References

1. R. Mauri, R. Shinnar, M. d'Amore, P. Giordano, A. Volpe. Solvent extraction of chromium and cadmium from contaminated soils. AIChEJ., 2001, 47, 509-512.
2. R.S. Juang, H.L. Huang. Mechanistic analysis of solvent extraction of heavy metals in membrane contactors. J. Membrane Sci., 2003, 213, 125-135.
3. H.R. Zafarani, M.E. Bahrololoom, M. Javidi, M.H. Shariat, J. Tashkhourian. Removal of chromate ion from aqueous solutions by sponge iron. Desalination & Water Treatment, 2014, 52, 7154-7162.
4. M. Constantin, I. Asmarandei, V. Harabagiu, L. Ghimici, P. Ascenzi, G. Fundueanu. Removal of anionic dyes from aqueous solutions by an ion-exchanger based on pullulan microspheres. Carbohydrate Polymers, 2013, 91, 74-84.
5. S. Choi, Y. Jeong. The removal of heavy metals in aqueous solution by hydroxyapatite/cellulose composite. Fibers & Polymers, 2008, 9, 267-270.
6. M. Muthukrishnan, B. Guha. Heavy metal separation by using surface modified nanofiltration membrane. Desalination, 2006, 200, 351-353.
7. Benatti CT, Tavares CRG, Lenzi E. Sulfate removal from waste chemicals by precipitation. Journal of environmental management. 2009, 90, 504-511.
8. R. Bazarqan-Lari, M. Bahrololoom, A. Nemati. Adsorption of Cu(II) ions from industrial wastewater on natural hydroxyapatite extracted from bone ash. J. Food Agric. Environ., 2011, 2, 652-657.
9. M. Shavandi, Z. Haddadian, M. Ismail, N. Abdullah, Z. Abidin. Removal of Fe(III), Mn(II) and Zn(II) from palm oil mill effluent (POME) by natural zeolite. J. Taiwan Ins. Chem. Eng., 2012, 43, 750-759.
10. A. Pelaez-Cid, I. Velazquez-Ugalde, A. Herrera-Gonzalez, J. Garcia-Serrano. Textile dyes removal from aqueous solution using *Opuntia ficus-indica* fruit waste as adsorbent and its characterization. J. Environ. Management, 2013, 130, 90-97.
11. A. Mudhoo, V.K. Garg, S. Wang. Removal of heavy metals by biosorption. Environ. Chem. Lett., 2012, 10, 109-117.
12. S. Saber-Samandari, M. Gazi. Cellulose-graft-polyacrylamide/hydroxyapatite composite hydrogel with possible application in removal of Cu(II) ions. Reactive & Functional Polymers, 2013, 73, 1523-1530.

13. X. Luo, J. Zeng, S. Liu, L. Zhang. An effective and recyclable adsorbent for the removal of heavy metal ions from aqueous system: Magnetic chitosan/cellulose microspheres. *Bioresource Technology*, 2015, 194, 403-406.
14. I. Oguz, F. Albayrak, A. Askin. Cu and Pb adsorption on some bentonitic clays. *Turkish J. Chem.*, 1998, 22, 243-252.
15. E. Erdem, N. Karapinar, R. Donat. The removal of heavy metal cations by natural zeolites. *J. Colloid & Interface Sci.*, 2004, 280, 309-314.
16. K. Krishnapriya, M. Kandaswamy. A new chitosan biopolymer derivative as metal complexing agent: synthesis, characterization, and metal(II) ion adsorption studies. *Carbohydrate Res.*, 2010, 345, 2013-2022.
17. A. Ghaei, M. Shariaty-Niassar, J. Barzin, T. Matsuura. Effects of chitosan membrane morphology on copper ion adsorption. *Chem. Eng. J.*, 2010, 165, 46-55.
18. R. Vieira, E. Guibal, E. Silva, M. Beppu. Adsorption and desorption of binary mixtures of copper and mercury ions on natural and crosslinked chitosan membranes. *Adsorption*, 2007, 13, 603-611.
19. Y.A. El-Reash, A. Abdelghany, A.A. El razak. Removal and separation of Cu(II) from aqueous solutions using nano-silver chitosan/polyacrylamide membranes. *International journal of biological macromolecules*. 2016, 86, 789-798.
20. R. Silva, G. Silva, O. Coutinho, J. Mano, R. Reis. Preparation and characterisation in simulated body conditions of glutaraldehyde crosslinked chitosan membranes. *J. Mater. Sci. : Mater. Med.*, 2004, 15, 1105-1112.
21. X.J. Huang, D. Ge, Z.K. Xu. Preparation and characterization of stable chitosan nanofibrous membrane for lipase immobilization. *Eur. Polymer J.*, 2007, 43, 3710-3718.
22. A. Bee, L. Obeid, R. Mbolantenaina, M. Welschbillig, D. Talbot. Magnetic chitosan/clay beads: A magsorbent for the removal of cationic dye from water. *J. Magn. & Magn. Mater.*, 2017, 421, 59-64.
23. R. Schmuhl, H. Krieg, K. Keizer. Adsorption of Cu(II) and Cr(VI) ions by chitosan: Kinetics & Equilibrium Studies, 2001, 27, 1-8.
24. Y. Zhang, S. Li, C. Wu. The in vitro and in vivo cementogenesis of CaMgSi₂O₆ bioceramic scaffolds. *J. Biomed. Mater. Res. A*, 2014, 102, 105-116.
25. C. Wu, Y. Ramaswamy, H. Zreiqat. Porous diopside (CaMgSi₂O₆) scaffold: a promising bioactive material for bone tissue engineering. *Acta Biomaterialia*, 2010, 6, 2237-2245.
26. T. Nonami, S. Tsutsumi. Study of diopside ceramics for biomaterials. *J. Mater. Sci.: Mater. Med.*, 1999, 10, 475-479.
27. T. Nonami, S. Tsutsumi, M. Fukuma, T. Urabe. Mechanical properties and biocompatibility of diopside. *J. Jpn. Soc. Biomater.*, 1994, 12, 274-278.
28. Y Ning, J. Li, W. Cai, X. Shao. Simultaneous determination of heavy metal ions in water using near-infrared spectroscopy with preconcentration by nano-hydroxyapatite. *Spectrochimica Acta A: Mol. & Biomol. Spectroscopy*, 2012, 96, 289-294.
29. S. Kongsri, K. Janpradit, K. Buapa, S. Techawongstien, S. Chanthai. Nanocrystalline hydroxyapatite from fish scale waste: Preparation, characterization and application for selenium adsorption in aqueous solution. *Chem. Eng. J.*, 2013, 215, 522-532.
30. F. Zhuang, R. Tan, W. Shen, X. Zhang, W. Xu, W. Song. Monodisperse magnetic hydroxyapatite/Fe₃O₄ microspheres for removal of lead(II) from aqueous solution. *J. Alloys & Comp.*, 2015, 637, 531-537.
31. W. Wei, R. Sun, J. Cui, Z. Wei. Removal of nitrobenzene from aqueous solution by adsorption on nanocrystalline hydroxyapatite. *Desalination*, 2010, 263, 89-96.
32. W. Wei, R. Sun, Z. Wei, H. Zhao, H. Li, F. Hu. Elimination of the interference from nitrate ions on oxalic acid in RP-HPLC by solid-phase extraction with nanosized hydroxyapatite. *J. Liquid Chromatography & Rel. Technol.*, 2008, 32, 106-124.
33. A. Moatary, A. Teimouri, M. Bagherzadeh, A.N. Chermahini, R. Razavizadeh. Design and fabrication of novel chitin hydrogel/chitosan/nano diopside composite scaffolds for tissue engineering. *Ceram. Int.*, 2016, 43, 1657-1668.
34. L. Wang, C. Li. Preparation and physicochemical properties of a novel hydroxyapatite/chitosan-silk fibroin composite. *Carbohydrate Polymers*, 2007, 68, 740-745.
35. S. Wang, L. Shen, Y. Tong, L. Chen, I. Phang, P. Lim, T. Liu. Biopolymer chitosan/montmorillonite nanocomposites: preparation and characterization. *Polymer Degradation & Stability*, 2005, 90, 123-131.
36. A. Teimouri, M. Azadi. Preparation and characterization of a novel chitosan/nanodiopside/nanohydroxyapatite composite scaffolds for tissue engineering applications. *Int. J. Polym. Mater. & Polym. Biomater.*, 2016, 65, 917-927.
37. J. Venkatesan, Z.J. Qian, B. Ryu, N.A. Kumar, S.K. Kim. Preparation and characterization of carbon nanotube-grafted-chitosan-natural hydroxyapatite composite for bone tissue engineering. *Carbohydrate Polymers*, 2011, 83, 569-577.
38. L. Ghorbanian, R. Emadi, S.M. Razavi, H. Shin, A. Teimouri. Fabrication and characterization of novel diopside/silk fibroin nanocomposite scaffolds for potential application in maxillofacial bone regeneration. *Int. J. Bio. Macromol.*, 2013, 58, 275-280.
39. W. Wei, L. Yang, W. Zhong, J. Cui, Z. Wei. Poorly crystalline hydroxyapatite: A novel adsorbent for enhanced fulvic acid removal from aqueous solution. *Appl. Sur. Sci.*, 2015, 332, 328-339.

40. Z. She, B. Zhang, C. Jin, Q. Feng, Y. Xu. Preparation and *in vitro* degradation of porous three-dimensional silk fibroin/chitosan scaffold. *Polymer Degradation & Stability*, 2008, 93, 1316-1322.
41. Y. Liu, M. Chen, H. Yongmei. Study on the adsorption of Cu(II) by EDTA functionalized Fe₃O₄ magnetic nanoparticles. *Chem. Eng. J.*, 2013, 218, 46-54.
42. Y.M. Hao, C. Man, Z.B. Hu. Effective removal of Cu(II) ions from aqueous solution by amino-functionalized magnetic nanoparticles. *J. Hazardous Mater.*, 2010, 184, 392-399.
43. G. Kousalya, M.R. Gandhi, C.S. Sundaram, S. Meenakshi. Synthesis of nanohydroxyapatite chitin/chitosan hybrid biocomposites for the removal of Fe(III). *Carbohydrate Polymers*, 2010, 82, 594-599.
44. A.M. Donia, A.A. Atia, K.Z. Elwakeel. Selective separation of mercury(II) using magnetic chitosan resin modified with Schiff's base derived from thiourea and glutaraldehyde. *J. Hazardous Mater.*, 2008, 151, 372-379.
45. F. Mai, C. Chen, J. Chen, S. Liu. Photodegradation of methyl green using visible irradiation in ZnO suspensions: determination of the reaction pathway and identification of intermediates by a high-performance liquid chromatography-photodiode arrayelectrosprayionization-mass spectrometry method. *J. Chromatogr. A*, 2008, 1189, 355-365.
46. Z.M. El-Bahy, M.M. Mohamed, F.I. Zidan, M.S. Thabet. Photo-degradation of acid green dye over Co-ZSM-5 catalysts prepared by incipient wetness impregnation technique. *J. Hazardous Mater.*, 2008, 153, 364-371.
47. A.E.H. Machado, J.A. de Miranda, R.F. de Freitas, E.T.F. Duarte, L.F. Ferreira, Y.D. Albuquerque, R. Ruggiero, C. Sattler, L. de Oliveira. Destruction of the organic matter present in effluent from a cellulose and paper industry using photocatalysis. *J. Photochem. & Photobiol. A*, 2003, 155, 231-241.

1.2.2.2. Effect of organic matter on hexavalent chromium reduction by nano zero valent iron in soil. /G. Allahverdiyeva/. Azerbaijan Chemical Journal. – 2019. – #2. – pp. 11-14. – eng.; abs.: rus., az., eng. <https://doi.org/10.32737/0005-2531-2019-2-11-14>

The experimental data are presented on lowering concentration of chrome(VI) ions in soil (with a purpose of rendering it harmless) by nanoparticles of iron in the presence of high concentration of an organic compound (saccharase) as a reducing agent. The reiterated repetition experiments has proved an effectiveness of synthesized before iron nanoparticles. The basic parameters of an effective reduction of Cr(VI) by nanoparticles of iron have been found out. They are as follows: the time of reduction (2 h) and concentration of organic substances (35.71 g/kg of soil). It has been established that a rise in reduction duration above 2 h lowers its effectiveness. The cause of this, revealed by us, consists in that a high concentration of the organic compound provides for inhibiting of Cr(VI). It occurs on account of 1) adsorption of substance on a surface of iron nanoparticles and as a result the decrease in their active surface, 2) interaction of the organic compound with the products of the reduction reaction, 3) dissolution of saccharase in liquid phase. As result of investigation the necessity of determining the significance of high concentration of the organic substance for an effective reduction of Cr(VI) in soil is exposed. At present the work in this direction is being continued. Fig. 1, Tab. 1, Ref. 13.

Keywords: Fe⁰ nanoparticles, organic matter, Cr(VI) reduction

References

- Xu X.R., Li H.B., Li X.Y., Gu, J.D. Reduction of hexavalent chromium by ascorbic acid in aqueous solutions. *Chemosphere*. 2004. V. 57. P. 609-613.
- Di Palma L, Gueye M.T., Petrucci E. Hexavalent chromium reduction in contaminated soil: a comparison between ferrous sulphate and nanoscale zero valent iron. *J. Hazard. Mater.* 2015. V. 281. P. 70-76.
- Jacobs J.A., Testa S.M., Overview of chromium(VI) in the environment: background and history. In: Guertin J., Jacobs J.A., Avakian C.P., editors. Chromium(VI) handbook. New York: CRC Press; 2005. P. 1-21.
- Alowitz M.J., Scherer M.M., Kinetics of nitrate, nitrite, and Cr(VI) reduction by iron metal. *Environ. Sci. Technol.* 2002. V. 36. P. 299-306.
- Oh., Y.J., Song H., Shin W.S., Choi S.J., Kim Y.H., Effect of amorphous silica and silica sand on removal of chromium(VI) by zero valent iron. *Chemosphere*. 2007. V. 66. P. 858-865.
- Wittbrodt P.R., Palmer C.D., Reduction of Cr(VI) in the presence of excess soil fulvic acid. *Environ. Sci. Technol.* 1995. V. 29. P. 255-263.
- Tandy S., Bossart K., Mueller R., Ritschel J., Hause L., Schulin R., Nowack B., Extraction of heavy metals from soils using biodegradable chelating agents. *Environ. Sci. Technol.* 2004. V. 38. P. 937-944.

8. Zhou H., He Y., Lan Y., Mao J., Chen S., Influence of complex reagents on removal of chromium(VI) by zero-valent iron. *Chemosphere* 2008. V. 72. P. 870-874.
9. Gonzini O., Plaza A., Di Palma L., Lobo M.C., Electrokinetic remediation of a soil contaminated soil enhanced by rhamnolipid. *J. Appl. Electrochemistry*. 2010. V. 40. P. 1239-1248.
10. Liu C., Evett J.B., Soil properties, Testing, Measurement, and Evaluation. 5th ed. Prentice-Hall. 2002. New-York, USA.
11. Bartlett R.J., Chromium cycling in soils and water: links, gaps, and methods. *Environ. Health Perspectives*. 1991. V. 92. P. 17-24.
12. He F., Zhao D., Manipulating the Size and Dispersibility of Zerovalent Iron Nanoparticles by Use of Carboxymethyl Cellulose Stabilizers. *Environ. Sci. Technol.* 2007. V. 41. P. 6216-6221.
13. Wu Y.J., Zhang J.H., Tong Y.F., Xu X.H., Chromium(VI) reduction in aqueous solutions by Fe₃O₄-stabilized FeO nanoparticles. *J. Hazard. Mater.* 2009. V. 172. P. 1640-1645.

3. NANOBIOLOGY

3.1. Uptake

1.3.1.1. Change of biomass of cyanobacteria *Spirulina platensis* in nutrient medium Zarrouk at introducing of known amount of chemical elements: Cr, Se, Zn, Ni, Cu, Ag, Cd, and Hg, separately. /N. Kuchava/. Nano Studies. – 2019. – #19. – pp. 263-272. – geo.; abs.: geo.

The problem of the change of biomass of one of the biotechnologically interesting blue-green cyanobacteria *Spirulina platensis* has been studied when the preliminarily determined amount of chemical elements such as chromium (Cr), selenium (Se), zinc (Zn), nickel (Ni), copper (Cu), silver (Ag), cadmium (Cd), and mercury (Hg) were introduced separately, during the initial stage of the experiment. The method of atomic absorption spectrometry was used. Tab. 8, Ref. 13.

Keywords: biomass, cyanobacteria, nutrient medium, chemical elements atomic absorption microscopy

References

1. L. P. Loseva, L. V. Dardinskaya. In: 6th Int. Cong. Appl. Algology, Czech Republic, 1993.
2. L.P. Loseva. In: 8th Int. Cong. Appl. Algology, Italy, 1999.
3. ნ. წიბაძემვილი, ა. რჩეულიშვილი, ე. გინტური, ნ. კუჭავა, ნ. ბაღდავაძე, ვ. გაბუნია. თეზ. კრ.: 1-ლი საერთაშორისო კონფ. „ნანოქიმია – ნანოტექნოლოგიები“, 2010, თბილისი, 37-38.
4. ნ. კუჭავა. საქართველოს მეცნიერებათა ეროვნული აკადემიის მაცნე (ქიმიის სერია – გარემოს ქიმია), 2017, 43, 1, 84-87.
5. ნ. კუჭავა. საქართველოს მეცნიერებათა ეროვნული აკადემიის მაცნე (ქიმიის სერია – გარემოს ქიმია), 2017, 43, 2, 232-235.
6. ნ. კუჭავა. Nano Studies, 2014, 9, 119-126.
7. ნ. კუჭავა. Nano Studies, 2014, 10, 111-116.
8. ლ.მ. მისულიშვილი, ა.ი. ველობილსკი, ე.ი. კირკესალი, ა.ი. ხივანიშვილი, ე.ნ. გინტური, ნ.ე. კუჩავა, მ.ვ. ფრონტაშვილი, ს.ს. პავლოვ, ნ.გ. აკსენოვა. Препринт ОИЯИ: Р18–2008, 2008.
9. Zh.Y. Li, S.Y. Guo, L.Li. J. Food Eng., 2006, 75, 1, 129-136.
10. ნ. კუჭავა. საქართველოს მეცნიერებათა ეროვნული აკადემიის მაცნე (ქიმიის სერია - გარემოს ქიმია), 2017, 43, 3-4, 363-366.
11. S. Chandrabhan, H.S. Viayarti, R. Verma, S.P. Seniya, Sh. Vyas, S. Singh Trivedia. J. Chem. Pharm. Res., 2012, 4, 4, 2250- 2254.
12. A. Khizanishvili, A. Belokobilsky, E. Ginturi, N. Kuchava, A. Rcheulishvili. In: Workbook 8th Int. Conf. Pharmacy & Appl. Phys. Chem. 2004, Monte Verita, Stefano Franscini, PO 24.
13. P. Jonson, A. Shubert. Nutr. Rep. Int., 1986, 34, 6, 1063-1071.

1.3.1.2. Influence of Mg on uptake of Zn, Cu, and Cr by *Arthrobacter globiformis* 151B. /A. Rcheulishvili, L. Tugushi, E. Ginturi, M. Osepashvili, O. Rcheulishvili, M. Gurjelidze/. Nano Studies. – 2019. – #19. – pp. 293-298. – geo.; abs.: geo.

To characterize metals joint action on highly resistant gram-positive aerobic bacteria, we have studied the influence of Mg ions on the accumulation capacity of Zn, Cu, and Cr ions by *Arthrobacter globiformis* 151B, which is promising candidate for good bioremediation. Bacterial cells were incubated at 26°C during different periods of time: 16, 24, 48, 96, and 144 h. For the experiment, the concentration of magnesium in the growth medium (TSB Broth) made up 1600 µg/ml. After that, they were harvested and tested on the content of research elements (zinc, copper and chromium) using atomic absorption spectroscopy. According to the research data, high concentration of magnesium can increase the accumulation capacity of zinc, copper and chromium over time. Fig. 3, Ref. 14.

Keywords: metals joint action, highly resistant gram-positive bacteria, accumulation capacity, bacterial cells

References

1. P.B. Tchounwou, C.G. Yedjou, A.K Patlolla, D.J. Sutton. Heavy metals toxicity and the environment. *Exp. Suppl.*, 2012, 101, 133.
2. Y. Xie, J. Fan, W. Zhu, E. Amombo, Y. Lou, L. Chen, J. Fu. Effect of heavy metals pollution on soil microbial diversity and bermudagrass genetic variation. *Front Plant Sci.*, 2016, 7, 755.
3. Zh. Yao, J. Li, He. Xie, C. Yu. Review on remediation technologies of soil contaminated by heavy metals. *Procedia Environ. Sci.*, 2012, 16, 722.
4. S. Khalid, M. Shahid, N.Kh. Niazi, B. Murtaza, I. Bibi, C. Dumat. A comparison of technologies for remediation of heavy metal contaminated soils. *Geochem. Exploration B*, 2017, 182, 247.
5. N.Y. Tsibakashvili, T.L. Kalabegishvili, A.N. Rcheulishvili, E.N. Ginturi, L.G. Lomidze, O.A. Rcheulishvili, D.N. Gvarjaladze, H.Y. Holman. Effect of Zn(II) on the reduction and accumulation of Cr(VI) by Arthrobacter species. *J. Ind. Microbiol.*, 2011, 38, 1803.
6. N.Y. Tsibakashvili, L.M. Mosulishvili, T.L. Kalabegishvili, D.T. Pataraya, M.A. Gurielidze, G.S. Nadareishvili, H.Y.N. Holman. Chromate-resistant and reducing microorganisms in Georgia basalts: their distribution and characterization. *Fresenius Environ. Bull.*, 2002, 11, 352-361.
7. N. Asatiani, T. Kartvelishvili, N. Sapojnikova, M. Abuladze, L. Asanishvili, M. Osepashvili. Effect of the simultaneous action of zinc and chromium on Arthrobacter spp. *Water Air Soil Pollut.*, 2018, 229.
8. M. Megraharaj, S. Avudainayagam, R. Neidu. Toxicity of hexavalent chromium and its reduction by bacteria isolated from soil contaminated with tannery waste. *Curr. Microbiol.*, 2003, 47, 51.
9. S.P. Kamaludeen, M. Megharaj, N. Sethunathan, R. Naidu. Chromium-microorganism interactions in soils: Remediation implications. *Rev. Environ. Contam. Toxicol.*, 2003, 178, 93.
10. მ. რჩეულიშვილი. გარემოში მობინადრე Arthrobacter-ის გვარის ბაქტერიების მიერქომის შეთვისება და დეტოქსიფიცია და აღნიშნულ პროცესზე სხვადასხვა მეტალური იონის გავლენა (სადოქტორო დისერტაცია), 2014, თბილისი.
11. Ch.A. McDevitt, A.D. Oggunniyi, E. Valkov, M.C. Lawrence, B. Kobe, A.G. McEwan, J.C. Paton. A molecular mechanism for bacterial susceptibility to zinc. *PLoS Pathog.*, 2011, 7, e1002357.
12. R. Choudhury, S. Srivastava. Zinc resistance mechanism in bacteria. *Curr. Sci.*, 2001, 81, 768.
13. E.A. Groisman, K. Hollands, M.A. Kriner, E.J. Lee, S.Y. Park, M.H. Pontes. Bacterial Mg²⁺ homeostasis, transport, and virulence. *Ann. Rev. Genet.*, 2013, 47, 625.
14. F. Altimira, C. Yanez, G. Bravo, M. Gonzalez, L.A. Rojas, M. Seeger. Characterization of copper-resistant bacteria and bacterial communities from copper-polluted agricultural soils of central Chile. *BMC Microbiol.*, 2012, 12, 193.

4. NANOTECHNOLOGY

4.1. Materials and Structures

1.4.1.1. Method for manufacturing nanocrystalline systems from liquid charge. /A. Mikeladze, O. Tsagareishvili, L. Chkhartishvili, R. Chedia, R. Tsiskarishvili/. Nano Studies. – 2019. – #19. – pp. 15-36. – rus.; abs.: rus.

A technology has been developed for the production of nanocrystalline systems of a wide range by spraying solutions of salts of the corresponding components in liquid organic compounds into a reactor with a reducing or carbide-forming medium. After pyrolysis, selective reduction processes and carbidization, powders are obtained, each particle of which can consist of nanocrystalline components (with crystallites sizes of 10-80 nm depending of initial solution) - metals, oxides, carbides, and quasibinary ceramics. In this way were obtained: (1) dispersion-strengthened transition metals and alloys with increased strength while retaining plastic properties; (2) titanium, tungsten, and boron carbides nanocrystalline cermets with high hardness and enhanced strength; and (3) nanocrystalline ceramics based on boron carbide/titanium diboride composition B_4C-TiB_2 with high hardness and increased impact toughness. Fig. 16, Ref. 10.

Keywords: nanocrystalline systems, spraying solution, carbide-forming medium, liquid charge

References

1. А.А. Гачечиладзе, В.В. Кияненко, А.Г. Микеладзе, Ф.Н. Тавадзе. Влияние неорганических дисперсных частиц на внутреннее трение и механические свойства железа и никеля. Сообщ. АН Груз. ССР, 1986, 124, 2, 357-360.
2. А.Г. Микеладзе. Литые дисперсно-упрочненные металлы. Литейное производство, 2001, 7, 4-6.
3. N. Jalabadze, A. Mikeladze, R. Chedia. New technology for the manufacturing of new crystalline materials. IEEE– Nano, 2002, TA 3, August, 273-276.
4. D. Stover, R. Vassen, A. Mikeladze, A. Gachechiladze, T. Mika Beridze, N. Zoidze, V. Gabunia. Elaboration of nanocrystalline metal-ceramics for functional layers of solid oxide fuel cells. Energy, 2003, 4 (28), 115-119.
5. A. Gachechiladze, D. Gabunia, A. Mikeladze, O. Tsagareishvili, L. Chkhartishvili. On technological process of boron carbide nanodispersive powders. In: Proc. 4th Int. Boron Symp. (Eds. A. Konuk, H. Kurama, H. Ak, M. Iphar), 2009, Ankara, Osmangazi Univ.–ТММОВ, 67-71.
6. А.А. Гачечиладзе, А.Г. Микеладзе, Д.Л. Габуния, Б.Г. Маргиев, Р.В. Чедия, Т.А. Цуладзе, Р.К. Зекалашвили, О.А. Цагарейшвили. На нанокристаллические твердые сплавы на основе карбида титана: Получение компактирование. Nano Studies, 2010, 2, 23-29.
7. M. Antadze, R. Chedia, O. Tsagareishvili, A. Mikeladze, A. Gachechiladze, B. Margiev, D. Gabunia, T. Tsuladze, D. Khantadze. Metal-ceramics based on nanostructured boron carbide. Solid State Sci., 2012, 14, 11-12, 1725-1728.
8. А.Г. Микеладзе, А.А. Гачечиладзе, Б.Г. Маргиев, О.А. Цагарейшвили. Металлокерамика на основе нанокристаллических абразивных карбидов. В сб. докл. 1-й Междунар. конф. «Современные технологии и методы неорганического материаловедения» (Ред. Дж.В. Хантадзе, Д.Л. Габуния, Д.Р. Рамазашвили, Л.С. Чхартишвили), 2012, Тбилиси, Меридиан, 149-160.
9. ა. მიქელაძე, ო. ცაგარეშვილი, ლ. ჩხარტიშვილი, რ. ჭედია. B_4C-TiB_2 ნანოსტრუქტურული კერამიკული შენადობების მიღება. საქ. მეცნ. ეროვ. აკად. მაცნე (მრავალდაროვანი კვლევების სერია), 2017, 2, 37-44.
10. A. Mikeladze, O. Tsagareishvili, L. Chkhartishvili, M. Darchiashvili, R. Chedia, K. Sarajishvili. Synthesizing of nanostructured composite ceramics B_4C-TiB_2 and their properties. In: Abs. 19th Int. Symp. Boron, Borides & Rel. Mater., 2017, Freiburg, Albert-Ludwigs Univ., 58-58.

1.4.1.2. Electrodeposition of tungsten nanostructured galvanic coatings from tungstate containing melts. /V. Malyshev, A. Gab, D. Bruskova, T. Dmytrenko, M. Gaune-Escard/. Nano Studies. – 2019. – #19. – pp. 77-86. – eng.; abs.: eng.

Electroreduction of dimer complexes of tungsten in metaphosphate-containing melts is studied using methods of voltammetry. Deposition of tungsten and tungsten-molybdenum coatings from oxyhalide and

oxide melts and the influence of the electrolysis parameters on the properties of the deposits are investigated. Tab. 1, Fig. 7, Ref. 13.

Keywords: electroredeposition, tungsten, nanostructured galvanic coating, electrolysis, deposits

References

1. V.I. Shapoval, V.V. Soloviev, V.V. Malyshev. Electrochemically active species and multielectron processes in ionic melts. Russ. Chem. Rev., 2001, 70, 2, 161-176.
2. C.K. Gupta. Chemical Metallurgy: Principles and Practice, 2003, John Wiley & Sons.
3. V.V. Malyshev, V.V. Soloviev, L.A. Chernenko, V.N. Rozhko. Management of composition cathodic products in the electrolysis of molybdenum-, tungsten- and carbon-bearing halogenide-ohide and oxide melts. Mater. Sci. & Eng. Technol., 2015, 45, 11, 67-72.
4. S. Glastone. An Introduction to Electrochemistry, 2006, Kindle Book.
5. V. Malyshev. Formation of cathode products of electrolysis of tungstate-molybdate melts. Mater. Sci., 2012, 48, 2, 208-212.
6. A.O. Omelchuk, S.V. Volkov, O.G. Zarubitskii. Modern electrochemical processes and technologies. J. Mining & Metall. B, 2003, 39, 1-2, 93-107.
7. V. Malyshev. Electrodeposition of different types of tungsten cathode deposits from ionic melts. Mater. Sci., 2011, 47, 3, 345-354.
8. V. Malyshev, A. Gab, A.M. Popescu, V. Constantin. Electroreduction of tungsten oxide(VI) in molten salts with added metaphosphate. Chem. Res. Chinese Univ., 2013, 29, 4, 771-775.
9. M. Erdogan, I. Karakaya. Electrochemical reduction of tungsten compounds to produce tungsten powder. Metall. & Mater. Trans. B, 2010, 41, 4, 798-804.
10. R. Abdulaziz, L.D. Brown, D. Inman, S. Simons, P.R. Shearing, D.J.L. Brett. Novel fluidised cathode approach for the electrochemical reduction of tungsten oxide in molten LiCl-KCl eutectic. Electroch. Commun., 2014, 41, 4, 44-46.
11. K. Nitta, T. Nohira, R. Hagiwara, M. Majima, S. Inazawa. Electrodeposition of tungsten from ZnCl₂-NaCl-KCl-WO₃ melt and investigation on tungsten species in the melt. Electrochimica Acta, 2010, 55, 3, 1278-1281.
12. H. Nakajima, T. Nohira, R. Hagiwara. Electrodeposition of metallic tungsten in ZnCl₂-NaCl-KCl-WCl₄ melt at 250°C. Electrochim. Solid-State Lett., 2005, 8, 7, C91-C94.
13. K. Nitta, M. Majima, S. Inazawa, T. Nohira, R. Hagiwara. Electrodeposition of Tungsten from Li₂WO₄-Na₂WO₄-K₂WO₄ based melts. Electrochim., 2009, 77, 8, 621-623.

1.4.1.3. Synthesis of nanostructured coatings molybdenum(tungsten)-nickel(cobalt) in tungstate melts.
/V. Malyshev, D. Shakhnin, A. Gab, T. Lukashenko, O. Dmytrenko/. Nano Studies. – 2019. – #19. – pp. 87-96.
– eng.; abs.: eng.

The electrochemical behavior of nickel and cobalt in a tungstate melt and the effect of the electrolysis conditions on the composition and structure of the deposits of the nickel (cobalt)-molybdenum (tungsten) alloys and intermetallic compounds in oxide tungstate melts are studied. As the molybdenum (tungsten) concentration in the melt increases or the nickel (cobalt) concentration decreases, the phase composition of the cathodic deposits changes from nickel through nickel (cobalt) - molybdenum (tungsten) alloys and intermetallics of different compositions to pure molybdenum (tungsten). Tab. 3, Fig. 7, Ref. 12.

Keywords: tungstate melt, molybdenum (tungsten) concentration, alloys, intermetallics, oxide tungsten melts

References

1. V.I. Shapoval, V.V. Soloviev, V.V. Malyshev. Electrochemically active species and multielectron processes in ionic melts. Russ. Chem. Rev., 2001, 70, 2, 161-176.
2. V.V. Malyshev, H.B. Kushkhov, V.I. Shapoval. High-temperature electrochemical synthesis of carbides, silicides, and borides of VI-B group metals in ionic melts. J. Appl. Electrochim., 2002, 32, 5, 573-579.
3. I.A. Novoselova, V.V. Malyshev, V.I. Shapoval, D.T. Supatashvili, Kh.B. Kushkhov. High-temperature electrochemical synthesis of binary and ternary intermetallic highmelting molybdenum and tungsten compounds with cobalt, nickel, and carbon. Russ. J. Appl. Chem., 1997, 70, 8, 1218-1224.
4. Eds. P. Kissinger, W.R. Heineman. Laboratory Techniques in Electroanalytical Chemistry, 1996, CRC Press.

5. C.H. Hamman. *Electrochemistry*, 1998, Wiley–VCH.
6. S. Glastone. *An Introduction to Electrochemistry*, 2006, Kindle Book.
7. J.F. Shackelford, M.K. Muralidhara. *Introduction to Materials Science for Engineers*, 2005, Pearson Int. Ed.
8. G.J. Janz. *Molten Salts Handbook*, 2013, Elsevier.
9. V.V. Malyshev, V.V. Soloviev, L.A. Chernenko, V.N. Rozhko. Management of composition cathodic products in the electrolysis of molybdenum-, tungsten- and carbon-bearing halogenide-ohide and oxide melts. *Mater. Sci. & Eng. Technol.*, 2015, 45, 11, 67-72.
10. K. Nitta, M. Majima, S. Inazawa, T. Nohira, R. Hagiwara. Electrodeposition of Tungsten from Li_2WO_4 – Na_2WO_4 – K_2WO_4 based melts. *Electrochim.*, 2009, 77, 8, 621-623.
11. W.D. Callister, D.G. Rethwisch. *Materials Science and Engineering*, 5, 2011, John Wiley & Sons.
12. D.A. Porter, K.E. Easterling, M. Sherif. *Phase Transformations in Metals and Alloys*, 2009, CRC Press.

1.4.1.4. Obtaining of new polytetrafluoroethylene based tribotechnical composite materials and studying their properties. /E. Kutelia, D. Gventsadze, L. Rurua/. *Nano Studies*. – 2019. – #19. – pp. 115-120. – geo.; abs.: geo.

Aim of the research was the obtaining the new polytetrafluoroethylene (PTFE) based composite materials by using natural mineral - chalcopyrite (CuFeS_2) - microparticles and studying their properties. On PTFE base, there were made polymeric composites filled with 2.5, 5, and 10 wt. % chalcopyrite powder. The wear intensity of these composites was determined at different friction speeds and results were compared with the characteristics of the best tribotechnical material - superfluvis. In case of further improving their tribological properties, we will get quite cheap and effective tribotechnical material. Tab. 1, Fig. 3, Ref. 4.

Keywords: polytetrafluoroethylene (PTFE), tribotechnicalcomposite, superfluvis, tribological properties

References

1. Д. Мур. *Основы и применения трибоники*, 2018, Москва, Мир, 122-125.
2. T. Xie, Zh.H. Zhou, Zh.X. Xu, J.W. Yu, M.H. Jiao. Characteristics of the transfer film and tribological properties of oxide / PTFE composites. *Adv. Mater. Res.*, 2013, 631-632, 172-175.
3. S.R. Yu, H.X. Hu, Y.B. Zhang, Y.H. Liu. Effect of transfer film on tribological behavior of polyamide 66-based binary and ternary nanocomposites. *Polymer Int.*, 2008, 57, 3, 454-462.
4. M.H. Cho. The role of transfer film and backtransfer behavior on the tribological performance of polyoxymethylene in sliding. *J. Mech. Sci. & Technol.*, 2009, 23, 8, 2291-2298.

1.4.1.5. Features of synthesis of gold nanoparticles using reagents of green chemistry. /T. Pavliashvili, E. Gelagutashvili, A. Tutunyan, G. Tservadze/. *Nano Studies*. – 2019. – #19. – pp. 121-124. – eng.; abs.: eng. Synthesis of gold nanoparticles from chloroauric acid using sodium citrate and tannin was studied. To prepare solutions, distilled water was used. The sodium citrate - tannin solution was prepared at room temperature, filtered and added with constant stirring to a 0.01% solution of HAuCl_4 . The optical absorption spectra of nanoparticles had an absorption maximum of 553 nm. The studies conducted with transmission electron microscopy showed that spheroidal and ellipsoid nanoparticles, as well as nano-triangles, were formed. The size of the formed gold nanoparticles turned out to be mainly in the range of 15-35 nm. Treatment of the solution in a microwave oven caused a decrease in the size of the nanoparticles and a shift in the optical absorption maximum towards the short waves (537 nm), apparently due to the further process of fragmentation of gold nanoparticles in the solution. Histograms were constructed for the distribution of gold nanoparticles in size before and after microwave exposure. Fig. 3, Ref. 5.

Keywords: gold nanoparticles, synthesis, green chemistry, histograms, solutions, microwave exposure

References

1. L.A. Dykman, N.G. Khlebtsov. Gold nanoparticles in biology and medicine: recent achievements and prospects. *Acta Naturae*, 2011, 3, 2 (9), 36-58.

2. T. Pavliashvili, T. Kalabegishvili, M. Janjalia, E. Ginturi, G. Tservadze. Synthesis of gold nanoparticles from chloroauric acid using red wine. *Eur. Chem. Bull.*, 2017, 6, 5, 192-195.
3. Yu.G. Frolov. A Course of Colloid Chemistry Surface Phenomena and Disperse Systems. 1989, Moscow, Khimya.
4. T. Pavliashvili, A. Tutunyan, Z. Akhvlediani, G. Tservadze. Nucleation of silicon dioxide nanoparticles in the film-forming tetraethoxysilane solution. *J. Pharm. & Appl. Chem.*, 2017, 3, 3, 185-187.
5. V.V. Makarov, A.J. Love, O.V. Sinitzyna, S.S. Makarova, I.V. Yaminsky, M.E. Taliantsky, N.O. Kalinina. "Green" nanotechnologies: Synthesis of metal nanoparticles using plants. *Acta Naturae*, 2014, 6, 1 (20), 35-44.

1.4.1.6. Effect of thermal treatment on hot-pressed isotopically modified boron carbide $^{10}\text{B}_4\text{C}$. /Z. Mestvirishvili, Sh. Zurabishvili, T. Mestvirishvili, Kh. Bluashvili, A. Sitchinava, N. Jalabadze/. *Nano Studies*. – 2019. – #19. – pp. 145-152. – rus.; abs.: rus.

Ceramic samples of $^{10}\text{B}_4\text{C}$ were obtained by hot pressing at temperature of 2100°C and pressure of 30 MPa. They were thermally treated in vacuum at 1800°C. Optical and scanning microscopic studies were carried out. The physical-mechanical properties of obtained ceramics were studied. The effect of thermal treatment on their machinability, fracture toughness, hardness, elastic modulus, as well as on the mechanical strength has been determined. Tab. 1, Fig. 5, Ref. 19.

Keywords: thermal treatment, boron carbide, ceramic samples, microscopic studies, fracture toughness, elastic modules

References

1. V. Domnich, S. Reynaud, R.A. Haber, M. Chhowalla. Boron carbide: Structure, properties, and stability under stress. *J. Am. Ceram. Soc.*, 2011, 94, 11, 3605-3628.
2. Z. Mestvirishvili, I. Bairamashvili, V. Kvatchadze, N. Rekhviashvili. Thermal and mechanical properties of $\text{B}_4\text{C}-\text{ZrB}_2$ ceramic composite. *J. Mater. Sci. & Eng. B*, 2015, 5, 9-10, 385-393.
3. K. Gillet, G. Roma, J.P. Crocombette, D. Gosset. The influence of irradiation induced vacancies on the mobility of helium in boron carbide. *J. Nucl. Mater.*, 2018, 512, 288-296.
4. Yu.A. Bykovskii, K.N. Zaitsev, P.J. Kervalishvili, I.N. Nikolaev, A.A. Portnov, S.O. Shalamberidze. Neutron fluence sensor based on boron carbide. *Tech. Phys. Lett.*, 1993, 19, 7, 457-458.
5. P.J. Kervalishvili, G.S. Karumidze, Sh.Sh. Shavelashvili, G.I. Kalandadze, S.O. Shalamberidze. Semiconductor sensor for neutrons. *Sensors & Actuators A*, 1993, 36, 43-45.
6. P.J. Kervalishvili, S.O. Shalamberidze, Yu.A. Bykovskii. Oriented boron carbide films produced by laser spraying. *Proc. Am. Inst. Phys.*, 1991, 524-527.
7. H. Werheit, M.H. Manghnani, U. Kuhlmann, A. Hushur, S. Shalamberidze. Mode Gruneisen parameters of boron carbide. *Solid State Sci.*, 2017, 72, 80-93.
8. A. Nino, A. Tanaka, Sh. Sugiyama, H. Taimatsu. Indentation size effect for the hardness of refractory carbides. *Mater. Trans.*, 2010, 51, 9, 1621-1626.
9. A.U. Khan, V. Domnich, R.A. Haber. Boron carbide-based armors: Problems and possible solutions. *Am. Ceram. Soc. Bull.*, 2017, 96, 6, 30-36.
10. В.Д. Рисованый, А.В. Захаров, Е.П. Клочков, В.Б. Пономаренко, Е.М. Мурашева, Т.М. Гусева. Поглощающие материалы стержней управления ядерных реакторов, 2012, Ульяновск, 57-95.
11. T.G. Abzianidze, A.M. Eristavi, S.O. Shalamberidze. Strength and creep in boron carbide (B_4C) and aluminum dodecaboride ($\alpha\text{-AlB}_{12}$). *J. Solid State Chem.*, 2000, 154, 1, 191-193.
12. G.I. Kalandadze, S.O. Shalamberidze, A.B. Peikrishvili. Sintering of boron and boron carbide. *J. Solid State Chem.*, 2000, 154, 1, 194-198.
13. H. Werheit, S. Shalamberidze. Advanced microstructure of boron carbide. *J. Phys. Cond. Matter*, 2012, 24, 38, 385-406.
14. T. Maruyama, Sh. Onose, T. Kaito, H. Horiuchi. Effect of fast neutron irradiation on the properties of boron carbide pellet. *J. Nucl. Sci. & Technol.*, 1997, 34, 10, 1006-1014.
15. Y. Morohashi, T. Maruyama, T. Donomae, Y. Tachi, Sh. Onose. Neutron irradiation effect on isotopically tailored $^{11}\text{B}_4\text{C}$. *J. Nucl. Sci. & Technol.*, 2008, 45, 9, 867-872.

16. I.A. Bairamashvili, M.V. Galustashvili, J.Sh. Jobava, V.G. Kvatchadze, Z.Z. Mestvirishvili. Composite ceramics based on boron carbide enriched in isotope ^{11}B as a promising radiation resistant structural material. *Nano Studies*, 2013, 8, 305-310.
17. T. Jianga, H. Jinb, Zh. Jina, J. Yanga, G. Qiaoa. An investigation of the mechanical property and thermal shock behavior of machinable $\text{B}_4\text{C}/\text{BN}$ ceramic composites. *J. Ceram. Proc. Res.*, 2009, 10, 1, 113-116.
18. B.M. Moshtaghoun, A.L. Ortiz, D. Gomez-Garcia, A. Dominguez-Rodriguez. Toughening of super-hard ultra-fine grained B_4C densified by spark-plasma sintering via SiC addition. *J. Eur. Ceram. Soc.*, 2013, 33, 1395-1401.
19. G.R. Anstis, P. Chantikul, B.R. Lawn, D.B. Marshall. A critical evaluation of indentation techniques for measuring fracture toughness – I. Direct crack measurements. *J. Am. Ceram. Soc.*, 1981, 64, 533-538.

1.4.1.7. A study of the condensed copper-containing nanomaterials. /D. Jishiashvili, Z. Shiolashvili, N. Makhatadze, A. Jishiashvili, A. Chirakadze, V. Gobronidze/. *Nano Studies*. – 2019. – #19. – pp. 285-290. – eng.; abs.: eng.

The development of new technologies for the vapor phase growth of copper-based materials is an important issue, which enables the formation of nano-sized structures. In this work, we have demonstrated that volatile Cu-containing molecules can be formed directly in the reaction zone by annealing Cu or CuO solid sources in the gaseous medium comprising ammonium chloride and hydrazine. We suggest that at the first stage of synthesis the CuCl volatile molecules were formed, followed by a Cu mass transfer to the Si substrate located above the source. In the case of Cu source, after reaching the Si substrate the CuCl molecules were reduced by hydrogen forming crystalline Cu microstructures. When CuO source was annealed at 700°C , the copper oxide nanostructures were produced on Si due to the presence of excess oxygen, which was released from CuO source. Fig. 4, Ref. 15.

Keywords: copper-based nanomaterials, nano-sized structures, copper-oxide nanostructures, excess oxygen

References

1. Ed.M. Seehra. Copper Metal for Semiconductor Interconnects. In: Noble and Precious Metals. Properties, Nanoscale Effects and Applications, 2018, IntechOpen. 430 pp.
2. M.R. Baklanov, Ch. Adelmann, L. Zhao, S. De Gendt. *ECS J. Solid State Sci.*, 2015, 4, Y1.
3. R. Motoyoshi, T. Oku, A. Suzuki, K. Kikuchi, Sh. Kikuchi, B. Jeyadevan, J. Cuya. *Adv. Mater. Sci. & Eng.*, 2010, 1-11.
4. P.G. Gordon, A. Kurek, S.T. Barry. *ECS J Solid State Sci. & Technol.*, 2015, 4, 3188.
5. P.C. Andricacos, C. Uzoh, J.O. Dukovic, J. Horkans, H. Deligianni. *IBM J. Res. & Develop.*, 1998, 48, 567.
6. Y. Sui, Y. Zhang, W. Fu, H. Yang, Q. Zhao, P. Sun, D. Ma, M. Yuan, Y. Li, G. Zou. *J. Cryst. Growth*, 2009, 311, 2285.
7. D. Ahn, S.-H. Park. *Sci. Rep.*, 2016, 6, 20718.
8. W.A. Herrmann. Synthetic Methods of Organometallic and Inorganic Chemistry: Copper, Silver, Gold, Zinc, Cadmium and Mercury, 1996, Georg Thieme Verlag, 248 pp.
9. K.H. Park, W.J. Marshall. *J. Am. Chem. Soc.*, 2005, 127, 9330.
10. R. Claessen. Design of Volatile Non-Halogenated Precursors for the Chemical Vapor Deposition (CVD) of Copper: Understanding Key Molecular Properties in a CVD Process Feasibility Study, 2002, Books on Demand, 212 pp.
11. D. Jishiashvili, Z. Shiolashvili, A. Chirakadze, N. Makhatadze, V. Gobronidze, A. Jishiashvili. *Nano Studies*, 2018, 17, 67.
12. Ed.A.R. Barron. Chemistry of Electronic Materials. Connexions, 2012, Houston, Rice University, 359 pp.
13. Z. Wang, G. Marin, G.F. Naterer, K.S. Gabriel. *J. Therm. Anal.*, 2015, 119, 815. 14. M. Burgelman, A. de Vos. *Thin Solid Films*, 1983, 102, 367.
15. C. Zamfirescu, I. Dincer, G.F. Naterer. *Int. J. Hydrogen Energy*, 2010, 35, 4839.

1.4.1.8. Metal-containing nanoparticles in a maleinized polyethylene matrix as nanofillers for polypropylene. /N. Kurbanova, Z. Huseynova, N. Ishenko, A. Aliyev, T. Guliyeva, S. Ragimova, S. Rzaeva/. *Azerbaijan Chemical Journal*. – 2019. – #1. – pp. 54-58. – eng.; abs.: rus., az., eng.

The effect of additives of metal-containing nanofillers including nanoparticles of zinc oxide stabilized on a polymer matrix of maleinized polyethylene on the properties of nanocomposites based on isotactic polypropylene by methods X-ray phase and thermogravimetric analyses was studied. The improvement of

the strength and rheological indexes, as well as the thermooxidative stability of the nanocomposites obtained was revealed. Fig. 2, Tab. 2, Ref. 11.

Keywords: isotactic polypropylene, metal-containing nanofillers, maleinized polyethylene matrix, physical and mechanical properties

References

1. Suzdalev I.P., Suzdalev P.I. Nanoclusters and nanocluster systems // Adv. Chem. 2001. V. 70. No 3. P. 203-240.
2. Pomogailo A.D., Rosenberg A.S., Uflyand I.E. Nanoparticles of metals in polymers. M.: Chemistry. 2000. 672 p.
3. Tretyakov A.O. Polymer nanocomposites – materials of the XXI century. Equipment and tools for professionals. 2003. V. 37. No 2. P. 18-20.
4. The encyclopedia of polymers. M.: The Soviet Encyclopedia. 1974. V. 2. P. 328.
5. Foster L. Nanotechnology. Science, innovation and opportunities. M.: Technosphere, 2008. 352 p.
6. Antipov E.M., Barannikov A.A., Gerasin V.A. et al. The structure and deformation behavior of nanocomposites based on polypropylene and modified clays // Highmol. comp., 2003. A. V. 45. No 11. P. 1885-1898.
7. Kovalchuk A.A., Shchegolikhin A.N., Dubnikova I.L. et al. Nanocomposites polypropylene/ multiwalled carbon nanotubes: in situ polymerization and properties // Plast. masses. 2008. № 5. P. 27-30.
8. Sister V.S., Ivannikova E.A., Lomakin S.M. Comparative analysis of thermal stability of polymer nanocomposites based on polypropylene // Nanomater. nanotechnol. 2012. № 3. P. 13-28.
9. Gubin S.P., Yurkov G.Yu., Kosobudsky I.D. Nanomater. based on metal-containing nanoparticles in polyethylene and other carbon-chain polymers // Int. J. Mater. Product Technol. 2005. V. 23. No 1–2. P. 2-25.
10. Kurbanova N.I., Alimirzoeva N.A., Guseinova Z.N., Nurullayeva D.R. Ecological Method of Preparation of Metal-Containing Nanoparticles in Polyethylene Matrix // 3st Int. Turkic World Conf. Chem. Sci. and Technol. (ITWCCST 2017). Baku. 10-13 Sept. 2017. Abstract. P. 98.
11. Technic. properties of polymeric materials: Educational-reference manual // Under the general ed. prof. Kryzhanovsky V.K. SPb.: The profession, 2007. 240p.

1.4.1.9. SmS thin films with nanosize surface architecture /Y. Aliyeva/. Azerbaijan Journal of Physics. –2019. – v. 25. – #3. – pp. 47-51 – eng.; abs.: eng.

Samarium sulfide thin films with thickness in the range from 40 to 100 nm were deposited on vacuum annealed sapphire plates that received ion beam treatment after annealing to reduce surface roughness. The deposited thin films were then subjected to vacuum annealing and ion beam treatment again and films surfaces good enough to use scribing to create nanosize surface elements were obtained. Diffraction grating was then prepared with the aid of a diamond coated cantilever and examined using electric force (EF) and magnetic force (MF) modes of atomic force microscope (AFM). The prepared grating was shown to have semiconducting channels between the metallic grooves separated from one another by 100 nm distance. The disclosed distinct properties of the grating are discussed in terms of variable valence of Sm ions. It is figured out that SmS thin films are very promising as materials to scribe on and obtain desirable surface architecture. Fig. 4, Ref. 10.

Keywords: Atomic Force Microscopy, SmS polycrystalline film, nanoscale structures, diffraction grating, Magnetic Force Microscopy, Electric Force Microscopy

References

1. V.V. Kaminskii, Shinji Hirai, Toshihiro Kuzuya, S.M. Solovyev, N.N. Stepanov, N.V. Sharenkova. Obtaining of SmS based semiconducting material and investigation of its electrical properties, Physics of the Solid State , 47, 2013, 1310-1313.
2. V.N. Antonov, A.P. Shpak, A.N. Yaresko. Electronic structure of mixed valent systems, Condensed Matter Physics 7, 2004, 211-246.
3. V.A. Sidorov, N.N. Stepanov, L.G. Khvostantsev, O.B. Tsiok, A.V. Golubkov, V.S. Oskotski and I.A. Smirnov. Intermediate valency state of samarium chalcogenides under high pressure, Semicond. Sci. Technol. 4, 1989, 286.
4. E. Rogers, P.F. Smet, P. Dorenbos, D. Poelman and E. van der Kolk. The thermally induced metal-semiconducting phase transition of samarium monosulfide (SmS) thin films, J.Phys. Condens Matter 22(1), 2010, 015005.

5. V.V. Kaminski and M.M. Kazanin. Thermo-voltaic Effect in Thin-Film Samarium-Sulfide-Based Structures, Technical Physics Letters 34, 2008, 361-362.
6. A. Monniera and M. Schnieper, R. Jaaniso, thin films of the Matlockite structure: Design, luminescence and hole-burning experiments, J. Appl. Phys. 82 (2), 1997, 536-547.
7. X.N. Xie, H.J. Chung, C.H. Sow, A.T.S. Wee. Nanoscale materials patterning and engineering by atomic force microscopy nanolithography, Materials Science and Engineering, R 54, 2006, 1-48.
8. Jung Yup Yang, Joo Hyung Kim, Jun Seok Lee, Seung Ki Min, HyunJung Kim, Kang L. Wang JinPyo Hong. Electrostatic force microscopy measurements of charge trapping behavior of Au nanoparticles embedded in metal insulator semiconductor structure, Ultramicroscopy 108, 2008, 1215-1219.
9. P.P. Deen, D. Braithwaite, N. Kernavanois, L. Paolasini, S. Raymond, A. Barla, G. Lapertot, and J. P. Sanchez. Structural and electronic transitions in the low-temperature, high-pressure phase of SmS, Phys. Rev. B, 71 245118, 2005, pp. 55-58.
10. E. Beaurepaire. Magnetism: A Synchrotron Radiation Approach, Springer Science & Business Media, 2006, 468 [11] U. Hartmann, Magnetic force microscopy, Annu. Rev. Mater. Sci. 29, 1999, 53-87.

5. NANOENGINEERING

5.1. Devices and Sensors

1.5.1.1. Highly photosensitive and radiant energy-saving molecular switches. /L. Devadze, J. Maisuradze, G. Petriashvili, Ts. Zurabishvili, N. Sepashvili, Sh. Akhobadze/. Nano Studies. – 2019. – #19. – pp. 97-102. – eng.; abs.: eng.

The work is based on an increase of effective photosensitivity of some spiropyrans (spirochromens) in a liquid-crystal matrix, which was discovered and studied by authors previously. The merocyanine molecules formed as a result of photoinduction with the ultraviolet (UV) light are structured as nano-particles, the micelles. The originated micelles deplete the base solution off merocyanine molecules and spiropyran molecules start to transform into the merocyanine molecules to restore the thermodynamic equilibrium. The merocyanine molecules formed photochemically are added by the molecules formed to restore thermodynamic equilibrium. The coloration of the system continues at the expense of the internal energy without consuming the radiant energy. By doping the composition in the polymer by using the innovative method developed by us, we obtained highly photosensitive and energy-saving polymer films with the ability of photo-regulation of their optical properties. Such an economy of the electrical power is quite significant when recording and reading large information massifs. Therefore, we consider it purposeful to use such polymer films as a photosensitive system in modern computers recording and processing the data, molecular machines, etc. Fig. 6, Ref. 13.

Keywords: molecular switches, photosensitivity, ultraviolet light, thermodynamic equilibrium, molecular machines

References

1. K. Japaridze, L. Devadze, J. Maisuradze, G. Petriashvili, Ts. Zurabishvili, I. Mzhavanadze, N. Sepashvili. Bull. Georgian Natl. Acad. Sci., 2013, 7, 57.
2. R. Klajn. Chem. Soc. Rev., 2014, 43, 148.
3. D.A. Parthenopoulos, P.M. Rentzepis. Science, 1989, 24, 842.
4. J.N. Israelachvili, D.J. Mitchell, B.W. Ninham. J. Chem. Soc. Faraday Trans. 2, 1976, 72, 1525.
5. R. Klajn, J. Fraser Stoddart, B.A. Grzybowski. Chem. Soc. Rev., 2010, 39, 2203.
6. K. Japaridze, Ts. Zurabishvili, G. Petriashvili. Georgian Patent P 5232, 2011.
7. V.A. Barachevsky. Rev. J. Chem., 2017, 7, 3, 334.
8. V.A. Barachevsky, T.M. Valova. Optics & Spectroscopy, 2017, 123, 404.
9. W. Tian, J. Tian. Dyes Pigm., 2014, 105.
10. K.G. Japaridze, L.V. Devadze, J.P. Maisuradze, N.O. Sepashvili, G.Sh. Petriashvili. In: Abs. 7th All-Russ. Conf. "Molecular Modeling", 2011, Moscow, 30.
11. J. Buback, M. Kullmann, F. Langhofer, P. Nuernberger, R. Schmidt, F. Wurthner. J. Am. Chem. Soc., 2010, 111, 132.
12. K.G. Japaridze, J.P. Maisuradze, G.G. Gachechiladze, E.S. Gomelauri. Chem. Heterocyclic Comp., 1971, 6, 775.
13. G. Petriashvili, M. Penelope de Santo, L. Devadze, Ts. Zurabishvili, N. Sepashvili, R. Gary, R. Barberi. Macromol. Rapid Commun., 2016, 37, 500.

1.5.1.2. Thin-film Sol transistor as base of generation of smart nanosensors of external influences with improved properties. /A. Leonov, V. Mordkovich, M. Pavlyuk/. Nano Studies. – 2019. – #19. – pp. 103-110. – eng.; abs.: eng.

It was demonstrated that magnetosensitive Sol (silicon-on-insulator) thin film transistors (TFT) with built in channel, double gates MOSOM (metal-oxide-Si-oxide-metal) control system, and opposite ohmic contacts to lateral channel sides (so called field effect Hall sensor - FEHS) provide the creation of smart magnetic nanoelectronic sensors with radically improved characteristics in comparison with conventional Si Hall elements. In particular FEHS improved magnetosensitivity operating temperature, radiation immunity, and power consumption. Such advantages are due of FEHS construction and Sol-technology for its manufacture. It was also demonstrated that SOI MOSOM TFT with built-in channel may be the attractive sensitive

elements for sensors of other external influences such as temperature, radiation, atmospheric gases composition, etc. Fig. 6, Ref. 18.

Keywords: Sol transistor, smart nanosensors, thin film transistors (TFT), Hall effect sensor (HES), radiation, gases, external influences

References

1. I.H. Huijsing. Smart Sensor System: Why? Where? Now? In: Smart Sensor System (Ed. G. Meiyer), 2012, Wiley & Sons Ltd.
2. Magnetic Field Sensor Market worth 2,9 billion USD.
3. Magnetic Sensors Market Analysis by Technology (February 2016).
<http://www.grandviewresearch.com/industry-analysis/magnetic-sensors-market>
4. R.S. Popovic. Hall effect devices, 2003, CRC Press.
5. М.И. Павлюк. Микроэлектронные датчики с расширенным функционалом на основе тонкоплёночного КНИ МДПДМ-транзистора со встроенным каналом. Электроника НТБ, 2017, 5, 70.
6. M.I. Kakoulin, A.V. Leonov, A.A. Malykh, V.N. Mordkovich, A.B. Odnolko, M.I. Pavlyuk. The development of constructive-technological decisions on creation of a family of microelectronic elements on the "silicon on insulator" (SOI) structures to provide the ability to create sensors of external influences of a various functional purpose. In: Proc. Sci.-Prac. Conf. "R&D – 2016", 2018, Springer Open, 31.
7. J.P. Colinge. Silicon-on-Insulator Technology: Materials to VLSI, 2004. Springer Science & Business Media.
8. A.V. Leonov, A.A. Malykh, V.N. Mordkovich, M.I. Pavlyuk. Field controlled Si Hall element with extended operation temperature range from liquid helium temperature up to 650 K. Procedia Eng., 2015, 120, 1197.
9. С. Зи. Физика полупроводниковых приборов, 1, 1984, Мир.
10. В.Л. Винецкий, Г.А. Холодарь. Радиационная физика полупроводников, 1979, Наукова думка.
11. К.Д. Щербачев, В.Т. Бублик, В.Н. Мордкович, Д.М. Пажин. Особенности образования радиационных дефектов в слое кремния структур «кремний на изоляторе». ФТП, 2011, 45, 754.
12. Д.М. Пажин. Влияние облучения на свойства КНИ структур и полевых элементов со встроенным каналом на их основе (Дисс. канд. физ.-мат. наук), 2013, ИПТМ РАН.
13. В.С. Першенков, В.Д. Попов, Г.М. Шальнов. Поверхностные радиационные эффекты в элементах интегральных микросхем, 1998, Энергоатомиздат.
14. Д.В. Громов, В.Н. Мордкович, Д.М. Пажин, П.К. Скоробогатов. Релаксационные эффекты в полевых датчиках Холла приводействии импульса ионизирующего облучения. Электронная техника (Сер. 2: Полупроводниковые приборы), 2011, 1, 19.
15. М.Л. Бараночников, А.В. Леонов, А.А. Малых, В.Н. Мордкович, В.Н. Мурашев, Д.М. Пажин. Датчики внешних воздействий с частотным выходом на основе полевого МДПДМ-транзистора со встроенным каналом. Нано- и микросистемная техника, 2013, 10 (159), 8.
16. A.V. Leonov, M.I. Pavlyuk. A stabilizer of micro- and small currents based on a field hall-effect sensor with autocompensation of the temperature effect. Instr. & Exp. Tech., 2016, 59, 721.
17. A.V. Leonov, A.A. Malykh, V.N. Mordkovich, M.I. Pavlyuk. Multichannel and multifunctional frequency-output sensors of physical effects based on a universal field transistor-type sensing element with a silicon-on-insulator structure. Instr. & Exp. Tech., 2018, 61, 299.
18. М.Л. Бараночников, А.В. Леонов, А.А. Малых, В.Н. Мордкович, В.Н. Мурашев, Д.М. Пажин. Датчики внешних воздействий с частотным выходом на основе полевого МДПДМ-транзистора со встроенным каналом. Нано- и микросистемная техника, 2013, 10 (159), 8.

1.5.1.3. Single channel laser rangefinder with aspheric concentrating mirror. /T. Khachidze, D. Zardiashvili, G. Mcchedlishvili/. Nano Studies. – 2019. – #19. – pp. 299-302. – geo.; abs.: geo.

The article refers to laser rangefinder (LRF) design issues. Single channel LRF with punctured aspheric mirror is proposed, which transfers laser beam to objective lens and transmits focused and reflected light to photodiode with minimal loses at the same time. Apart from that aspheric surface focuses light unlike flat one and makes it possible to receive laser beam spot of smaller size on photodiode. By changing the focal length of the mirror, it becomes possible to vary distance to photodiode without changing the focal length of the objective lens. Fig. 2, Ref. 4.

Keywords: laser rangefinder (LRF), design, punctured aspheric mirror, focal length of the mirror, objective lens

References

1. В.Б. Бокшанский, Д.А. Бондаренко, М.В. Вязовых, И.В. Животовский, А.А. Сахаров, В.П. Семенков. Лазерные приборы и методы измерения дальности (Учебное пособие), 2012, Москва, Изд. МГТУ им. Н.Э. Баумана, 92 стр.
2. И.Ф. Балашов. Энергетическая оценка импульсных лазерных дальномеров, 2000, Санкт-Петербург, ГИТМО, 37 стр.
3. А. Медведев, А. Гринкевич, С. Князева. Однозрачковые системы со встроенным лазерным дальномером. Фотоника, 46, 4, 50-60.
4. А.В. Медведев, А.В. Гринкевич. Однозрачковая мультиспектральная оптическая система со встроенным лазерным дальномером. Охранный документ № 0002646436 (05.03.2018).

1.5.1.4. Study of hydrogen peroxide vapors sensors made on the base of carbon nanotubes coated with tin oxide nanoparticles. /Z. Adamyan, A. Sayunts, E. Khachaturyan, V. Arakelyan, V. Aroutiounian, B. Joost/. Proceedings of NAS RA. Physics. – 2019. – v. 54. – #1. – pp. 75-84. – rus.; abs.: rus., arm., eng.

In this work, we present the results of studies of the nanocomposite MWCNTs/SnO₂ hydrogen peroxide vapor sensors. The technology of manufacturing of these sensors has been developed. 100°C optimal operating temperature of the studied sensors has been found as a result of the measurements of the temperature characteristics. The response and recovery curves of the sensors were investigated in the presence of different concentrations of hydrogen peroxide vapor in the atmosphere. Sufficiently high response is observed when low concentrations of the target gas present in the air. The linear dependence of the response of the sensor on the concentration of hydrogen peroxide vapor is observed in a double logarithmic scale in a certain concentration range. The minimal registered gas concentration is 1 ppm or less. Fig. 5, Tab. 2, Ref. 23.

Keywords: hydrogen peroxide vapor sensor, nanocomposites, temperature characteristics, gas concentration

References

1. W.T. Hess, Hydrogen Peroxide in Kirk-Othmer Encyclopedia of Chemical Technology, 13, Wiley, New York, 1995, p. 961.
2. I. Taizo, A. Sinichi, K. Kawamura. PDAJ. Pharm. Sci. Tech, 52, 13 (1998).
3. F. Dubnikova, R. Kosloff, et al. J. Amer. Chem. Soc., 27, 1146 (2005).
4. S. Banerjee, S.K. Mohapatra, et al. Nanotechnology, 20, 075502 (2009).
5. W. Xu, Y. Fu, Y. Gao, et al. Chem. Commun., 51, 10868 (2015).
6. S. Reisert, B. Schneider, et al. Phys. Stat. Sol. A, 210, 5, 898 (2013).
7. S. Reisert, H. Geissler, et al. Phys. Stat. Sol. A, 208, 6, 1351 (2011).
8. V.M. Aroutiounian, V. Araqelyan, M. Aleksanyan, et al. JSSS, 7, 281 (2018).
9. D.-J. Lee, S.-W. Choi, Y.T. Byun. Sens. Actuators B, 256, 744 (2018).
10. A.L. Verma, S. Saxena, G.S.S. Saini, et al. Thin Solid Films, 519, 8144 (2011).
11. F.I. Bohrer, C.N. Colesniuc, J. Park, et al. J. Amer. Chem. Soc., 130, 3712 (2008).
12. V.M. Aroutiounian, A.Z. Adamyan, et al. Sens. Actuators B, 177, 308 (2013).
13. V. Aroutiounian, Z. Adamyan, A. Sayunts, et al. IJETST, 1, 8, 1309 (2014).
14. Z. Adamyan, A. Sayunts, V. Aroutiounian, et al. Sensors & Transducers, 213, 6, 38 (2017).
15. З.Н. Адамян, А.Г. Саюնц, и др. Изв. НАН РА, Физика, 51, 192 (2016).
16. N. Sinha, J. Ma, J.T.W. Yeow. J. Nanoscience and Nanotechnology, 6, 573 (2006).
17. R. Leghrib, A. Felten, J.J. Pireaux, E. Llobet. Thin Solid Films, 520, 966 (2011).
18. M. Penza, R. Rossi, M. Alvisi, et al. Sens. Actuators B, 135, 289 (2008).
19. M. Penza, G. Cassano, R. Rossi, et al. Appl. Phys. Lett., 90, 173123-1 (2007).
20. M.S. Wagh, G.H. Jain, D.R. Patil, et al. Sens. Actuators B, 122, 357 (2007).
21. Z. Adamyan, A. Sayunts, V. Aroutiounian, et al. JSSS, 7, 31 (2018).

22. Z. Nemeth, B. Reti, Z. Pallai, Phys. Stat. Sol. (B) Basic Research, 251, 2360 (2014).
 23. Z. Nemeth, Z. Pallai, B. Reti, et al. J. Coating Sci. Tech., 1, 137 (2014).

1.5.1.5. Concept of an optical magnetometer based on the spectroscopy of alkali vapors confined in nanometric-thick cells. /E. Klinger/. Proceedings of NAS RA. Physics. – 2018. – v. 53. – #4. – pp. 417-431. – rus.; abs.: rus., eng.

In this article, we present the concept of an optical scalar magnetometer based on the spectroscopy of hot alkali vapors confined in nanometric-thick cells. We present an appropriate theoretical model to describe the interaction of linearly and circularly polarized light with atomic alkali vapors confined in extremely thin cells where a longitudinal magnetic field is applied. This model can be used to perform consecutive fittings of experimental spectra recorded by derivative selective reflection method, in order to measure the value of magnetic field. We illustrate the model with various calculated spectra for natural Rb vapor, while equivalent results hold for other alkalis (Na, K, and Cs). We analyze the feasibility of the concept for different cases depending on light polarization and cell thickness, and discuss possible limitations of the technique. Fig. 5, Ref. 17.

Keywords: optical magnetometer, spectroscopy, alkali vapors, nanometric-thick cells, feasibility, Rb vapor, light polarization, cell thickness

References

1. W. Demtröder. Laser spectroscopy: basic concepts and instrumentation. *Laser Spectroscopy: Basic Concepts and Instrumentation*, Berlin: Springer-Verlag, 2013.
2. W. Yang, D.B. Conkey, B. Wu, D. Yin, A.R. Hawkins, H. Schmidt. *Nature Photonics*, 1, 331 (2007).
3. D. Budker, D.F.J., Kimball. *Optical magnetometry*. Cambridge University Press, 2013.
4. P. Tremblay, A. Michaud, M. Levesque, S. Thériault, M. Breton, J. Beaubien, N. Cyr. *Phys. Rev. A*, 42, 2766 (1990).
5. A. Sargsyan, G. Hakhamyan, A. Papoyan, D. Sarkisyan, A. Atvars, M. Auzinsh, *Appl. Phys. Lett.*, 93, 021119 (2008).
6. A. Sargsyan, E. Klinger, G. Hakhamyan, A. Tonoyan, A. Papoyan, C. Leroy, D. Sarkisyan. *JOSA B*, 34, 776 (2017).
7. А. Саргсян, Э. Клингер, Е. Пашиян-Леруа, К. Леруа, А. Папоян, Д. Саркисян. *Письма в ЖЭТФ*, 104, 222 (2016).
8. Э. Клингер, А. Саргсян, К. Леруа, Д. Саркисян. *ЖЭТФ*, 152, 641 (2017).
9. D.A. Steck. Rubidium 85 D line data, <http://steck.us/alkalidata/rubidium85numbers.pdf>
10. D.A. Steck. Rubidium 87 D line data, <http://steck.us/alkalidata/rubidium87numbers.pdf>
11. C. Cohen-Tannoudji, B. Diu, F. Laloë. *Quantum Mechanics, Volume I*, ISBN 978-0-471-16433-3, Wiley-VCH, 830 (2005).
12. B. Zambon, G. Neinhuis. *Opt. Commun.*, 143, 308 (1997).
13. G. Dutier, S. Saltiel, D. Bloch, M. Ducloy. *JOSA B*, 20, 793 (2003).
14. L. Weller, K.S. Kleinbach, M.A. Zentile, S. Knappe, C.S. Adams, I.G. Hughes. *J. Physics B: Atomic, Molecular and Optical Physics*, 45, 215005 (2012).
15. A. Sargsyan, G. Hakhamyan, C. Leroy, Y. Pashayan-Leroy, A. Papoyan, D. Sarkisyan. *Optics Letters*, 37, 1379 (2012).
16. K.A. Whittaker, J. Keaveney, I.G. Hughes, A. Sargsyan, D. Sarkisyan, B. Gmeiner, V. Sandoghdar, C.S. Adams. *J. Physics: Conference Series*, 635, 122006 (2015).
17. M.T. Hummon, S. Kang, D.G. Bopp, Q. Li, D.A. Westly, S. Kim, C. Fredrick, S.A. Diddams, K. Srinivasan, V. Aksyuk, J.E. Kitching. *Optica*, 5, 443 (2018).

1.5.1.6. Plasmonic Nanoparticles Arrangements for Biosensing (Proceedings of the Int. Conference on “Microwave and THz Technologies and Wireless comm.”). /H. Parsamyan, T. Yezekyan, H. Haroyan/. Armenian Journal of Physics. – 2018. – v. 11. – #4. – pp. 241-245. – rus.; abs.: rus., arm., eng.

The localized surface plasmon resonance (LSPR) of gold and silver spherical and cubic nanoparticles (NP), as well as NP dimers is studied. We use numerical simulations based on finite element method to derive extinction cross sections (ECS) and LSPR wavelengths of the investigated structures. Study of the dependence of the LSPR of the cubic NPs on the different values of chamfering radius shows the red shift of the resonant wavelength while decreasing the chamfering radius. For the silver nanocube the analysis

suggests three extinction peaks on ~ 333 nm, ~ 364 nm due to strong absorption and on ~ 400 nm due to scattering, while for gold nanocube there is only one extinction peak on ~ 526 nm. The diameter of the spheres and the sides of the cubes is 50 nm. According to the simulations the localized surface plasmon properties of the silver nanoparticles are much stronger and more diverse than that of corresponding gold nanoparticles. The study of nanocube and nanosphere dimers is carried out as well. Fig. 3, Ref. 10.

Keywords: surface plasmon, plasmonic nanoparticles, biosensing, finite element method.

References

1. S. Unser, I. Bruzas, Jie He and L. Sagle. "Localized surface plasmon resonance biosensing: current challenges and approaches Sensors", vol. 15, no 7, pp. 15684-15716, Jul. 2015.
2. M. Hentschel, et al. "Linear and nonlinear optical properties of hybrid metallic-dielectric plasmonic nanoantennas", Beilstein J Nanotechnol, vol. 7, pp. 111-120, Jan. 2016.
3. I.I. Slowing, Juan L. Vivero-Escoto, Chia-Wen Wu, Victor S.-Y. Lin. "Mesoporous silica nanoparticles as controlled release drug delivery and gene transfection carriers," Advanced drug delivery reviews, vol. 60, no. 11, pp. 1278-1288, Aug. 2008.
4. I. Brigger, C. Dubernet, P. Couvreur. "Nanoparticles in cancer therapy and diagnosis", Advanced drug delivery reviews, vol. 60, pp. 24-36, Dec. 2012.
5. V. Amendola, et al. "Surface plasmon resonance in gold nanoparticles: a review", Journal of Physics: Condensed Matter, vol. 29, no. 20, pp. 203002, Apr. 2017.
6. M. Jahn, et al. "Plasmonic nanostructures for surface enhanced spectroscopic methods", Analyst, vol. 141, no. 3, pp. 756-793, Dec. 2015.
7. S. Chen, M. Svedendahl, M. Käll, L. Gunnarsson and A. Dmitriev. "Ultrahigh sensitivity made simple: nanoplasmonic label-free biosensing with an extremely low limit-of-detection for bacterial and cancer diagnostics", Nanotechnology, vol. 20, no. 43, pp. 434015, Oct. 2009.
8. T.K. Sau and C.J. Murphy. "Room temperature, high-yield synthesis of multiple shapes of gold nanoparticles in aqueous solution", Journal of the American Chemical Society, vol. 126, no. 28, pp. 8648-8649, Jul. 2009.
9. A.S.D.S. Indrasekara, S. Meyers, S. Shubeita, L.C. Feldman, T. Gustafsson and L. Fabris. "Gold nanostar substrates for SERS-based chemical sensing in the femtomolar regime", Nanoscale, vol. 6, no. 15, pp. 8891-8899, May. 2014.
10. M. Thiele, A. Knauer, D. Malsch, A. Csáki, Th. Henkel, J.M. Köhler and W. Fritzsch. "Combination of microfluidic high-throughput production and parameter screening for efficient shaping of gold nanocubes using Dean-flow mixing", Lab on a Chip, vol. 17, no. 8, pp. 1487-1495, May. 2017.

1.5.1.7. Generation of Optical Radiation in Nanotube Undulators. /K. Gevorgyan, L. Gevorgyan/. Armenian Journal of Physics. – 2018. – v. 11. – #3. – pp. 125-129. – eng.; abs.: eng.

The possibility to generate an sufficiently intense, monochromatic and directed radiation in the near ultraviolet range was investigated. Such radiation has important practical application in biology and medicine. It is shown that for a certain choice nanotube parameters channeled relativistic positrons in it the bunch will radiate as gamma photons and the "water window" photons also. Ref. 12.

Keywords: tunable radiation, channeling, nanotube undulator, positron bunch, medium polarization, water window

References

1. V.L. Ginzburg. "On radiation of micro-radio-waves and their absorption in air", Izvestiya Akad. Nauk SSSR, Ser. Fiz. 11(1947) 165-182.
2. N.A. Korhmazyan. "Radiation of Fast Charged Particles in Transverse Sinusoidal Electric Fields (in Russian)", Izv. Akad. Nauk Arm. SSR, Ser. Fiz. 5(1970) 287-8 & 418.
3. H. Motz. "Cherenkov effect and undulator radiation", J. Appl. Phys. 22(1951) 527.
4. L.A. Gevorgyan, N.A. Korkhmazyan. "Undulator radiation in dispersive media", JETP 76(1979) 1226-1235.
5. J. Lindhard. "The influence of crystal lattice on the motion of fast charged particles", Physcis Uspekhi 99(2) N2(1969) 249-296.
6. M.A. Kumakhov. "On the theory of electromagnetic radiation of charged particles in crystal", Phys. Lett. 57A(1976) 17-18.

7. V.V. Kaplin, S.V. Plotnikov, S.A. Vorobev. "Radiation by charged particles in deformed crystal", *Zh. Tekh. Fiz.* 50(1980) 1079-1081.
8. A.A. Korol, A.V. Solov'yov, W. Greiner. "Coherent radiation of an ultrarelativistic charged particle channeled in a periodically bent crystal", *J. Phys. G: Nucl. Part. Phys.* 24 (5) (1998)L45-L53.
9. R.O. Avakian, L.A. Gevorgian, K.A. Ispirian, R.K. Ispirian. "Spontaneous and stimulated radiation of particles in crystalline and nanotube undulators", *NIM B* 173(2001) 112-120.
10. L.A. Gevorgian, K.A. Ispirian, R.K. Ispirian. "High energy particle channeling in nanotubes", *NIM B* 145(1998) 155-159
11. K.L. Gevorgyan, L.A. Gevorgyan, H.L. Gevorgyan. Positron Bunch Radiation in the System of Tightly-Packed iv:1512.08282.
12. N. Rohriger et al. *Nature*, 481 (2012) 488.

6. NANOMEDICINE

6.1. Medical Physics

1.6.1.1. Actual problems of boron-neutron-capture-therapy and prospects for its implementation in Georgia. /D. Arutinovi/. Nano Studies. – 2019. – #19. – pp. 153-202. – geo.; abs.: geo.

On the one hand, the increase in mortality in the World due to tumors and, on the other hand, fast development of particles' accelerators allows us to take a fresh look at the issue of increasing the effectiveness of the use of various radiation therapies for providing high-tech care to oncological patients. At the same time, it should be taken into account that, according to modern ideas, the presence of a malignant tumor is so-called systemic disease and methods of treatment of standard-type tumors are different for disseminated and systemic diseases, as well as for locally distributed processes of similar diseases. Radiation therapy with this or that radiation (X- or gamma-rays, electrons, protons, neutrons, ions, etc.) acts directly on the tumor. Among the methods of clinical oncology, radiation therapy occupies one of the leading places. According to the World Health Organization, 70-75% of cancer patients need radiotherapy. To date, radiotherapy is the leading method of conservative treatment and for cancer patients both radical and palliative variants of such treatment are used. Every fifth patient can be recovered by radiotherapy. The advantages of radiotherapy include patients relatively tolerability to treatment, possibility of a sustainable cure of some tumors in the early stages while maintaining the function of the injured organ, etc. And with independent palliative therapy in inoperable cases, the use of radiation treatment is most effective among other medical interventions. Specifically, boron neutron capture therapy (BNCT) is a multidisciplinary task. For its realization, it is necessary to co-operate specialists of different profiles - doctors, biologists, physicists, chemists, and engineers. The priority areas of research are: development the new delivery agents of boron in order to increase the effectiveness of this therapy, reducing the risk of damaging of healthy cells, using the new neutron sources and improving the existing ones, dosimetry), resolving a number of applied problems of this therapy, and the implementation of obtained results in clinical practice. The implementation of such methods allows achieving high results of treatment if the provided radiation dose is satisfactorily tolerant for the normal tissues adjacent to the tumor. BNCT is practically a non-alternative method for treating patients with brain tumors such as, for example, multiform glioblastomas and anaplastic atrocytomas. BNCT essentially improves efficiency of recovery in combination with other modern treatment options. Successful clinical trials of treatment with BNCT of numerous liver metastases (which to this day remain practically incurable) have been started. Analysis of the materials of the International Scientific Conferences and Meetings on BNCT shows that this direction is being developed in about 30 scientific centers and more than 200 laboratories in the World. Range of applications of this therapy is expanding more and more, accelerating technique that forms a beam of epithermal neutrons is being improved, intensive synthesis and selection of those new drugs that have the ability to be accumulated in the tumor are continuing. Nuclear reactors, cyclotrons with neutron-forming targets, linear accelerators, isotopic sources, etc. are used as neutron sources. In the paper, the significant role that is assigned to radiation therapies, in general, in the fight against tumor diseases and, in particular, the contribution that BNCT can make in this fight, is analyzed. Briefly are characterized the available, including the latest, literature. The radiotherapy is characterized. The therapeutic effects of irradiation by heavy particle fluxes, in general, and neutrons, in particular, and also the basic parameters of the equipment necessary to achieve this goal are described. The methodology the actual treatment by the boron neutron capture therapy approach has been handed over, neutron sources with the necessary parameters are characterized as well, conditions for chemical compounds intended for use in BNCT and methods for measuring boron content in biological objects are described, and information is provided on modern international centers of the BNCT. There are presented the interviews with Georgian radiation therapy experts and, to the possible extent, summarizes and generalizes their responses. There are formulated recommendations for the implementation of the BNCT in Georgia developed by the author based on the analysis of information from the scientific literature and the results of the interviews with experts that he had conducted. These recommendations address all the main aspects of the problem such as: conditions for the operation of the equipment, composition and qualifications of the operating personnel, technical cooperation and communication with clinics and associations working in the

appropriate direction, implementation of procedures, preparation, provision of equipment resources, geographic factors, and other issues related, for example, to public relations, etc. In author's afterword and conclusions, the key factors that contribute to and impede the implementation of the BNCT in Georgia are sharpened. Tab. 3, Fig. 6, Ref. 16.

Keywords: boron neutron capture therapy (BNCT), non-alternative method, beam of epithermal neutrons, scientific literature, biological objects, recommendations

References

1. G.L. Locher. Biological effects and therapeutic possibilities of neutrons. Am. J. Roentgenol. Radium Ther., 1936, 36, 1-13.
2. L. Chkhartishvili. Interaction between neutron-radiation and boron-containing materials. In: Radiation Synthesis of Materials and Compounds (Eds. B.I. Kharisov, O.V. Kharissova, U.O. Mendez), 2013, Boca Raton, CRC Press – Taylor & Francis Group, Ch. 3 – 43-80.
3. IAEA-TECDOC-1223. Current status of neutron capture therapy. 2001, Vienna, International Atomic Energy Agency.
4. Ed. N.S. Hosmane. Boron Science: New Technologies and Applications. 2011, Boca Raton, CRC Press.
5. Eds. W.A.G. Sauerwein, A. Wittig, R. Moss, Y. Nakagawa. Neutron Capture Therapy. Principles and Applications., 2012, Berlin–Heidelberg, Springer-Verlag.
6. С. Ульяненко, С. Корякин. Нейтрон-захватная терапия. Актуальные проблемы и возможные подходы к перспективам развития медицинской технологии, 2012, Saarbruken, Palmarium Acad. Publ.
7. R.F. Barth, A.H. Soloway, R.G. Fairchild, R.M. Brugger. Boron neutron capture therapy for cancer: Realities and prospects. Cancer, 1992, 70, 2995-3007.
8. A.H. Soloway, W. Tjarks, B.A. Barnum, F.G. Rong, R.F. Barth, I.M. Codogni, J.G. Wilson. The chemistry of neutron capture therapy. Chem. Rev., 1998, 98, 1515-1562.
9. J.A. Coderre, G.M. Morris. The radiation biology of boron neutron capture therapy. Radiat. Res., 1999, 151, 1-18.
10. L. Dong, M.I.M. Craig, D. Khang, C. Chen. Applications of nanomaterials in biology and medicine. J. Nanotechnol., 2012, 816184, 1-2.
11. M.A. Abdel-Wahhab, F. Marquez. Nanomaterials in biomedicine. Soft Nanosci. Lett., 2015, 5, 53-54.
12. R.N. Grimes. Boron clusters come of age. J. Chem. Educ., 2004, 81, 5, 657-672.
13. G. Wu, R.F. Barth, W. Yang, R.J. Lee, W. Tjarks, M.V. Backer, J.M. Backer. Boron containing macromolecules and nanovehicles as delivery agents for neutron capture therapy. Anticancer Agents Med. Chem., 2006, 6, 167-184.
14. J. Ch., L. Chow. Application of nanoparticle materials in radiation therapy. In: Handbook of Ecomaterials (Eds. L. Martinez, O. Kharissova, B. Kharisov), 2018, Cham, Springer, 1-21.
15. გიგილაშვილი. ნეიტრონების შთანთქმის თერაპიაში ბორშემცველი ნანოსტრუქტურების გამოყენების შესაძლებლობები (მაგისტრის აკდემიური ხარისხის მოსაპოვებლად წარმოდგენილი ნაშრომი), 2017, თბილისი, საქ. ტექ. უნივ.
16. J. E. Martin. Physics for Radiation Protection. A Handbook, 2006, Weinheim, Wiley-VCH Verlag GmbH & Co. KGaA.

1.6.1.2. Influence of physical factors on blood outdraining volume. /M. Chikhladze, M. Chikhladze, O. Ketiladze/. Nano Studies. – 2019. – #19. – pp. 259-262. – geo.; abs.: geo.

The paper deals with influence of blood pressure gradient and hemodynamic resistance on the blood outdraining volume. Based on physical statements, blood motion and mechanism of arising of pressure gradient are explained. Normal values of blood speed and of "excess" pressure are given for different blood vessels. Fig. 1, Ref. 3.

Keywords: physical factors, vlood outdraining volume, pressure gradient, normal valued blood vessels

References

1. Т. Педли. Гидродинамика крупных кровеносных сосудов. 1983, Москва, Мир.
2. В.Г. Лещенко, Г.К. Ильич. Медицинская биологическая физика. 2012, Минск – Москва, Новое знание – ИНФРА-М.
3. მ. ჩიხლაძე, მ. გიგინეიშვილი, ო. კეთილაძე. სამედიცინო ფიზიკის საფუძლები, 2018, თბილისი, ტექნიკური უნივერსიტეტი. ISBN 978-9941-20-998-7. CD4522.pdf.

Author Search

Abdullayev N. 1.2.1.6.
Abramishvili R. 1.1.2.3.
Adamyan Z. 1.5.1.4.
Aghayev M. 1.2.1.4.
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Braude I. 1.1.2.1.
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Chedia R. 1.4.1.1.
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Chikhladze M. 1.6.1.2.
Chirakadze A. 1.4.1.7.
Chkhaidze S. 1.1.2.3.
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Gaune-Escard M. 1.4.1.2.
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Gelagutashvili E. 1.4.1.5.
Gevorgian L. 1.1.2.10.
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Gevorgyan L.1.5.1.7.
Gigineishvili A. 1.1.2.6.
Ginturi E. 1.3.1.2.
Gobronidze V. 1.4.1.7.
Gongadze N. 1.2.1.3.
Guliyeva T. 1.4.1.8.
Gurielidze M. 1.3.1.2.
Gventsadze D. 1.4.1.4.
Haroyan H. .1.5.1.6.
Harutyunyan V. 1.1.2.12.
Hasanova I. 1.2.1.4.
Hayrapetyan D. 1.1.1.1., 1.1.2.12.
Hryhorova T. 1.1.2.5.
Huseynova Z. 1.4.1.8.
Ibragimov T. 1.1.2.15.
Iluridze G. 1.1.2.6.
Imamaliyev A. 1.1.2.15.
Ishenko N. 1.4.1.8.
Ismailov T. 1.1.2.16.
Jabua Z. 1.1.2.6.
Jalabadze N. 1.4.1.6.
Jishiashvili A. 1.4.1.7.
Jishiashvili D. 1.4.1.7.
Joost B. 1.5.1.4.
Kagramanov K. 1.2.1.6.
Kakhramanov N. 1.1.2.14.
Kazaryan E. 1.1.2.12.
Ketiladze O. 1.6.1.2.
Khachaturyan E. 1.5.1.4.
Khachidze T. 1.5.1.3.
Khaldeeva I. 1.1.2.2.
Khalilova K. 1.2.1.6.
Klinger E. 1.5.1.5.
Korostil A. 1.1.2.4.
Kuchava N. 1.3.1.1.
Kurbanova N. 1.4.1.8.
Kutelia E. 1.4.1.4.
Kveselava V. 1.2.1.3.

Lavrynenko O. 1.2.1.1.
Lavrynenko O. 1.2.1.2.
Leonov A. 1.5.1.2.
Lototskaya V. 1.1.2.1.
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Mahmudov H. 1.2.1.4.
Maisuradze J. 1.5.1.1.
Makhatadze N. 1.4.1.7.
Malyshev V. 1.4.1.2., 1.4.1.3.
Matevosyan L. 1.1.2.11.
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